

THE
THERMODYNAMIC PROPERTIES
OF AMMONIA

COMPUTED FOR THE USE OF ENGINEERS FROM NEW EXPERIMENTAL DATA DERIVED FROM INVESTIGATIONS MADE AT THE MASSACHUSETTS INSTITUTE OF TECHNOLOGY

BY

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PREFACE

These tables of the properties of saturated and superheated ammonia are based for the most part on an experimental investigation carried out during the course of several years in the Research Laboratory of Physical Chemistry of the Massachusetts Institute of Technology. This investigation was suggested by Professor Edward F. Miller of the Institute's Mechanical Engineering Department, which generously supplied many of the facilities needed in prosecuting the experimental investigation.

The original intention was to determine the vapor-pressure curve and the specific heat-capacity of liquid ammonia with the view of utilizing the results obtained as a partial basis for the computation of a new table of the thermodynamic properties of ammonia which would prove useful in controlling the performance of refrigerating machines. After the completion of the preliminary work, in connection with which the already existing data had been critically examined, it appeared desirable to carry out a more comprehensive experimental investigation. Throughout the whole work we have been indebted to Professor Miller for his advice and support.

The experimental work was carried out by Henry A. Babcock, Harvey S. Benson and Robert B. Brownlee, senior and graduate students in the Mechanical Engineering Department, under the direction of Frederick G. Keyes, a member of the Research Laboratory staff. Mr. Babcock took up the portion of the work bearing on the heat-capacity of liquid ammonia. Messrs. Benson and Brownlee began the determination of the vapor-pres-
sures, the liquid specific volumes, and the isotherms of the substance and continued this work during the following year. Mr. Brownlee collabo-
rated with Mr. Keyes in working over all the data and in constructing
the necessary diagrams for the tables.

The computation of the tables was carried out by George W. Clark, Instructor in the Mechanical Engineering Department of the Institute. Mr. Clark's task was especially difficult because of the form of the equation of state employed; and it was carried out by him with great skill and intelligence. This part of the work was aided financially by a generous grant from the Rumford Fund of the American Academy. The computed values have been thoroughly and independently checked by F. G. Keyes.

The experimental methods employed and the details of the data ob-
tained will, it is hoped, soon be ready for publication. It was decided,
however, to print the tables in advance of the publication of the experi-

mental research on account of the technical need of more accurate tables than have been hitherto accessible.

The treatment of the experimental results obtained and the critical study of other observers' data have resulted in some new methods of examining experimental data which are here presented in considerable detail, for it is hoped that they will be of service to others interested in similar studies. The form of the equation of state employed is very different from those which have hitherto been employed in computing tables. The usual equations employed give the volume explicitly, while the equation used in computing the present tables possesses five values of the volume. The multiple value of the volume is an obvious physical necessity from the point of view of the continuity of the phases; and a careful study of the application of the equation, not only to the vapor phase of ammonia but also to the existing data for several other substances, has shown its use to be justified. The use of the equation for practical purposes has moreover led to the development of special methods of application which greatly lighten the labor involved in computing.

The tables have been brought into the usual forms convenient for engineering practice. In addition to the tables an accurate "Mollier" diagram has been prepared which has proved to be of very material assistance for rapidly solving engineering problems.

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PART I

DISCUSSION OF THE DATA AND COMPUTATIONS

LIST OF SYMBOLS USED

dQ : Element of heat absorbed when the work dW is represented by $p dv$.

dH : Element of heat absorbed when the work dW is represented by $d(pv)$.

dU : Increment of the internal energy.

1 : Subscript referring to the vapor phase.

2 : Subscript referring to the liquid phase.

s : Subscript referring to the saturation state.

p : Pressure (or force per unit area).

v : Volume of a unit of weight.

T : Temperature reckoned on the absolute scale or on the (substantially equivalent) hydrogen scale.

C : Specific heat where the heat element dQ is considered.

γ : Specific heat where the heat element dH is considered.

Φ : Entropy defined as $\int \frac{dQ}{T}$.

L : Heat of evaporation of unit-weight of a substance.

R : Absolute gas constant defined by $pv = RT$.

δ : Correction term of the volume in the equation of state.

a and l : Constants of the cohesive pressure term $\frac{a}{(v - l)^2}$ in the equation of state.

One 15° calorie = 4.182 joules.*

One 15° B.t.u. = 777.17 standard ft. lb.

Absolute temperature of the ice-point: 273.1° C. or 459.58° F.

* This value is very nearly identical with the mean value, 4.1826 joules, obtained from a recomputation of Barnes work by A. W. Smith (*Phys. Rev.*, 1911).

The Thermodynamic Properties of Ammonia

1. FUNDAMENTAL THERMODYNAMIC RELATIONS

THE First Law of Thermodynamics requires that the heat absorbed dQ from the surroundings by a system which undergoes any change in state be equal to the increase dU in its internal energy plus the work dW produced in the surroundings. Since the work commonly consists in a change in volume dv against a pressure p , this relation is commonly expressed by the equation *

$$dQ = dU + p dv. \quad (1)\dagger$$

The quantity dQ , although of infinitesimal magnitude, is not a differential of any finite quantity which, like the internal energy U , is fully determined by the state of the system. It is therefore convenient for many purposes to consider another energy quantity dH which is defined by the equation

$$dH = dU + d(pv) \quad (2)$$

and which therefore is related to dQ in the way expressed by the equation

$$dH = dQ + v dp.$$

The quantity dH is evidently a complete differential, — one whose value is fully determined by the change in state of the system (since the values of dU and $d(pv)$ are so determined).

Various useful relations may be deduced from these equations by expressing the state of the substance (constituting the system) in terms of the variables p , v , and T . We shall consider in connection with equation (1) first the case where the independent variables are p and T and then the case where they are v and T .

Differentiating equation (1) with respect to T we get

$$\left(\frac{\partial Q}{\partial T}\right)_p = C_p = \left(\frac{\partial U}{\partial T}\right)_p + p \left(\frac{\partial v}{\partial T}\right)_p. \quad (3)$$

* It is customary in writing thermodynamic equations to sometimes insert a factor J or its reciprocal, depending on the units employed. In the equations here presented this factor has not been inserted. The simplification may be obtained by suitably choosing the units.

† In representing this heat-quantity Clausius wrote dQ to indicate that the quantity of heat was infinitesimal. However, dQ is not the differential of a known finite quantity Q , and some writers make use of other notations to avoid a misunderstanding of the quantity. The element dH , on the other hand, is evidently a perfect differential.

But

$$dU = \left(\frac{\partial U}{\partial T}\right)_p dT + \left(\frac{\partial U}{\partial p}\right)_T dp.$$

and it may be easily shown * that $\left(\frac{\partial U}{\partial p}\right)_T = -T\left(\frac{\partial v}{\partial T}\right)_p - p\left(\frac{\partial v}{\partial p}\right)_T$, and since $dv = \left(\frac{\partial v}{\partial T}\right)_p dT + \left(\frac{\partial v}{\partial p}\right)_T dp$, (1) may be written

$$dQ = C_p dT - T\left(\frac{\partial v}{\partial T}\right)_p dp. \quad (4)$$

From (4) may be written the specific heats along the saturation line of either the liquid or vapor,† or

$$C_s = C_p - T\left(\frac{\partial v}{\partial T}\right)_p \frac{dp}{dT} = C_p - \left(\frac{\partial v}{\partial T}\right)_p \cdot \frac{L}{v_1 - v_2} \quad (5)$$

since $T \frac{dp}{dT} (v_1 - v_2) = L$. Thus the difference in the specific heats along the saturation lines of liquid and vapor may be written

$$C_{s_1} - C_{s_2} = C_{p_1} - C_{p_2} - \frac{L}{v_1 - v_2} \left[\left(\frac{\partial v_1}{\partial T}\right)_{p_1} - \left(\frac{\partial v_2}{\partial T}\right)_{p_2} \right]. \quad (6)$$

The specific heat of a liquid as measured is usually that defined by (5), while it is impossible to measure C_{s_1} directly. By means of the equation

$$C_{p_1} - C_{p_2} = \frac{dL}{dT} - \frac{L}{T} + T \frac{dp}{dT} \left[\left(\frac{\partial v_1}{\partial T}\right)_{p_1} - \left(\frac{\partial v_2}{\partial T}\right)_{p_2} \right]$$

one obtains easily from (6) $C_{s_1} - C_{s_2} = \frac{dL}{dT} - \frac{L}{T}$.

It is evident that since $\frac{dL}{dT}$ is negative, C_{s_1} may be negative or positive.

If C_{s_1} is negative the vapor would superheat on compression.

From (4) the expression for the entropy becomes

$$\Phi = \int \frac{dQ}{T} = \int \frac{C_p dT}{T} - \int \left(\frac{\partial v}{\partial T}\right)_p dp + \Phi_0, \quad (7)$$

and accordingly it follows on differentiating that

$$\left. \begin{aligned} \left(\frac{\partial \Phi}{\partial T}\right)_p &= \frac{C_p}{T}, \\ \left(\frac{\partial \Phi}{\partial p}\right)_T &= -\left(\frac{\partial v}{\partial T}\right)_p. \end{aligned} \right\} \quad (8)$$

The first of these equations differentiating with respect to p and the second with respect to T may be equated, giving the equation

$$\left(\frac{\partial C_p}{\partial p}\right)_T = -T\left(\frac{\partial^2 v}{\partial T^2}\right)_p.$$

* Max Planck, "Thermodynamik," 3rd ed., page 128.

† The subscript 1 will refer to the vapor phase and 2 to the liquid phase.

The first of equations (8) is a necessary condition of consistency which applies to tables of thermodynamic properties. For example, the differences in the total heats at constant pressure are the values of the mean specific heat between the two temperatures at which the difference is taken. This quantity divided by the average absolute temperature should be equal to the difference in the entropies for the corresponding temperatures.

The choice of v and T as independent variables leads to a general expression for dQ . For this purpose the definition $\left(\frac{\partial U}{\partial T}\right)_v = C_v$, and the equation $\left(\frac{\partial U}{\partial v}\right)_T = T \left(\frac{\partial p}{\partial T}\right)_v - p$ are needed. Proceeding in the same manner as with equation (4) there is obtained

$$dQ = C_v dT + T \left(\frac{\partial p}{\partial T}\right)_v dv. \quad (9)$$

From this equation it follows at once that the specific heats along the saturation line may be written

$$C_s = C_v + T \left(\frac{\partial p}{\partial T}\right)_v \frac{dv}{dT}, \quad (10)$$

and the difference of the specific heats becomes

$$C_{s_1} - C_{s_2} = C_{v_1} - C_{v_2} + T \left[\left(\frac{\partial p}{\partial T}\right)_{v_1} \cdot \frac{dv_1}{dT} - \left(\frac{\partial p}{\partial T}\right)_{v_2} \cdot \frac{dv_2}{dT} \right]. \quad (11)$$

If C_s is replaced by C_p one obtains the familiar relation

$$C_p = C_v + T \left(\frac{\partial p}{\partial T}\right)_v \cdot \left(\frac{\partial v}{\partial T}\right)_p, \quad (12)$$

and the difference between C_s and C_p is evidently

$$C_s - C_p = T \left(\frac{\partial p}{\partial T}\right)_v \left[\frac{dv}{dT} - \left(\frac{\partial v}{\partial T}\right)_p \right]. \quad (13)$$

The equation for the entropy and its derivatives becomes

$$\Phi = \int \frac{C_v dT}{T} + \int \left(\frac{\partial p}{\partial T}\right)_v dv + \Phi_0, \quad (14)$$

$$\left. \begin{aligned} \left(\frac{\partial \Phi}{\partial T}\right)_v &= \frac{C_v}{T}, \\ \left(\frac{\partial \Phi}{\partial v}\right)_T &= \left(\frac{\partial p}{\partial T}\right)_v. \end{aligned} \right\} \quad (15)$$

Applying to (15) a process similar to that employed with equations (8) leads to the equation

$$\left(\frac{\partial C_v}{\partial v}\right)_T = T \left(\frac{\partial^2 p}{\partial T^2}\right)_v.$$

6 THE THERMODYNAMIC PROPERTIES OF AMMONIA

Returning to equation (2) it follows as with (4) that

$$dH = C_p dT - \left[T \left(\frac{\partial v}{\partial T} \right)_p - v \right] dp \quad (16)$$

and $dH = \left[C_v + v \left(\frac{\partial p}{\partial T} \right)_v \right] dT + \left[T \left(\frac{\partial p}{\partial T} \right)_v + v \left(\frac{\partial p}{\partial v} \right)_T \right] dv. \quad (17)$

For the specific heats at constant pressure and constant volume the two equations lead to the equations

$$\left. \begin{aligned} \left(\frac{\partial H}{\partial T} \right)_p &= C_p, \\ \left(\frac{\partial H}{\partial T} \right)_v &= C_v + v \left(\frac{\partial p}{\partial T} \right)_v. \end{aligned} \right\} \quad (18)$$

The specific heats at constant pressure are seen to be equivalent while the specific heat at constant volume is greater than the specific heat defined from equation (1). The specific heat along the saturation line becomes

$$\begin{aligned} \gamma_s &= \frac{dH}{dT} = C_p - \left[T \left(\frac{\partial v}{\partial T} \right) - v \right] \frac{dp}{dT} \\ &= \left[C_v + v \left(\frac{\partial p}{\partial T} \right)_v \right] + \left[T \left(\frac{\partial p}{\partial T} \right)_v + v \left(\frac{\partial p}{\partial v} \right)_T \right] \frac{dv}{dT}. \end{aligned} \quad (19)$$

If the relation $pv = RT$ is applied to (16) and (17) there results for the former $dH = C_p dT$ while (17) becomes $dH = (C_v + R) dT = C_{p_1} dT$. The difference in the saturation specific heats is

$$\gamma_{s_1} - \gamma_{s_2} = C_{p_1} - C_{p_2} - T \frac{dp}{dT} \left[\left(\frac{\partial v_1}{\partial T} \right)_p - \left(\frac{\partial v_2}{\partial T} \right)_p \right] + \frac{L}{T}. \quad (20)$$

Taking account of the equation for $C_{p_1} - C_{p_2}$ in equation (5) and L one obtains the relation

$$\gamma_{s_1} - \gamma_{s_2} = \frac{dL}{dT}. \quad (21)$$

The equation $C_{s_1} - C_{s_2} = \frac{dL}{dT} - \frac{L}{T}$ provides a further relation which assists in comprehending the difference between the definition of heat contained in equations (1) and (2). Writing $\Delta C_{s_{12}}$ for this equation and $\Delta \gamma_{s_{12}}$ for (21) it follows that

$$L = T (\Delta \gamma_{s_{12}} - \Delta C_{s_{12}}). \quad (22)$$

The equations for the Joule-Thomson experiment are at once deducible from equations (16) and (17); assuming that H is constant one obtains:

$$\frac{dT}{dp} = \frac{T \left(\frac{\partial v}{\partial T} \right)_p - v}{C_p}. \quad (23)$$

$$\frac{dT}{dv} = \frac{T \left(\frac{\partial p}{\partial T} \right)_v + v \left(\frac{\partial p}{\partial v} \right)_T}{C_v + v \left(\frac{\partial p}{\partial T} \right)_v}. \quad (24)$$

The above relations serve as a general basis for using the quantity defined by either (1) or (2). The present tables are based on equation (1), which represents the heat added to a fluid within an envelope, while (2) evidently represents the quantity of heat supplied when a fluid is forced to flow from one p, v, T condition to another p, v, T condition.

General Formulae Deducible from Equation (1)

Independent Variables p, T .

$$dQ = C_p dT - T \left(\frac{\partial v}{\partial T} \right)_p dp.$$

$$C_v = C_p - T \left(\frac{\partial v}{\partial T} \right)_p \frac{dp}{dT}.$$

$$\left(\frac{\partial U}{\partial p} \right)_T = - T \left(\frac{\partial v}{\partial T} \right)_p - p \left(\frac{\partial v}{\partial p} \right)_T.$$

$$\Phi = \int \frac{C_p dT}{T} - \int \left(\frac{\partial v}{\partial T} \right)_p dp + \Phi_0.$$

$$\left(\frac{\partial \Phi}{\partial T} \right)_p = \frac{C_p}{T}.$$

$$\left(\frac{\partial \Phi}{\partial p} \right)_T = - \left(\frac{\partial v}{\partial T} \right)_p.$$

$$\left(\frac{\partial C_p}{\partial p} \right)_T = - T \left(\frac{\partial^2 v}{\partial T^2} \right).$$

Independent Variables v, T .

$$dQ = C_v dT + T \left(\frac{\partial p}{\partial T} \right)_v dv.$$

$$C_p = C_v + T \left(\frac{\partial p}{\partial T} \right)_v \frac{dv}{dT}.$$

$$\left(\frac{\partial U}{\partial v} \right)_T = T \left(\frac{\partial p}{\partial T} \right)_v - p.$$

$$\Phi = \int \frac{C_v dT}{T} + \int \left(\frac{\partial p}{\partial T} \right)_v dv + \Phi_0.$$

$$\left(\frac{\partial \Phi}{\partial T} \right)_v = \frac{C_v}{T}.$$

$$\left(\frac{\partial \Phi}{\partial v} \right)_T = \left(\frac{\partial p}{\partial T} \right)_v.$$

$$\left(\frac{\partial C_v}{\partial v} \right)_T = T \left(\frac{\partial^2 p}{\partial T^2} \right)_v.$$

For Q Constant

$$\left(\frac{dT}{dp} \right)_Q = \frac{T \left(\frac{\partial v}{\partial T} \right)_p}{C_p}.$$

For Q Constant

$$\left(\frac{dT}{dv} \right)_Q = - \frac{T \left(\frac{\partial p}{\partial T} \right)_v}{C_v}.$$

$$\left(\frac{\partial v}{\partial T} \right)_p \cdot \left(\frac{\partial p}{\partial v} \right)_T \cdot \left(\frac{\partial T}{\partial p} \right)_v = -1.$$

$$C_p - C_v = T \left(\frac{\partial p}{\partial T} \right)_v \cdot \left(\frac{\partial v}{\partial T} \right)_p = - T \left(\frac{\partial p}{\partial v} \right)_T \cdot \left(\frac{\partial v}{\partial T} \right)_p^2.$$

$$C_{p_1} - C_{p_2} = \frac{dL}{dT} - \frac{L}{T} + \frac{L}{v_1 - v_2} \left[\left(\frac{\partial v_1}{\partial T} \right)_p - \left(\frac{\partial v_2}{\partial T} \right)_p \right].$$

$$C_v - C_u = \frac{dL}{dT} - \frac{L}{T}.$$

General Formulae Deducible from Equation (2)

$$\begin{aligned}
 & \text{Independent Variables } p, T. & \text{Independent Variables } v, T. \\
 dH = C_p dT - \left[T \left(\frac{\partial v}{\partial T} \right)_p - v \right] dp. & dH = \left[C_v + v \left(\frac{\partial p}{\partial T} \right)_v \right] dT + \left[T \left(\frac{\partial p}{\partial T} \right)_v + v \left(\frac{\partial p}{\partial v} \right)_T \right] dv. \\
 \gamma_s = C_p - \left[T \left(\frac{\partial v}{\partial T} \right)_p - v \right] \frac{dp}{dT}. & \gamma_s = \left[C_v + v \left(\frac{\partial p}{\partial T} \right)_v \right] + \left[T \left(\frac{\partial p}{\partial T} \right)_v + v \left(\frac{\partial p}{\partial v} \right)_T \right] \frac{dv}{dT}. \\
 \left(\frac{\partial H}{\partial T} \right)_p = \left(\frac{\partial Q}{\partial T} \right)_p = C_p. & \left(\frac{\partial H}{\partial T} \right)_v = C_v + v \left(\frac{\partial p}{\partial T} \right)_v. \\
 \frac{1}{T} \left(\frac{\partial H}{\partial T} \right)_p = \left(\frac{\partial \Phi}{\partial T} \right) = \frac{C_p}{T}. & \frac{1}{T} \left(\frac{\partial H}{\partial T} \right)_v = \left(\frac{\partial \Phi}{\partial T} \right)_v + \frac{v}{T} \left(\frac{\partial p}{\partial T} \right)_v. \\
 \left(\frac{\partial T}{\partial p} \right)_H = \frac{T \left(\frac{\partial v}{\partial T} \right)_p - v}{C_p} = \mu. & \left(\frac{\partial T}{\partial v} \right)_H = - \frac{T \left(\frac{\partial p}{\partial T} \right)_v + v \left(\frac{\partial p}{\partial v} \right)_T}{C_v + v \left(\frac{\partial p}{\partial T} \right)_v}. \\
 \left(\frac{\partial H}{\partial p} \right)_T = - \left[T \left(\frac{\partial v}{\partial T} \right)_p - v \right] = - C_p \mu. & \left(\frac{\partial H}{\partial v} \right)_T = T \left(\frac{\partial p}{\partial T} \right)_v + v \left(\frac{\partial p}{\partial v} \right)_T. \\
 \gamma_{s_1} - \gamma_{s_2} = \frac{dL}{dT}.
 \end{aligned}$$

$H_1 - H_2 = U_1 - U_2 + p_1 v_1 - p_2 v_2.$

2. FUNCTIONAL EXPRESSIONS FOR THE CHANGE OF VAPOR-PRESSURE WITH THE TEMPERATURE

A knowledge of the relation between the pressure and the temperature along the saturation line is of first importance in constructing a table of thermodynamic quantities. From this relation may be obtained: (1) the specific volumes of the saturated vapor through the equation of state of the vapor phase; (2) values of $\frac{dp}{dT}$ for use in calculating the heat of evaporation by the Clapeyron Equation. The problem of obtaining accurate values of $\frac{dp}{dT}$ depends evidently upon the accuracy with which it is possible to relate p to T by means of a suitable equation. Considerable attention has in the past been devoted to this subject, but most of the methods of

* Since $\left(\frac{\partial p}{\partial T} \right) = \frac{p}{T} = \frac{R}{v}$ from $pv = RT$ it follows that for a perfect gas

$$\left(\frac{\partial H}{\partial T} \right)_v = T \left(\frac{\partial \Phi}{\partial T} \right)_v + R = C_v + R = C_{p_0},$$

where C_{p_0} refers to the specific heat at constant pressure as the pressure approaches zero. It may be further observed that on applying the same pressure volume relation to $\left[T \left(\frac{\partial v}{\partial T} \right)_p - v \right]$ the specific heat γ_s becomes equivalent to C_p . C_{s_1} on the other hand reduces to $C_{s_1} = C_{p_1} - R \frac{d \log p}{d \log T}$.

attacking the problem have their starting point in certain integrations of the Clapeyron equation $L = T \frac{dp}{dT} (v_1 - v_2)$, which may also be written:

$dp = \frac{L}{T(v_1 - v_2)} dT$. It is easily seen, for example, that the latter equation, L being represented by $L_0 + a_1 T + a_2 T^2 + \dots + a_n T^n$, v_1 by $\frac{RT}{p}$, and neglecting v_2 , leads to the expression

$$\begin{aligned} \log p &= \int \frac{L_0 + a_1 T + a_2 T^2 + \dots + a_n T^n}{RT^2} dT + m \dots \\ &= -\frac{L_0'}{T} + a_1 \log T + a_2' T + \dots + a_n' T^{n-1} + m, \end{aligned} \quad (25)$$

where $L_0' = \frac{L_0}{R}$ and a_1' , a_2' have an obvious significance, m being a constant of integration calculable from a single value of p at a definite temperature.

In practice the constants of formula (25) are evaluated from several smoothed data suitably spaced with reference to the temperature. Consideration reveals at once, however, that a method of evaluating the constants is to be preferred whereby the inevitable inconsistencies of the experimental data will be disclosed and eliminated. The following method of procedure was accepted in the case of ammonia as a means of accomplishing this object.

Van der Waals inferred from certain considerations relative to his equation of state that

$$\log \frac{p_c}{p} = a \left(\frac{T_c}{T} - 1 \right) \quad (26)$$

should be valid where a is a constant. It is well known, however, that a is not constant, but varies with T .* If, however, it is possible to represent a accurately as a function of T the correct relation between p and T will result. Writing (26) as $T \left(\frac{\log p_c/p}{T_c - T} \right) = a$ it is possible to calculate values of a throughout the extent of the vapor-pressure data available. For a number of substances a lies on a curve resembling a parabola.

It is a matter of experience that it is impossible to draw a representative curve through experimental data where a minimum occurs, but to avoid this difficulty it is only necessary to plot the a 's with $(T_c - T)$ as an ordinate and by extrapolation obtain the value of a indefinitely near the critical temperature. This section of the curve is fortunately very nearly linear. Assuming that a may be represented as

$$a_0 + a_1 (T_c - T) + a_2 (T_c - T)^2 + \dots + a_n (T_c - T)^n,$$

* See H. Happel, *Ann. d. Physik*, **13**, 340 (1904); also Marks, *Jour. Am. Soc. Mech. Eng.*, **33**, 563 (1911).

a_0 would be the value of a indefinitely near the critical temperature. If now the variable is changed to

$$Z = \frac{a - a_0}{T_c - T} = a_1 + \dots + a_n (T_c - T)^{n-1},$$

the difficulty of being obliged to draw a smooth curve through a minimum will be avoided.*

Rewriting (26), taking account of the Z function, there results

$$\log p = -\frac{Z(T_c - T)^2 + a_0 T_c}{T} + (\log p_c + a_0);$$

writing $a_0 T_c = \omega$ and $(\log p_c + a_0) = m$,

$$\log p = -\frac{Z(T_c - T)^2 + \omega}{T} + m. \quad (27)$$

This equation in practice is most convenient for calculations since $Z(T_c - T)^2$ may be obtained with sufficient accuracy with a 20-inch slide rule; there remaining only the division of $\omega + Z(T_c - T)^2$ by T to be carried out by logarithms. If desired, the critical constants may be absorbed in the constants of the equation

$$\log p = -\frac{b}{T} + C + dT + eT^2 + \dots \quad (28)$$

3. THE VAPOR-PRESSURE DATA FOR AMMONIA

The most reliable and systematic data in connection with the vapor-pressure of ammonia are due to Regnault.† Other measurements have been made at isolated sections of the vapor-pressure temperature curve by Faraday,‡ Blumcke,§ Brill,|| and Davies.¶ The data due to all these observers have been admirably treated by Goodenough and Mosher** and also recently by Holst.†† In the Holst treatment of the data a few new measurements carried out by Holst were included. These additions consist of three measurements between -32° and -44° C. and also one each at 19.58° and 45.05° C. The Holst treatment, however, does not lead to values which differ materially from the Mosher values although Holst perceived that the Regnault pressures above zero were too low.

The vapor-pressure values used in the present tables depend entirely on the data obtained at the Research Laboratory of Physical Chemistry of the Massachusetts Institute of Technology. Measurements of pressure were made by equilibrating the pressure exerted by the ammonia

* Sometimes, owing to inaccurate data, the value of a_0 at the critical temperature is difficult to determine. For several substances, however, it has been found that no appreciable error is made by assuming 3.00 as the value of a_0 .

† *Mem. de l'Inst. de France*, **26**, 598 (1847).

¶ *Proc. Roy. Soc.*, **78-A**, 41 (1906).

‡ *Phil. Trans.*, **135**, 170 (1845).

|| *Ann. der Physik*, **21**, 170 (1906).

§ *Wiedemann's Annalen*, **34**, 10 (1888).

** *Univ. of Illinois Bull.*, **18** (1913).

†† *Les prop. therm. de l'ammoniaque et du chlorure de méthyle*, Leiden (1914).

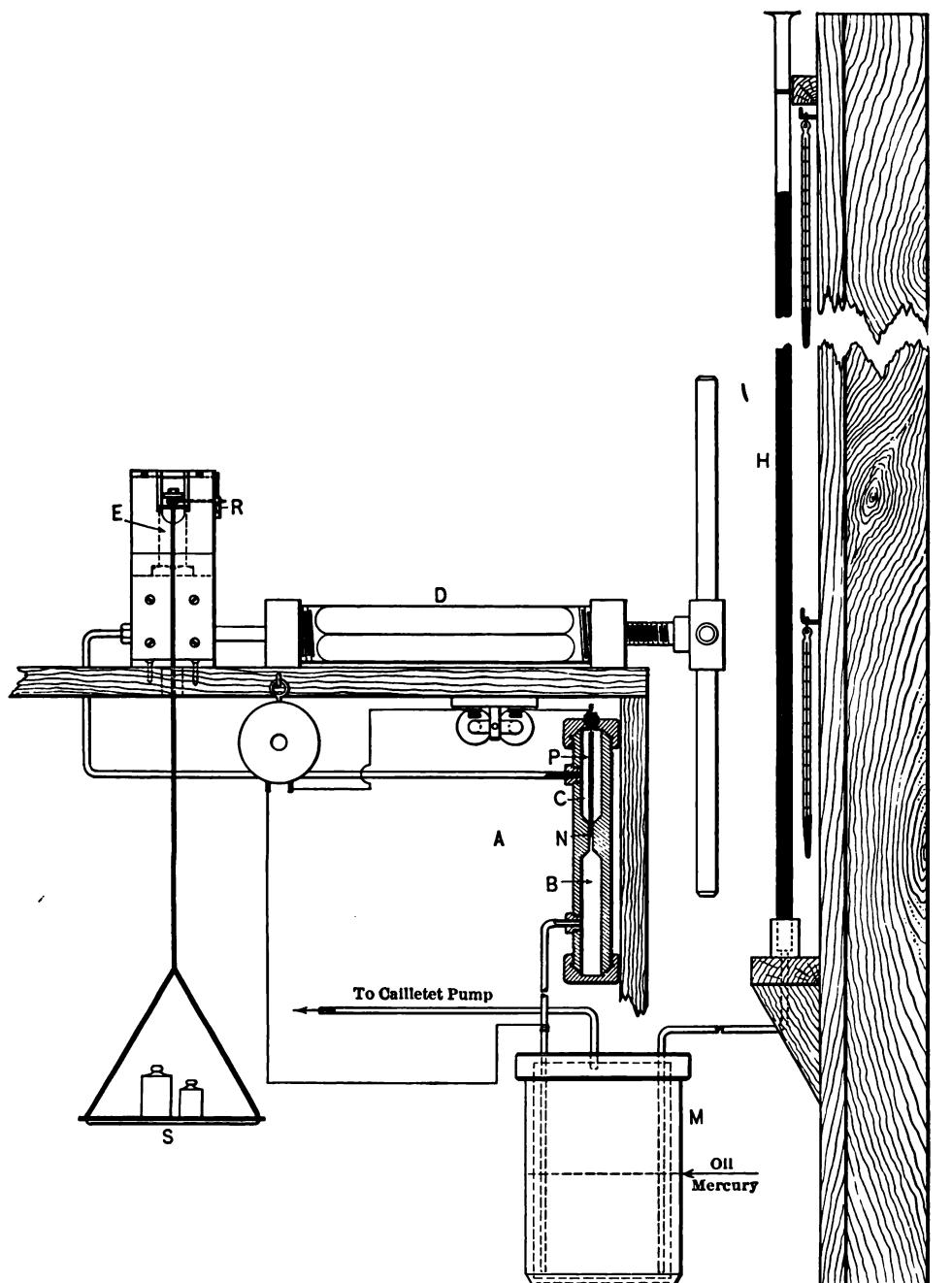


FIG. I.

against a piston by means of weights, and the final temperature measurements were made with a platinum resistance thermometer. The calibration of the pressure piston was accomplished by direct comparison with a column of mercury 12.8 meters in length. The device used for determining the equilibrium of the piston is illustrated in Fig. I.

The mercury-in-glass column is indicated at *H* and communicates with the mercury in the steel cylinder at *M*. Leading from the cover of the steel cylinder are two steel tubes, one of which passes to the Cailletet pump which serves to elevate the mercury in the column *M*, while the other leads from the mercury to the steel device *A*. From *A* the oil tube connects with the pressure-measuring piston *E*, which has attached a motor-driven device at *R* for reciprocating the piston through an angle regularly. The adjusting of the oil on the piston side of *A* is accomplished by the screw-pump *D*. The temperature of the mercury column was read by means of thermometers placed at intervals along the column. The average temperature was then obtained by graphically determining the area on a rectangular diagram between the curve drawn through the temperature readings and the axis of column length. Division of the area by the column length thus gives the true average temperature.

It has always been customary to consider the piston in equilibrium when the piston appeared to neither rise nor fall. Since the correct calibration of the piston was a matter of primary importance considerable study was devoted to the problem of investigating the sources of error that attend detecting the true equilibrium of the piston. The method finally adopted consisted in observing the motion of the mercury at its junction with the oil at *N* in the steel capillary *A* by means of a telephone receiver connected in series with the secondary of a small induction coil adjusted to the proper frequency. The connections of the circuit are evident from the drawing. An insulating joint is provided at *I* through which passes the pointed platinum wire (*p*). When the weights on the scale pan *S* are insufficient the mercury will rise in the capillary at *N* and excite the telephone receiver. If it is desired to confirm the observation the circuit is broken by the injection of a minute quantity of oil by means of the pump *D*. The weights are adjusted until the removal of 0.1 gram causes contact to be made and the addition of 0.1 gram permanently prevents contact. Since the diameter of the capillary was about 0.15 cm. while the diameter of the piston was about 0.476, a motion of 0.1 cm. of mercury in the capillary corresponds with only 0.01 cm. vertical motion of the measuring piston. The leak of oil at the piston, of course, would cause the mercury to make contact even if equilibrium had been attained. The diameter of the hole into which the piston was fitted, however, was only 0.01 mm. greater than the piston and observations on the rate of leak were made. The rate of leak under the calibrating pressures was 7×10^{-4} c.c. per hour per atmosphere and thus the arrangement permitted readings being taken rapidly and accurately. The average of two sets of the gauge calibrations agreed to about one part in eight thousand. The telephone device was employed in all the final vapor-pressure measurements and an improved thermostatic arrangement, containing the ammonia under measurement, permitted a given temperature to be maintained constant to within 0.005° C. for long periods of time. The real difficulty in making

accurate measurements lies, however, in securing true equilibrium between the vapor and liquid. To aid in securing equilibrium, the ammonia in the container was agitated by shaking the container during the course of the measurements.

The data of other observers in relation to the measurements carried out at the Research Laboratory of Physical Chemistry is illustrated in a *Z* plot, Fig. 2. It will be noted that the *Z* function serves well in making evident inconsistencies in the trend of the various observations.

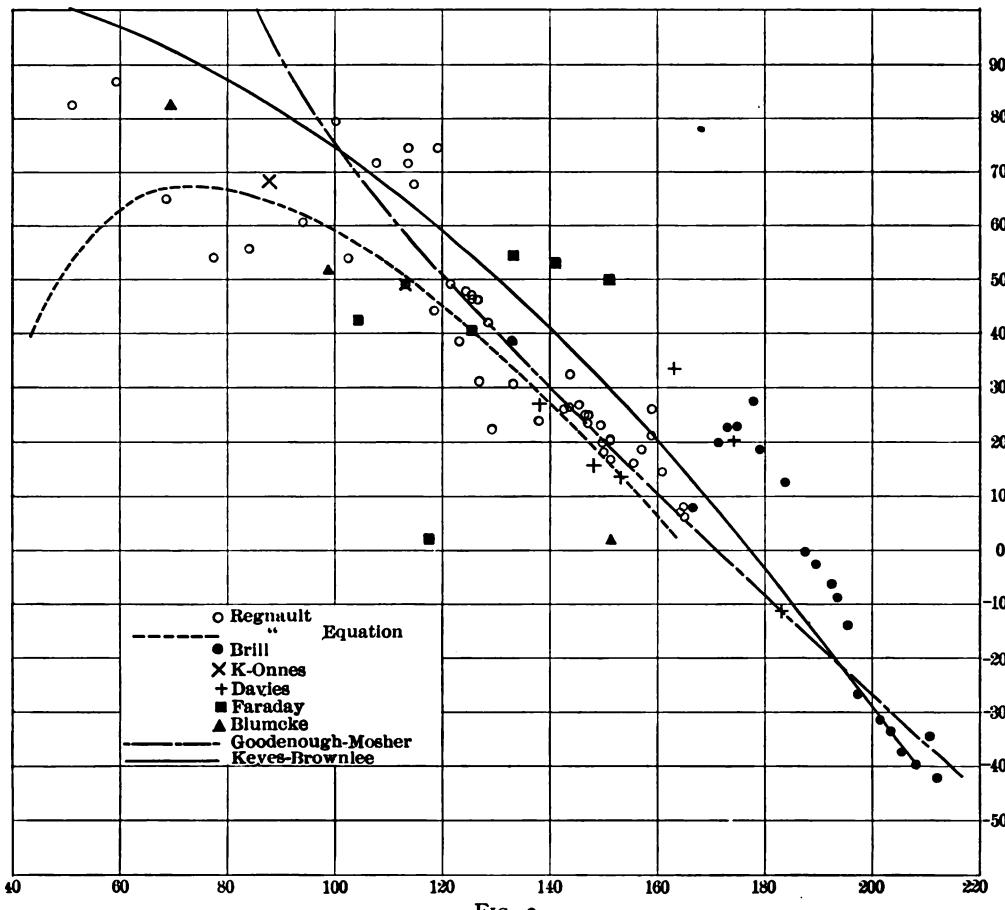


FIG. 2.

The equations of the vapor pressure are as follows:

$$\log_{10} p = 7.91121 - \frac{1209.88 + Z(T_c - T)^2}{T}. \quad (29)$$

$$Z = 10^{-4} [-11.901 + 1.0018 \cdot 10^{-2}(T_c - T) + 3.2715 \cdot 10^{-4}(T_c - T)^2],$$

or eliminating the critical constant T_c ,

$$\begin{aligned} \log_{10} p = & -\frac{1969.65}{T} + 16.19785 - 0.0423858 T + 5.4131 \cdot 10^{-6} T^2 \\ & - 3.2715 \cdot 10^{-8} T^3. \end{aligned} \quad (30)$$

The critical constants are as follows:

$$\begin{aligned}T_c &= 132.9^\circ \text{ C.} \\p_c &= 112.31 \text{ atmos.} \\v_c &= 4.236 \text{ cc. per gram.}\end{aligned}$$

The value given of the critical volume was derived from the vapor and liquid saturation specific volumes by means of the rule of the "rectilinear diameter."** The formula for $\frac{dp}{dT}$ follows from the above vapor-pressure equation in its second form and reads:

$$\frac{dp}{dT} = p \left[\frac{4535.28}{T^2} - 0.074571 + 2.49282 \cdot 10^{-4} T - 2.25987 \cdot 10^{-7} T^2 \right]. \quad (31)$$

4. THE SPECIFIC VOLUME OF THE LIQUID AMMONIA

Before discussing the experimental data the question of the empirical equation which is to represent the specific volumes of the liquid may be considered. After modifying it somewhat the equation of Avenarius seemed to be the best suited for the purpose. The Avenarius equation reads:

$$v = a + b \log (T_c - T), \quad (32)$$

where T_c is the critical temperature. Any empirical equation must satisfy the terminal conditions of the curve at the critical temperature, and must yield a finite value for the specific volume at absolute zero, assuming that superfusing could, of course, take place to that extent. The equation does satisfy the latter condition, but gives an infinite value to the volume at the critical temperature. Granting its validity, if v is plotted with the logarithm of $(T_c - T)$ on a rectangular diagram a straight line should result. It is needless to state that such a condition could scarcely be expected to hold. Since, however, the Avenarius form of equation does rectify the v, T curve to a considerable extent, the procedure adopted for the ammonia liquid volumes was to assume the equation $v = a - \log (\alpha - T)$, where α is a function of the temperature. To use the equation it is merely necessary to determine the constant a for some accurately known specific volume far below the critical temperature, arbitrarily assuming α to have the same numerical value as the critical temperature. Values of α for all the remaining data can then be calculated and plotted with the temperature as a coördinate. The curve for ammonia resembles a parabola with a minimum at about 48 degrees. The total change in α is only a few per cent, and its representation as a function of the temperature is comparatively easy. The particular advantage of the method lies in the fact that slight inconsistencies in the experimental data become at once evident. In Fig. 3 the α values corresponding to the experimental data of Dieterici and Lange are plotted on

* *Compt. rend.* 102, 1202 (1886); *Compt. rend.* 104, 1563 (1887); *Phil. Mag.*, 50, 291 (1900).

the same scale as the Research Laboratory experimental data, the latter being represented by the curve. The dotted line represents the smoothed data as given by Mosher.

It is evident that there is lack of agreement between the Lange data and the present work. It is also evident from the figure that the specific

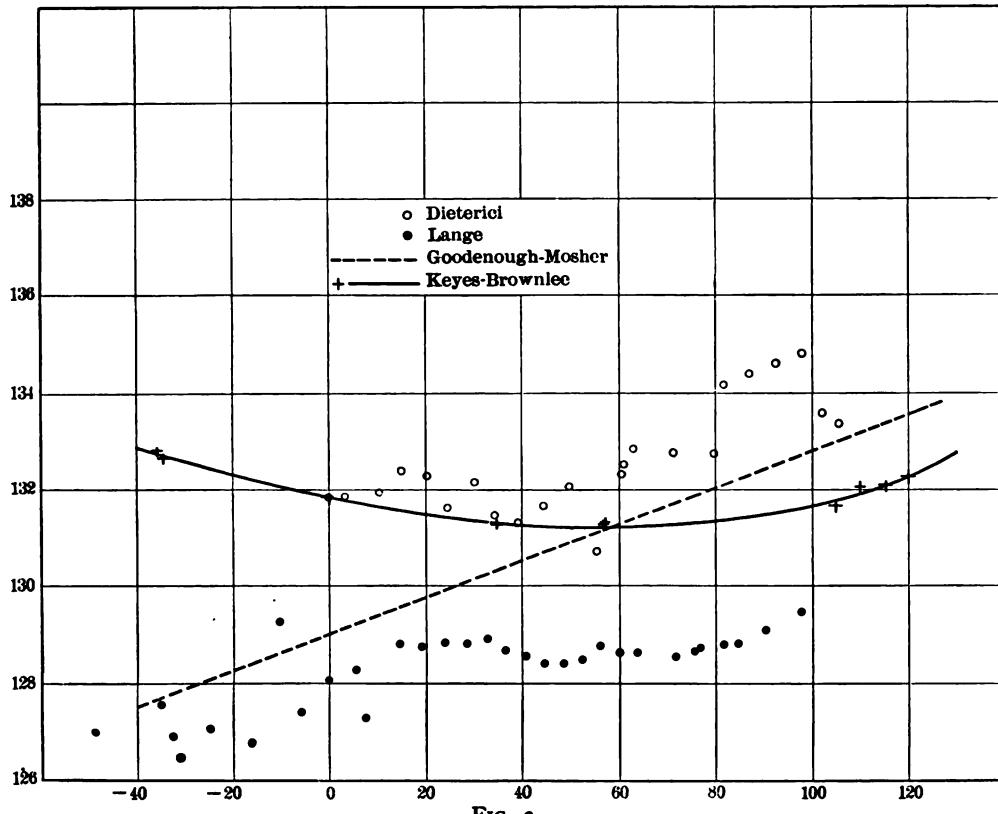


FIG. 3.

volume curve can be accurately obtained from about five exact measurements of the specific volume at suitably selected temperatures. This latter fact led Mr. Brownlee to measure accurately the volumes at the temperatures -33.5° , 0° , 35° , 68° , 110° , 120° , 125° , and the full line in the figure is drawn through the volumes found at these temperatures.

TABLE I. SPECIFIC VOLUME OF LIQUID AMMONIA

Temperature, °C.	α from R. L. of P. C. Data	Dieterici	Lange	Brownlee
-50	133.22	1.4375	1.4227
-30	132.61	1.4895	1.4745
-10	132.06	1.5480	1.5332
0	131.83	1.5656	1.5795	1.5657
+20	131.42	1.6342	1.6503	1.6387
40	131.21	1.7227	1.7383	1.7256
60	131.225	1.8250	1.8487	1.8331
80	131.40	1.9595	1.9982	1.9747
100	131.775	2.1525	2.1836
120	132.49	2.5891

Table I gives a survey of the agreement between the Lange* and Dieterici values and the values based on the Research Laboratory measurements. Attention may be directed to the agreement in the value obtained by Dieterici at 0° C. The deviations in the Dieterici measurements and the recent measurements lie in the direction of the difference between the hydrogen scale and the mercury scale. The two latter scales, of course, agree at the freezing point and the boiling point of water. Without knowing the kind of glass from which the mercury thermometer was constructed it would be difficult to correct the Dieterici data. The maximum difference between the two thermometer scales occurs at about 40° C. and would likely not exceed 0.12° C.† A greater error than the thermometric error, however, would result from the temperature expansion of the glass container used and its dilation due to pressure. The latter cannot be calculated but must be determined experimentally, and even then the glass used must be carefully annealed after having been blown. The temperature expansion of German soda glass is about 3.0×10^{-5} , which gives 0.0021 as the correction due to temperature expansion. The thermometric error amounts to 0.0006 c.c., thus giving a total of 0.0027 c.c. as the amount by which the Dieterici value is too small. Dieterici‡ gives 1.7227 as the result of smoothing his experimental data. The value arrived at in the recent work is 1.7255 c.c. Correcting the Dieterici value leads to the value 1.7254 c.c. The attempt to correct the Dieterici work at higher temperatures is difficult owing to the unknown stretch of the glass due to the increasing pressure. It will be noticed from the α figure that the Dieterici experimental values become increasingly small as the temperature increases, which would be predicted in fact owing to the pressure and temperature dilation effect.

5. EQUATIONS OF STATE IN GENERAL

The state of any substance in either of its three phases may be represented as some function of the variables p , v , and T , but for practical requirements in connection with ammonia refrigeration machines it is the vapor phase which is required to be accurately represented by such a function. The number of formulæ§ proposed are very numerous, but

* The volumes inserted in the table were calculated from Fig. 3 by drawing a representative line through the Lange values.

† See Guillaume: "Traité Pratique de la thermométrie de précision."

‡ Winkelmann, Handbuch der Physik, 3, 965.

§ Starting with the general equation of the Joule-Thomson experiment (23) the Callendar equation may be derived by assuming μ to depend on the temperature; as $\frac{\alpha}{T^n}$, and independent of the pressure. For example:

$$\frac{T\left(\frac{\partial v}{\partial T}\right)_p - v}{C_p} = \frac{dT}{dp} = \mu = \frac{\alpha}{T^n}.$$

If C_p be assumed constant the equation may be written $\frac{T dv - v dT}{T^2} = C_p \frac{\alpha}{T^{n+1}}$. Integrating

equations of the general form suggested by Callendar* seem to have received the preference. It seems not unlikely that this preference may be attributed to the fact that formulae of the Callendar type give the volume explicitly, and such an equation is most convenient in preparing tables of "properties" since volumes at constant pressure are desired. The matter of primary importance would appear, however, to concern the general consistency of the deductions and inferences which follow from the proposed equation rather than the saving of labor to the calculator in preparing tables of "properties."

The history of the subject of equations of state may be considered as included in the much broader attempt to increase our understanding concerning the continuity of matter from the point of view of an explanation of the phenomena in terms of the motion of the discreet particles of which the substances are assumed to be composed. Van der Waals,† considerations led him to a rational equation which represented the continuity of the vapor and liquid phases in its general aspects. The equation of van der Waals, however, while leading to many generalizations fails to represent accurately the p, v, T relations, even in the vapor phase, with sufficient accuracy. Many of the numerous formulæ, for the most part wholly empirical, which have appeared since van der Waals' work was published, may accordingly be regarded as an attempt to provide an equation which would represent with sufficient accuracy the p, v, T relations of substances required in technical work. The vapor phase of water for example has received much attention on account of its technical importance, and an empirical equation due to Knoblauch, Linde, and Klebe‡ has been accepted as representing the vapor phase accurately within the range required in engineering practice. The equation reads:

$$v = \frac{BT}{p} - (1 + ap) \left[C \left(\frac{373}{T} \right)^3 - D \right].$$

this equation there results $\frac{v}{T} = f(p) - \frac{C_p}{n+1} \frac{\alpha}{T^{n+1}}.$

When the temperature is high and the volume large $pv = RT$ may be assumed to represent the behavior of the gas. This identifies $f(p)$ with $\frac{R}{p}$.

The complete Callendar equation may then be written

$$v = \frac{RT}{p} - C_p \frac{\alpha}{n+1} \frac{1}{T^n}.$$

* Proc. Roy. Soc., 67, 266 (1900).

† J. D. van der Waals, Kontinuität, 1872. See also the Van't Hoff lectures.

NOTE.—The van der Waals equation reads

$$p = \frac{RT}{v-b} - \frac{a}{v^2}, \text{ where } a \text{ and } b \text{ are constants.}$$

‡ Verein deutscher Ingenieure, Heft, 21 (1905). Berlin. Also see Winkelmann, *Handbuch der Physik*, Vol. III, 1121.

This equation represents the somewhat restricted range of the measurements made by Knoblauch, Linde, and Klebe, but begins to fail as small volumes are approached.* If $\frac{\partial^2 p}{\partial T^2}$ is formed from the equation above it turns out to be a function of p and T . Examination of the Linde data just mentioned shows on the other hand that the pressure is a linear function of the temperature which would make $\frac{\partial^2 p}{\partial T^2}$ equal to zero. The consequence of $\frac{\partial^2 p}{\partial T^2}$ being equal to zero or a function of the temperature is of considerable significance because of its relation to the general equation:

$$\left(\frac{\partial C_v}{\partial v}\right)_T = T \left(\frac{\partial^2 p}{\partial T^2}\right)_v.$$

If $\left(\frac{\partial^2 p}{\partial T^2}\right)_v$ is a function of the volume, $[C_v]_T$ must be a function of the volume; but if the second derivative of the pressure is zero, C_v is a function of the temperature solely. Unfortunately it is not easy to measure C_v ,† and therefore C_v is measured. The general relation $C_v = C_p - T \left(\frac{\partial p}{\partial T}\right)_v \cdot \left(\frac{\partial v}{\partial T}\right)_p$ permits on the other hand the computation of C_v only when $T \left(\frac{\partial p}{\partial T}\right)_v \cdot \left(\frac{\partial v}{\partial T}\right)_p$ is accurately known. The strongest proof of the independence of C_v from the volume is therefore at present furnished by the linear increase of the pressure when the vapor is heated at constant volume.

* Mr. R. D. Mailey, at the Research Laboratory, has made a very careful study of the properties of water, liquid, and vapor phase, and over a range of temperature exceeding the critical temperature and to pressures above 500 atmospheres. One of the writers has had the privilege of examining the data in the vapor phase and finds the equation used in these tables to apply. This would indicate that equations of the type of the Linde Equation are defective in form.

† J. Joly. *Proc. Roy. Soc.*, **41**, 352 (1886), (1887); *Chem. News*, **58**, 271 (1888); *Proc. Roy. Soc.*, **45**, 218 (1890); *Phil. Trans.*, **182a**, 73 (1892), **185a**, 943 (1894); *Proc. Roy. Soc.*, **55**, 390 (1894).

Joly employing his steam calorimeter measured the specific heat at constant volume of air, carbon dioxide, and hydrogen. The measurements of the latter substance were not carried to completion. A. Winkelmann (Winkelmann, *Handbuch der Physik*, Vol. III, 228) discusses the air data and points out that Joly's values are too large from a comparison of the ratio C_p/C_v . This latter quantity has been measured by a number of observers and the ratio is close to 1.405. The Joly values of C_v on the other hand lead to 1.390. The values obtained by Joly moreover seem to indicate that C_v is a function of the volume. Consideration shows, however, that Joly's C_v is in reality $C_v + T \left(\frac{\partial p}{\partial T}\right)_v \frac{dv}{dT} + \Delta H$, where dv is the sum of combined thermal expansion and pressure expansion of the copper sphere which was used to contain the gas under measurements. The term ΔH represents a quantity of heat absorbed by the copper sphere containing the gas, which results from the altered heat capacity of copper under tension and also the absorption of heat due to the stretching of the copper sphere when the pressure increases from the pressure at ordinary temperature to the pressure at the final temperature of the steam. The latter quantity is small, but becomes significant at the higher pressures employed by Joly. A note discussing and recalculating Joly's experiments is in course of preparation.

The Joule-Thomson measurements* have been frequently regarded as furnishing a crucial test of the correctness of form or the accuracy with which the constants may have been determined in an accepted equation of state. The Joule-Thomson numbers indeed do furnish a sound basis for testing equations of state, but the measurements are unfortunately most difficult to make and experimenters who have occupied themselves with the problem have not always arranged to carry out the measurements in such a way as to yield numbers readily interpreted in connection with the Joule-Thomson thermodynamic equation of the porous plug experiment. For example, while the equation requires the difference in temperature of the gas before and after the plug for a small difference in pressure very often what has been measured is the difference in temperature corresponding to a large difference in pressure.†

The Joule-Thomson measurements in the case of ammonia are due to Wobsa.‡ Wobsa's measurements exhibit the anomaly of making the coefficient μ diminish with increasing pressure at constant temperature which would lead to the inference that ammonia vapor compressed at constant temperature approaches more nearly the ideal gas state.§ From measurements of the boiling point of liquid ammonia supplied in wrought iron cylinders it is possible to compute the per cent of water present by means of the Van't Hoff formula. The per cent of water appears to be of the order of 0.5 to 0.7 per cent. The presence of water accordingly in the commercial ammonia employed by Wobsa may possibly account in part for the apparently anomalous trend in the measurements.¶ The Wobsa measurements have been admirably discussed by Goodenough and Mosher and nothing can be added to their treatment until further measurements have been made.

* A lucid discussion of this quantity is given in Noyes and Sherrill's *General Principles of Chemistry*.

† W. P. Bradley and C. F. Hale, *Phys. Rev.*, **29**, 258 (1909).

‡ *Zeitschr. f. d. ges. Kalte Industrie*, **61** (1907).

§ An ideal gas is one following the equation $pV = RT$ for which μ would be zero.

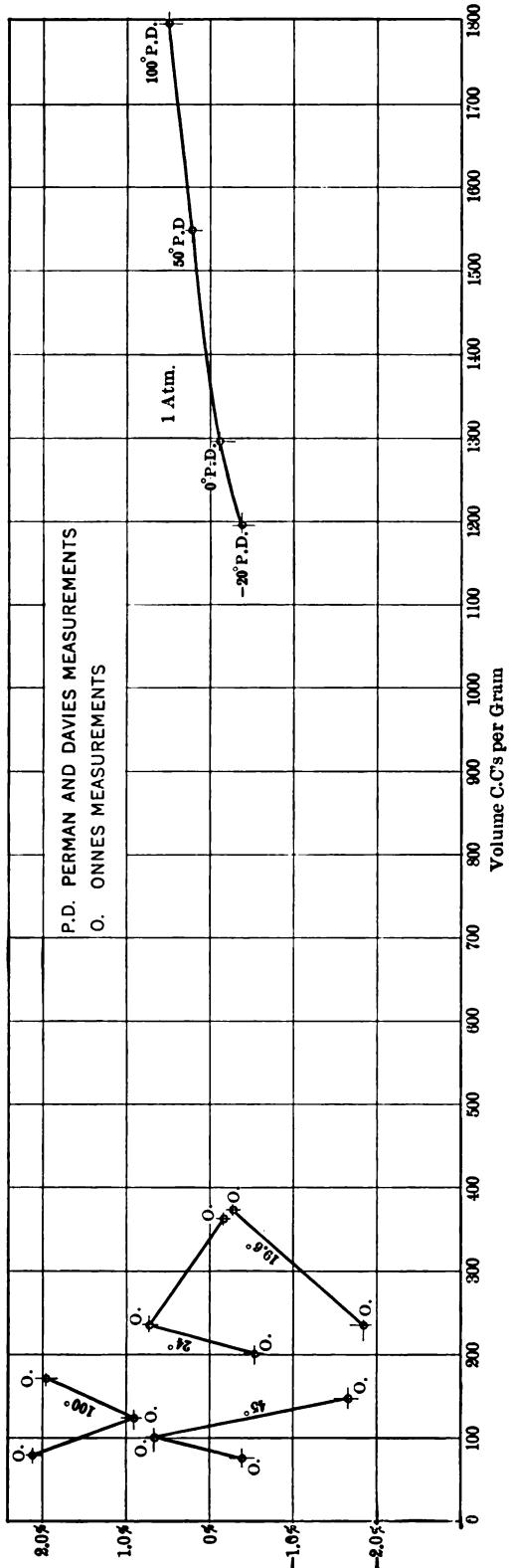
¶ In a mixture of two gases the constants of the equation (33) for any given constant composition would be a function of the constants of the components. For example, assume that (*a*) of the cohesive pressure term may be written, where *x* is the fraction of the first component:

$$a_x = a_1(x)^2 + 2 a_{12}(x)(1-x) + a_2(1-x)^2,$$

a_{12} being the cohesive pressure constant for the unlike molecules. If the attraction were large between unlike molecules, as is the case for ammonia and water, a_{12} would be many times larger than either a_1 or a_2 . The equation (33) used in connection with the Joule-Thomson equation (23) gives for moderate pressures:

$$\frac{dT}{dp} = \frac{\frac{2a}{RT} - \beta}{C_{p_0}}.$$

Now $\frac{2a}{RT}$ is the principal term of the numerator, and in a mixture, a and β would be replaced by a_x , β_x . From the comment above it is easily seen that a_x might be larger for a mixture than it would be for either pure substance alone since ammonia and water have considerable mutual affinity.



6. THE EQUATION OF STATE FOR AMMONIA VAPOR

The equation of state used in the computation of the present tables has already been briefly discussed in connection with a number of other substances.* The equation reads:

$$p = \frac{R}{v - \delta} T - \frac{a}{(v - l)^2}. \quad (33)$$

The constants of this equation for ammonia vapor have been derived from the measurements made at the Research Laboratory of Physical Chemistry of the Massachusetts Institute of Technology. The values of these constants are:

$$R = 4.8177;$$

$$\log_{10}\delta = 0.98130 - \frac{3.08}{v}$$

$$a = 34610.1; \text{ and } l = -1.173.$$

Small volumes and high pressures are best suited for the purpose of determining the constants since the deviations from the relation $pV = RT$ are greatest at small volumes. In the present case volumes less than 15 c.c. were not used in evaluating the constants of the equation above. A number of measurements were made at large volumes, but great difficulty was experienced in obtaining accurate data owing apparently to the adsorption phenomena due to the steel walls of the container.

The comparison of the work of other observers may be most

* Frederick G. Keyes, *A. S. R. E. Journal*, Vol. 1, 9 (1914).

easily compared with equation (33) by substituting the measured volumes and temperatures in the equation and comparing the pressures. It is perhaps better in the present instance, however, to compute the volumes for the measured pressures and temperatures. The result of such a comparison is given in Fig. 4, where the per cent volume difference is given at the calculated volumes. The Onnes* groups of data are substantially at constant temperature while the Perman and Davies measurements are at one atmosphere pressure. Holst states that the isotherm in the Onnes data at 45 degrees is in error. The magnitude or the nature of the error is not stated however. It is well known that ammonia is adsorbed on glass surfaces to a more marked extent than any other gas, consequently there exists the possibility that the somewhat erratic trend of the Onnes measurements may be due to this disturbing effect. If this were true it may be stated that as the temperature falls, increased adsorption would cause the volume to grow small too rapidly at constant pressure, while heating the glass bulb loaded at room temperature would cause ammonia to be given up and hence give too large a volume. Perman and Davies†

* H. Kamerlingh Onnes' "Report of the Third International Congress of Refrigeration," Sept. 15 to Oct. 1, 1913.

† Perman and Davies, to satisfy themselves that there was no adsorption, measured the density in two glass globes of different capacities. One bulb had a volume of about 0.5 liter, surface 3.22 dm², and the other 1.77 liters, surface 7.10 dm². The ratio of surfaces, accordingly, of the larger globe to the smaller is 2.2 times. Langmuir has measured the total quantity of water vapor evolved in a good vacuum from a glass surface, in passing to 360° C. The globe was a 40-watt tungsten lamp, which has a surface of approximately 1.61 dm², and the quantity of water vapor evolved amounted to 0.3 c.c. 0°/760 or about 2.41×10^{-4} grams. The weight of water is accordingly 1.5×10^{-4} grams per dm². If it is assumed that ammonia dissolves in the water film in the same manner as in liquid water, there would be 1.35×10^{-4} grams of ammonia adsorbed per dm². If ρ is the true density of the ammonia, v the volume of the globe in which it is proposed to measure the density of the ammonia, ω the weight of ammonia adsorbed per square dm. of surface, and s the surface of the globe, we may write:

$$\text{Total weight of ammonia in the globe is } \rho V + \omega s = W \text{ or } \rho = \frac{W}{V} - \omega \frac{s}{V}.$$

Let the subscript 1 denote a globe of radius r_1 , and subscript 2 denote a second sphere of larger size and radius r_2 , then $\frac{W_1}{V_1}$, the apparent density obtained in the first bulb, is equal to $\rho + \omega \frac{S_1}{V_1}$ and, similarly, $\frac{W_2}{V_2} = \rho + \omega \frac{S_2}{V_2}$.

If Δ is the difference in the measured densities,

$$\Delta = \frac{W_1}{V_1} - \frac{W_2}{V_2} = \left(\rho + \omega \frac{S_1}{V_1} \right) - \left(\rho + \omega \frac{S_2}{V_2} \right) = \omega \left(\frac{S_1}{V_1} - \frac{S_2}{V_2} \right) = 3 \omega \left(\frac{1}{r_1} - \frac{1}{r_2} \right). \quad (a)$$

The value of the density calculated by the equation of state is 0.76994, while the density found by Perman and Davies is 0.77085, the difference being 0.00091 gram. From the equation above,

$$\omega = \left(\frac{W}{V} - \rho \right) \frac{V}{S} = 0.00091 \left(\frac{V}{S} \right).$$

Now for a liter sphere $r = 0.621$ dm². Since $\frac{V}{S} = \frac{1}{3} r$ or 0.207, $\omega = 0.000188$ gram per dm².

The number obtained above by means of the Langmuir datum is of the same order of magnitude

work was carried out with the greatest care and the average of two series of duplicate measurements at zero degrees and one atmosphere are in good agreement.

Measurements have been made by Dieterici* of the specific volumes of the vapor along the saturation curve. The measurements were made in

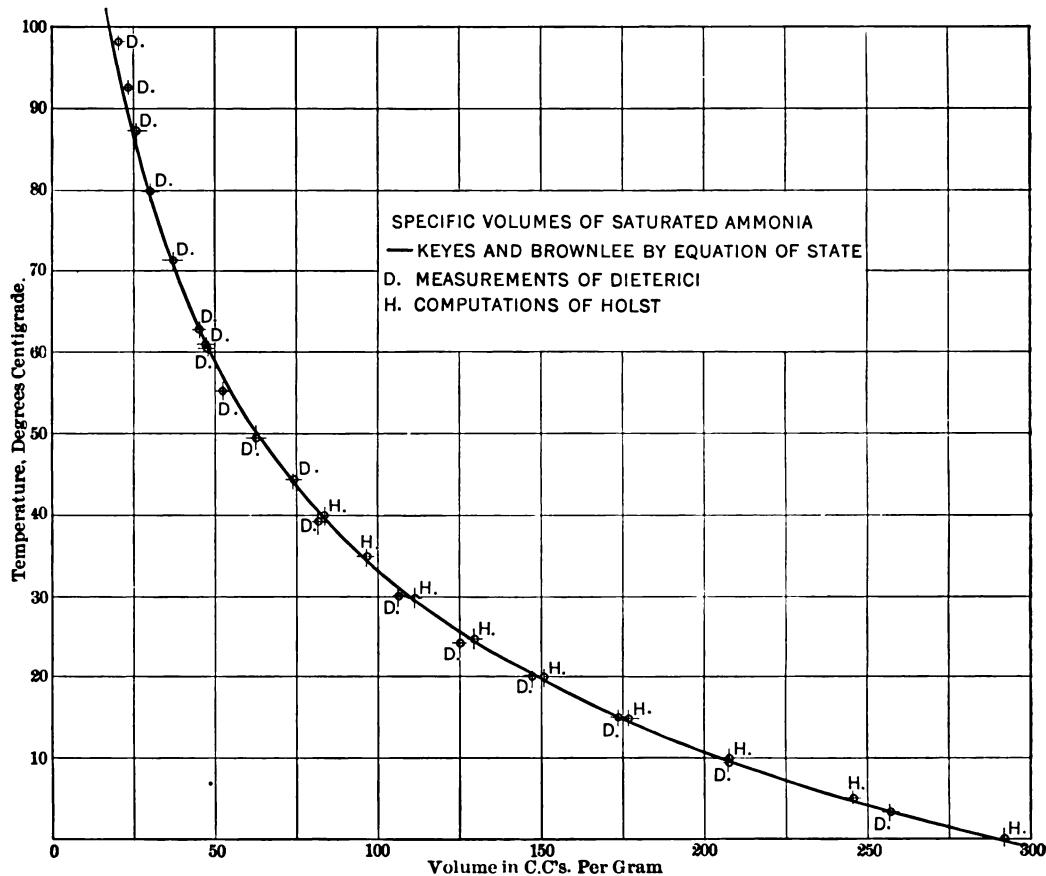


FIG. 5.

glass† and depend on the accuracy with which the liquid volumes are known. Holst has also computed by means of the Onnes virial equation

(0.000135). Applying the values to the two spheres used by Perman and Davies there results:

$$\left\{ \begin{array}{l} \omega_L = 0.000135. \\ \omega_{P.D.} = 0.000188. \\ r_1 = 0.4925 \text{ dm}^2. \\ r_2 = 0.7510 \text{ dm}^2. \end{array} \right. \quad \begin{array}{l} \Delta_{P.D.} = 3 \omega_{P.D.} \left(\frac{r_2 - r_1}{r_1 r_2} \right) = + 0.000327. \\ \Delta_L = 3 \omega_L \left(\frac{r_2 - r_1}{r_1 r_2} \right) = + 0.000235. \end{array}$$

The difference in the apparent densities therefore would amount to only about a quarter of a milligram. In view of the difficulties attending the accurate weighing of large globes it would appear that the detection of adsorption by varying the surface is not very sensitive at one atmosphere pressure and zero degrees.

* Dieterici, *Zeit. für die Desam. Kalte Industrie*, **21** (1904).

† Young, *Trans. Chem. Soc.*, **59**, 37, 126, 929 (1891).

the specific volumes of the saturated vapor,* the constants of the equation being based on the measurements of the vapor phase isotherms obtained at the Leiden Laboratory. The saturated specific volumes used in the present tables were computed by means of the equation of state (33). This computation requires the saturation pressures which were determined from the vapor-pressure equation (31). In Fig. 5 the full line is drawn through the computed saturation specific volumes while the experimental values of Dieterici are entered as indicated together with the values computed by Holst. The full line is a representative line through the Dieterici and Holst values up to about 70 degrees, when the Dieterici data assumes a distinctly different trend. The specific volumes of the liquid obtained by Dieterici on the other hand show a trend in the opposite direction.

7. THE HEAT OF VAPORIZATION OF LIQUID AMMONIA

Measurements involving the heat of vaporization of ammonia were made by Regnault, and of these measurements twelve† survived the reign of the Commune and were later published. A careful consideration of Regnault's data involving the heat of vaporization of liquid ammonia has been given by Jacobus and Denton.‡ Franklin and Kraus§ measured the quantity at the boiling point of liquid ammonia (-33.2), their value differing considerably from the value obtained by Estreicher and Schuerr.|| The original communication containing the Estreicher and Schuerr measurements is not available and hence a critical examination of the method used or a review of the data used in making necessary corrections is precluded. The method pursued by Franklin and Kraus consisted in vaporizing a definite volume of liquid ammonia at atmospheric pressure by supplying heat electrically. The calculation of the heat of vaporization requires the electrical energy, the density of liquid ammonia at -33.2° C., and the value of the calorimetric equivalent of the joule at 15 degrees. The value of the electrochemical equivalent of copper used by Franklin and Kraus was retained in the recalculation. The recomputed mean of the Franklin and Kraus measurements using the latest density data and the 15 degree cal. employed in the present tables accordingly is 336.58 Cal. at -33.2 degrees. A confirmation of the general correctness of this value may be obtained from the data concerning the elevation of the boiling point of liquid ammonia. The data in this connection is also due to Franklin and Kraus.¶ The mean value of

* The Onnes equation reads:

$$pv = RT + \frac{a_1}{V} + \frac{a_2}{V^2} + \dots$$

† *Ann. de Chim. et de Physique* (4) **24**, 375 (1871).

‡ Jacobus, *Trans. Am. Soc. Mech. Eng.*, **12**, 307 (1891).

§ *Jour. Phys. Chem.*, **2**, 555 (1907).

|| *Acad. Soc. Cracovie, Bull.*, **7A**, 345 (1910).

¶ *Am. Chem. Jour.*, **20**, 841 (1898).

k in Van't Hoff's formula $\left(L = \frac{0.019885 T^2}{k} \right)$, taken from the elevation of the boiling point where water and alcohol were used as solutes, is 3.398. The Van't Hoff formula leads to the value 336.8 Cal. as the heat of vaporization at -33.2 degrees. The value obtained by means of the Clapeyron-Clausius relation, $T \frac{dp}{dT} (v_1 - v_2) = L$, where the quantities on the left of the equation are obtained from the vapor-pressure equation (30), the equation of state (33), and the equation of the liquid volume, leads to the value 336.5 at -33.2 degrees.

The equation relating the heat of vaporization of liquid ammonia to the temperature depends on the heats of vaporization calculated by means of the Clapeyron-Clausius equation from the data obtained at the Research Laboratory. The value of L was computed at 80° , 40° , 0° , and at -70° C. The value of L at -70 degrees, on account of the uncertainty of the vapor-pressure equation at the lowest temperatures, cannot be considered to possess the same relative accuracy as the values of the heat of vaporization computed for the higher temperatures. The values calculated, however, at the temperatures mentioned were related to the temperature by means of a modified formula due to Thiesen. The Thiesen formula connecting the heat of vaporization with the temperature is $L = C (T_c - T)^n$, where c and n are constants and T_c the critical temperature. The equation satisfies the current ideas concerning the terminal conditions of the curve (L, T) — namely, it yields a finite value of the heat of vaporization at the absolute zero and a zero value at the critical temperature. Taking logarithms of both sides of the equation there results $\log L = \log C + n \log (T_c - T)$, consequently the logarithm of L is a linear function of the logarithm of $\log (T_c - T)$ and $\frac{d \log L}{d \log T} = n$.

The equation was not found to hold strictly for liquid ammonia although it does satisfy the values very nearly. To modify the equation it was assumed that the differential could be expressed as $\frac{d \log L}{d \log T} = a + b (T_c - T)$.

The resulting equation was then integrated, yielding the equation

$$\log_{10} L = 1.56817 - 2.822 \cdot 10^{-5} (T_c - T) + 0.43387 \log_{10} (T_c - T). \quad (34)$$

The values of L calculated by the latter equation together with the Regnault-Jacobus and Kraus values are given in Table 2.

Gilles Holst,* in a recent publication concerning the properties of ammonia, computed the heats of vaporization of liquid ammonia from Regnault's data, using the more accurate Dieterici specific heat values now available in computing the corrections. The average temperature of Regnault's twelve measurements is 11.68° C., and the rate of change of the heat of vaporization with temperature may be taken from equation

* Gilles Holst, *Les propriétés thermiques de l'ammoniaque et du Chlorure de méthyle*, Leiden (1914).

(34) to reduce each of Regnault's measurements to the average temperature. The result of this averaging is 295.7 Cal., whereas equation (34) leads to the value 294.3 Cal. at 11.68 degrees. The average Regnault value is accordingly 0.5 per cent higher than the value derived from the equation.

TABLE 2. VALUES OF HEAT OF VAPORIZATION L

Temperature	Calculated by equation (34)	Observed	Percentage difference	Observer
-33.4	336.7	321.3	-4.58	
-33.2	336.5	336.6	+0.03	
-33.2	336.5	336.8	+0.08	
-23.71	328.1	316.1	-3.6	
-19.55	324.5	335.1	+3.2	
-9.72	316.0	317.0	+0.3	
7.80	298.4	293.0	-1.8	
9.52	296.7	295.0	-0.6	
10.15	296.0	292.4	-1.2	
10.73	295.4	288.1	-2.5	
10.90	295.3	287.0	-2.8	
10.99	295.1	293.3	-0.6	
11.00	295.1	291.3	-1.3	
11.04	295.1	292.5	-0.9	
11.90	294.8	285.8	-2.9	
12.60	293.5	291.6	-0.6	
12.94	293.1	283.8	-3.1	
15.53	290.4	285.2	-1.8	
16.00	289.9	294.0	+1.4	
17.00	288.9	296.8	+2.7	
19.53	286.0	296.5	+3.7	
28.18	276.4	292.0	+5.8	
29.22	275.4	291.8	+5.3	
30.92	273.4	285.0	+4.2	

Holst in his summing up of the properties of ammonia has computed the heats of vaporization by means of the Clapeyron equation, using the specific volumes of the saturated vapor based on the Onnes "virial" equation of state and using values of $\frac{dp}{dT}$ derived from a careful study of Regnault's vapor-pressure measurements. Table 3 contains the Holst values compared with the values computed by means of equation (34).

TABLE 3

Temperature	Computed by Holst	Computed by equation (34)	Percentage difference
-40	328.5	341.98	-3.9
-30	322.5	333.62	-3.3
-20	316.0	324.86	-2.4
-10	309.0	315.74	-2.1
0	301.4	306.01	-1.5
+10	293.2	296.01	-0.9
+20	284.4	285.49	-0.4
+30	274.8	274.43	+0.14
+40	264.2	262.67	+0.67

Regnault's values of the vapor-pressure were perceived to be low by Holst, and it is now known that the corrections applied by Regnault to his mercury thermometer below zero were in error. This circumstance would make uncertain values of $\frac{dp}{dT}$ resulting from the vapor-pressure curve and also the values of L calculated by means of the Clapeyron equation. A glance at the Z diagram makes evident at once the erratic trend of the Regnault data.

Accurate experimental values of the heat of vaporization of liquid ammonia are necessary in the temperature interval between 0 degrees and 20 degrees; however, the Franklin-Kraus value at the boiling point and the mean Regnault value at 11.68 would indicate that equation (34) represents the heat of vaporization with substantial accuracy.

8. THE SPECIFIC HEAT-CAPACITY OF AMMONIA VAPOR

The measurements of the specific heat-capacity of ammonia have been reviewed by Nernst* and the results summarized in the equation:

$$C_{p, \text{1 atm.}} = 8.62 + 0.002 t + 7.2 \cdot 10^{-7} t^3. \quad (35)$$

The equation gives the heat-capacity for one formula weight or 17.034 grams. The equation of state (33) applied in connection with the equation $C_v = C_p - T \left(\frac{\partial p}{\partial T} \right)_v \left(\frac{\partial v}{\partial T} \right)_p$ leads to the equation for C_v as follows:

$$C_v = 0.35116 + 1.055 \cdot 10^{-4} T + 6.05 \cdot 10^{-8} T^2. \quad (36)$$

Accordingly, the specific heat at constant pressure is given by the equation

$$C_p = 0.35116 + 1.055 \cdot 10^{-4} T + 6.05 \cdot 10^{-8} T^2 + \frac{R}{\left(1 - \frac{\alpha \delta}{v^2} \right) - \frac{2 a (v - \delta)^2}{RT (v - l)^3}}, \quad (37)$$

since $\left(\frac{\partial p}{\partial T} \right)_v = \frac{R}{(v - \delta)}$ and $\left(\frac{\partial v}{\partial T} \right)_p = \frac{v - \delta}{T} \left(\frac{1}{\left(1 - \frac{\alpha \delta}{v^2} \right) - \frac{2 a (v - \delta)^2}{RT (v - l)^3}} \right)$

The quantity $\int C_p dT$ or the integral heat along a constant pressure curve must be obtained by integrating graphically the term

$$\frac{R dT}{\left(1 - \frac{\alpha \delta}{v^2} \right) - \frac{2 a (v - \delta)^2}{RT (v - l)^3}}.$$

Fig. 6 gives a picture of the C_p field. At high temperatures and low pressures the equation (37) tends to resolve into $C_p = C_v + R$ as in fact the diagram illustrates.

* *Zeitschr. f. Electrochemie*, 16, 96 (1910).

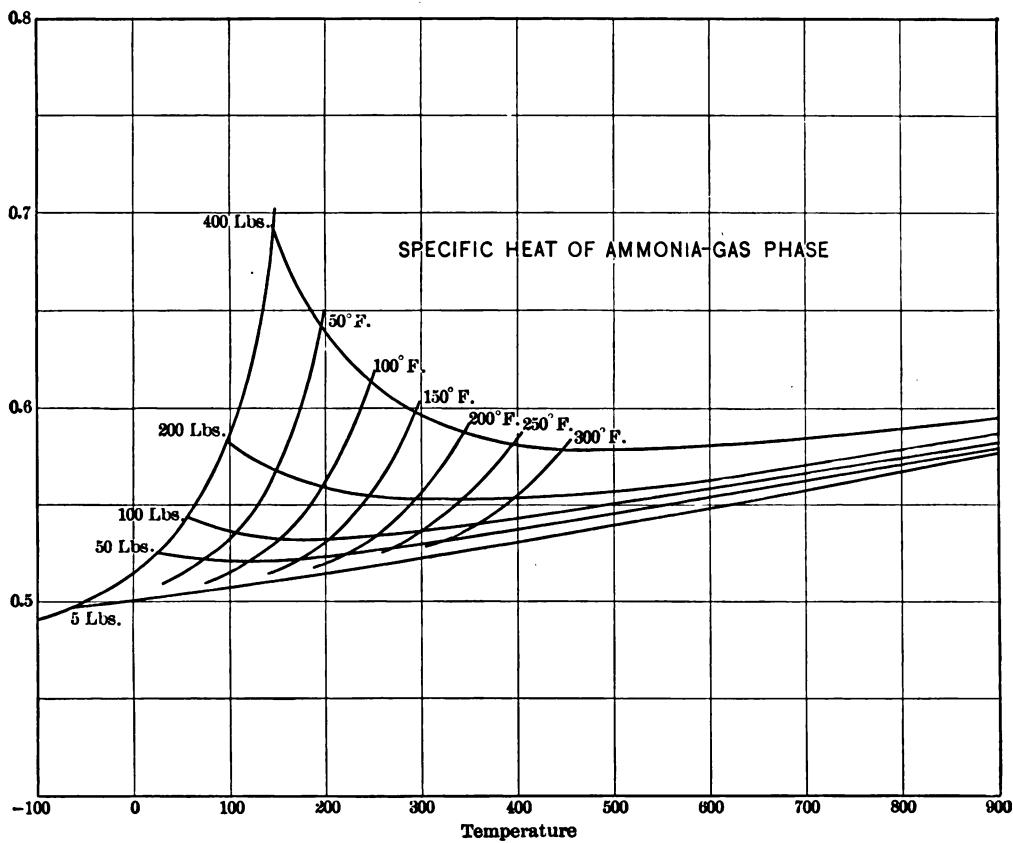


FIG. 6.

9. THE ENTROPY OF AMMONIA VAPOR

Equation (14) gives for the entropy $\Phi = \int \frac{C_v dt}{T} + R \int \frac{dv}{(v - \delta)} + \Phi_0$.

The main difficulty that arises in the use of this equation is the integration of $\frac{dv}{v - \delta}$ which contains the transcendental δ . Up to the present time no integral has been found for this expression in terms of ordinary functions. One method, however, of integrating the function is as follows:

$$\text{assume } \int \frac{dv}{v - \delta} = \log(v - \delta) + \alpha\beta \int \frac{1}{v^2} \cdot \frac{e^{-\frac{\alpha}{v}} dv}{(v - \delta)}. \quad (38)$$

The problem now is reduced to the integration of the second term of the right-hand member, and since $d \frac{1}{v} = d\rho = -\frac{1}{v^2} dv$ it follows that

$$\alpha\beta \int \frac{\rho d\rho}{e^{\alpha\rho}} \cdot \frac{1}{1 - \frac{\beta\rho}{e^{\alpha\rho}}} = \alpha \left[\int \frac{\beta\rho d\rho}{e^{\alpha\rho}} + \int \left(\frac{\beta\rho}{e^{\alpha\rho}} \right)^2 d\rho + \dots + \int \left(\frac{\beta\rho}{e^{\alpha\rho}} \right)^n d\rho \right]. \quad (39)$$

It can be easily shown that the series part of (38) converges for all values of the variable ρ .

Let $e^{\alpha\rho} = z$, whence

$$\int \frac{\beta \rho}{e^{\alpha \rho}} d\rho = \frac{\beta}{\alpha^2} \int \frac{\log z}{z^2} dz,$$

and integrating this equation in z there results

$$\begin{aligned} \int \frac{(\log z)^n}{z^{n+1}} dz &= -\left[\frac{1}{n} \left(\frac{\log z}{z} \right)^n + \frac{1}{n} \left(\frac{\log z}{z} \right)^{n-1} \frac{1}{z} + \frac{n-1}{n^2} \left(\frac{\log z}{z} \right)^{n-2} \frac{1}{z^2} \right. \\ &\quad \left. + \frac{(n-1)(n-2)}{n^3} \left(\frac{\log z}{z} \right)^{n-3} \frac{1}{z^3} + \dots \text{ to } (n+1) \text{ terms.} \right] \quad (40) \end{aligned}$$

Applying (40) to each member of the series, collecting and rearranging, leads to the expression

$$\alpha \left[\int \frac{\beta \rho}{e^{\alpha \rho}} d\rho + \dots \right] = \log \left(1 - \frac{\beta \rho}{e^{\alpha \rho}} \right) - \begin{cases} + \left(k + \frac{1}{2 \cdot 2} k^2 + \dots + \frac{(n-1)!}{n^n} \cdot k^n \right), \\ p \left(\frac{1}{2} k + \frac{2!}{3 \cdot 3} k^2 + \dots + \frac{n!}{(n+1)^n} k^n \right) \\ + p^2 \left(\frac{1}{3} k + \frac{1}{2!} \cdot \frac{3!}{4 \cdot 4} k^2 + \dots \right. \\ \quad \left. + \frac{(n+1)!}{2! (n+2)^n} \cdot k^n \right) \\ + \dots \dots \dots \dots \dots \dots \dots \dots \\ + p^m \left(\frac{1}{m+1} k + \dots \right. \\ \quad \left. + \frac{1}{m!} \frac{(n+m-1)}{(n+m)^n} k^n \right) \end{cases}$$

where $p = \frac{\beta}{\alpha} \cdot \frac{\log z}{z}$, $k = \frac{\beta}{\alpha} \cdot \frac{1}{z}$.

Table 4 represents the values of these series terms for various values of the density ρ of ammonia. In spite of the somewhat formidable appearance of the series portion of the integral it is seen that the series converge with such rapidity that the labor of calculating is not excessive. The computation of the series at 25 c.c., 50 c.c., 100 c.c., and 500 c.c. is sufficient, since beyond 500 c.c. the series gives a constant. In practice a plot of $\Sigma f(\rho)$ is most convenient. The complete expression for the entropy may be written as follows:

$$\Phi = 0.80859 \log_{10} T + 1.055 T \cdot 10^{-4} + 3.025 T^2 \cdot 10^{-8} + 0.2688 \log_{10}(v - \delta) + R\Sigma f(\rho) + \Phi_0. \quad (41)$$

For volumes greater than 200 c.c. per gram $R\Sigma f(\rho)$ may with sufficient accuracy be assured constant.

TABLE 4

$\rho=0.05$	$\rho=0.02$	$\rho=0.01$	$\rho=0.002$	$\rho=0.001$	$\rho=0.0002$	Series Equation Above
1.26177	1.70690	1.9020	2.0802	2.1057	2.1241	k series
0.27012	0.02732	0.00935	0.00042	0.00011	p series
0.07114	0.00407	0.00077	0.00001	p^2 series
0.02026	0.00063	0.00006	p^3 series
0.00608	0.00010	p^4 series
0.00182	p^5 series
0.00055	p^6 series
0.00016	0.19587	0.12140	0.02904	0.01465	0.00297	p^7 series
0.40939	0.18182	0.09344	0.01900	0.00056	0.00192	$\log\left(1 - \frac{\beta\rho}{e^{ap}}\right)$
2.04122	2.11671	2.12702	2.12876	2.12902	$\Sigma f(\rho)$

10. THE SPECIFIC HEAT-CAPACITY OF LIQUID AMMONIA

In carrying out accurate heat-capacity measurements by the method of mixtures, it is necessary to have a large difference of temperature between the thermostat and the calorimeter, in order that the resulting temperature-change may be sufficiently large to permit the necessary percentage accuracy. There is considerable difficulty involved in obtaining the precise value for the water-equivalent of the calorimeter, and also in obtaining the value of the specific heat of the steel container or other receptacle containing the ammonia to be experimented on.

The following method of measuring heat-capacities was suggested by Dr. Charles A. Kraus of the Research Laboratory. A steel bomb containing liquid ammonia under the pressure of its saturated vapor is brought to a constant temperature in a thermostat above the calorimeter; it is then dropped into a calorimeter containing a definite weight of water, and the temperature-change of the calorimeter is observed. Another steel bomb identical with the first, but containing water under the pressure of its saturated vapor, is placed in the thermostat and dropped into the calorimeter. The weight of water in the water bomb is adjusted by repeated experiments until it gives practically the same temperature-change as the ammonia bomb. Omitting corrections, the specific heat-capacity of the ammonia would vary as the ratio of the weight of water to the weight of ammonia multiplied by the heat-capacity of the water.

This method does away with many of the objections to the method of mixtures. It is, of course, dependent for operation on a large temperature-difference between the thermostat and the calorimeter, but the errors are the same or nearly the same for both the ammonia and the water experiments and consequently compensate.

Mr. Henry A. Babcock after having completed the measurements of the specific heat-capacity by the method already outlined undertook the development of a method which would permit the measurement of the heat-capacity over very small temperature intervals and at various temperatures approaching the critical temperature. The method consisted

essentially in rotating a steel ammonia container submerged in oil in a silvered Dewar tube. Measured amounts of electrical energy were introduced by means of a combined platinum resistance heater and thermometer. The amount of energy necessary to heat the oil and container was determined at a number of intervals between 20 and 120 degrees. The electrical energy necessary to raise the apparatus through one degree with the ammonia present was then obtained. The heat capacity of the

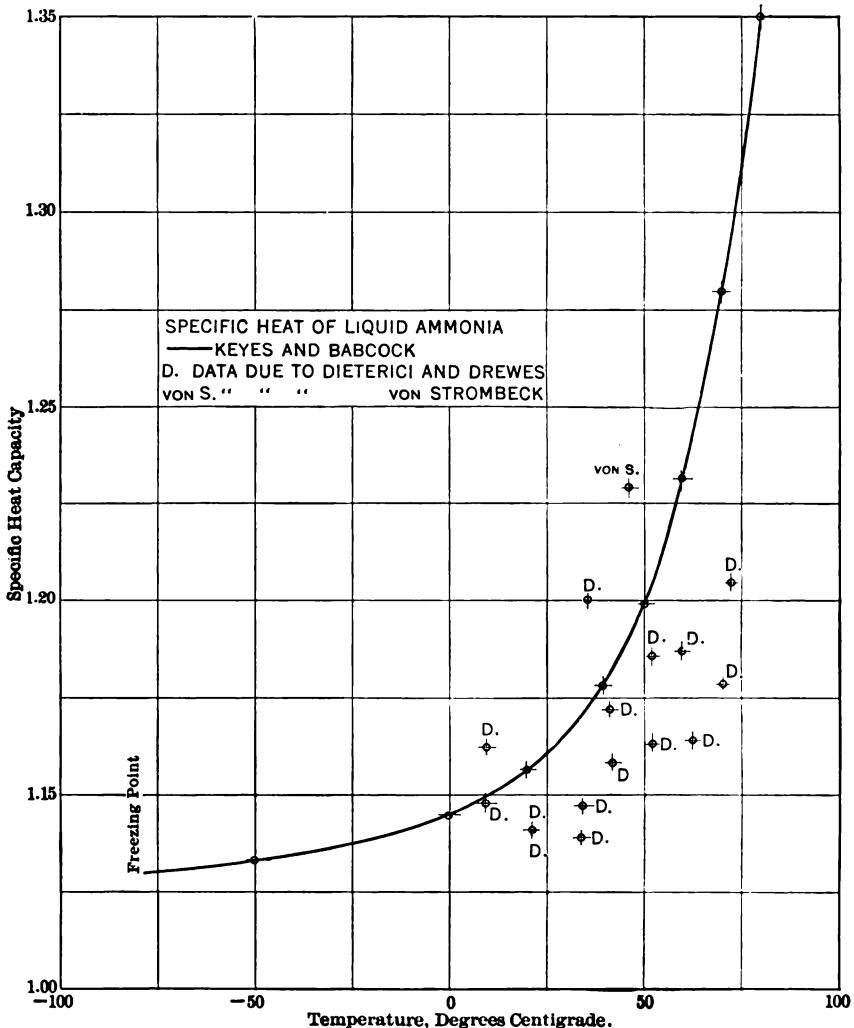


FIG. 7.

ammonia, after making the necessary corrections, is thus the difference between the two series of measurements at corresponding temperatures. These latter measurements have not been included in the equation used in the present tables, since the two more accurate values obtained by Mr. Babcock cover the practical range of temperature required in engineering work.

The equation taken to represent the saturation liquid specific heat-capacity is somewhat arbitrary. It seems probable that the heat-capacity, because of the term $\frac{dv}{dT}$, becomes infinite at the critical temperature. For this reason the empirical equation chosen to represent the heat capacity was

$$C_s = 1.13747 - \frac{5.7575}{(T_c - T)} + \frac{898.53}{(T_c - T)^2}. \quad (42)$$

This equation passes through the two measurements made by Mr. Babcock. The course of the values is illustrated by Fig. 7 in which have been inserted the values reported by Dieterici. The equation for the integral heat referred to 0°C . is

$$\int_{273.1}^T C_s dT = 1.13747 T + 13.257 \log_{10}(T_c - T) + \frac{898.53}{T_c - T} - 345.556. \quad (43)$$

Table 5 gives a list of values as smoothed by Dieterici and reported in the "Landolt and Bornstein Tabellen." The Dieterici or Drewes values should be increased by about one per cent because of the calorie in which the results are expressed. The values given by other observers are also included for comparison.

TABLE 5. SPECIFIC HEAT-CAPACITY OF LIQUID AMMONIA

	$0^\circ\text{C}.$		
Drewes.....	0.876		1.1450
Drewes.....	1.140		1.1501
Elleau & Ennis.....	1.02		1.1501
Drewes.....	1.19		1.1570
Ludeking & Starr.....	0.886		1.1642
Drewes.....	1.218		1.1664
Drewes.....	1.231		1.1796
von Strombeck.....	1.220		1.1883
Drewes.....	1.239		1.1988
Drewes.....	1.240		1.2275

PART II

TABLES OF THE THERMODYNAMIC CONSTANTS OF AMMONIA

DESCRIPTION OF THE TABLES

Table I gives the thermodynamic properties of saturated ammonia with the temperature as the argument, while Table II gives the properties with the pressure as the argument. The lower limit of the temperature table is -100° F.; the values being tabulated for each degree to 150° and to 200° for each five degrees.

The pressure table (Table II) is complete for every pound pressure from five to two hundred pounds pressure, from two hundred pounds for every two pounds to three hundred pounds pressure, for every ten pounds to five hundred pounds pressure, and for every twenty-five pounds to seven hundred pounds pressure.

The superheat table (Table III) gives the temperature, the total heat of the liquid, the vapor volume of liquid and vapor, and the entropies of the liquid and vapor corresponding from the saturation pressure to four hundred pounds. The total heat, the volume of the vapor and the entropy of the vapor is extended into the superheat three hundred degrees, every ten degrees of superheat being tabulated to two hundred degrees and every fifty degrees from two hundred degrees to three hundred degrees superheat.

In calculating the various quantities appearing in the tables large graphs were constructed from the values calculated from the equations already discussed. The vapor pressures were calculated corresponding to each 18° F. interval using equation (30). A check on the values tabulated from the graphs was subsequently obtained by calculating the pressure at temperatures nine degrees from the pressures which served to construct the graph.

The heat of the liquid was obtained by calculating the values needed for the graph from equation (43). This equation is obtained by integrating equation (42) with respect to the temperature. The values of entropy of the liquid are given by the equation which results from the integration of (42), after first dividing by the temperature. A general check on the values obtained by the method outlined was subsequently obtained by calculating the rates of change of the quantities from the corresponding equations which result by differentiation with respect to the temperature and comparing these calculated rates of change with the successive differences of the tabulated quantities.

The volumes at constant superheat (Table III) were calculated at suitable intervals from equation (33) and their reciprocal, or the density plotted against the saturation pressure. The resulting graph is nearly

a straight line which greatly facilitated the reading of the densities. The densities were afterward converted into volumes by means of a table of reciprocals.

The total heat was obtained by graphically integrating the graph of the specific heat-capacity of the vapor plotted against the temperature and adding the quantities so found to the total heat of the saturated vapor. The graph of the entropy of the vapor was calculated from equation (41).

The consistence of the superheat table was checked by applying the equation $\left(\frac{\partial \phi}{\partial T}\right)_p = \frac{C_p}{T}$ in the following manner: Values of C_p were calculated for the vapor and compared with the differences in the total heat quantities at constant pressure. These values divided by the average absolute temperature were then compared with the differences in the tabulated entropy at constant pressure.

The Mollier Chart. — The solution of many refrigeration problems is greatly facilitated by the use of the usual heat content-entropy or Mollier diagram. The diagram accompanying the tables, I, II, and III, has curves at constant pressure, curves at constant quality, and constant temperature. The ordinates are heat contents and the abscissæ are entropies.

To bring the diagram into convenient compass oblique coördinates have been used. The horizontal lines are lines of constant heat content. The oblique lines are inclined at an angle of thirty degrees to the horizontal axis. Attention is directed to the fact that twice the vertical distances are equal to distances along the oblique lines.

TEMPERATURE

Temp. °F.	Pressure, lb. per sq. in.	Sp. vol., cu. ft. per lb.	Density, lb. per cu. ft.	$\frac{1}{V_1}$	Heat content of liquid	Heat content of vapor	Latent heat of evap.	Q ₁	L	Internal energy B.t.u.			Entropy		
										Evap.	U ₁	Vapor	U ₁	Liquid	Evap.
t	p	V ₁								ϕ ₁	ϕ ₂	ϕ ₁	ϕ ₂	ϕ ₁	ϕ ₂
-100	1.132	196.9	0.00508	.00619	663.4	621.7	1.8437
-95	1.468	101.5	0.00757	.00757	659.6	617.3	1.8094
-90	1.740	132.1	0.00917	.00917	655.8	612.9	1.7750
-85	2.134	100.1	0.01100	.01100	652.0	608.4	1.7415
-80	2.595	99.9	0.01332	.01332	648.1	603.9	1.7080
-75	3.108	76.16	0.01313	.01313	644.2	599.4	1.6765
-70	3.780	64.20	0.01558	.01558	640.3	594.9	1.6452
-65	4.533	54.20	0.01834	.01834	636.3	590.4	1.6142
-60	5.378	46.00	0.02170	.02170	632.3	585.9	1.5842
-55	6.373	39.50	0.02532	.02532	628.3	581.3	1.5547
-50	7.513	33.90	0.02950	.02950	624.2	531.0	532.6	531.0	576.8	483.6	484.6	-0.2076	1.5247	1.3171	1.3171
-49	7.761	32.87	0.03042	.03042	621.1	531.3	532.4	531.3	575.8	483.8	484.8	-0.2048	1.5189	1.3141	1.3141
-48	8.014	31.89	0.03136	.03136	622.5	531.6	532.5	531.6	574.9	484.0	485.0	-0.2020	1.5132	1.3112	1.3112
-47	8.272	30.97	0.03229	.03229	621.9	531.9	532.9	531.9	574.0	484.2	485.2	-0.1992	1.5075	1.3083	1.3083
-46	8.541	30.09	0.03323	.03323	620.9	532.3	532.9	532.3	573.0	484.4	485.4	-0.1965	1.5019	1.3054	1.3054
-45	8.810	29.24	0.03420	.03420	620.1	532.6	532.1	532.6	572.1	484.6	485.6	-0.1938	1.4963	1.3025	1.3025
-44	9.088	28.42	0.03519	.03519	619.3	532.9	532.4	532.9	571.2	484.8	485.8	-0.1911	1.4907	1.2996	1.2996
-43	9.377	27.62	0.03521	.03521	618.4	533.2	533.2	533.2	570.2	485.0	486.0	-0.1884	1.4851	1.2967	1.2967
-42	9.675	26.84	0.03726	.03726	617.5	533.5	533.5	533.5	569.3	485.2	486.2	-0.1857	1.4795	1.2938	1.2938
-41	9.974	26.09	0.03833	.03833	616.7	533.8	533.8	533.8	568.4	485.4	486.4	-0.1830	1.4740	1.2910	1.2910
-40	10.282	25.36	0.03943	.03943	615.9	534.1	534.1	534.1	567.4	485.6	486.6	-0.1803	1.4685	1.2882	1.2882
-39	10.600	24.65	0.04057	.04057	615.1	534.4	534.4	534.4	566.5	485.8	486.8	-0.1776	1.4630	1.2854	1.2854
-38	10.930	23.97	0.04172	.04172	614.3	534.7	534.7	534.7	565.6	486.0	487.0	-0.1749	1.4575	1.2826	1.2826
-37	11.261	23.31	0.04290	.04290	613.4	535.0	535.0	535.0	564.6	486.2	487.2	-0.1722	1.4521	1.2799	1.2799
-36	11.601	22.67	0.04411	.04411	612.6	535.3	535.3	535.3	563.7	486.4	487.4	-0.1695	1.4467	1.2772	1.2772
-35	11.952	22.06	0.04533	.04533	611.7	535.6	535.6	535.6	562.8	486.6	487.6	-0.1668	1.4413	1.2745	1.2745
-34	12.312	21.47	0.04658	.04658	610.9	535.9	535.9	535.9	561.8	486.8	487.8	-0.1641	1.4359	1.2718	1.2718
-33	12.682	20.89	0.04787	.04787	610.1	536.2	536.2	536.2	560.9	487.0	488.0	-0.1614	1.4305	1.2691	1.2691
-32	13.056	20.33	0.04919	.04919	609.2	536.4	536.4	536.4	560.0	487.2	488.2	-0.1588	1.4252	1.2664	1.2664
-31	13.439	19.79	0.05053	.05053	608.3	536.7	536.7	536.7	559.0	487.4	488.4	-0.1561	1.4199	1.2638	1.2638

TEMPERATURE.—Continued

TABLE I

37

Temp. F.	Pressure, lb. per sq. in.	Sp. vol., cu. ft. per lb.	V ₁	Density, lb. per cu. ft.	$\frac{I}{v_1}$	Heat content of liquid	Latent heat of evap.	Q ₁	Heat content of vapor	U ₁	Internal energy B.t.u.		Entropy Vapor	Vapor ϕ_1
											Evap. U ₁	L		
-30	13.830	19.27	0.05189	0.05189	-70.5	607.5	537.0	558.1	487.6	-0.1535	1.4146	1.2611		
-29	14.232	18.77	-0.5328	-0.5328	-69.3	606.0	537.3	557.2	487.8	-.1509	1.4093	1.2584		
-28	14.645	18.28	0.5470	0.5470	-68.2	605.8	537.6	556.2	488.0	-.1482	1.4040	1.2558		
-27	15.067	17.80	-0.5618	-0.5618	-67.0	604.9	537.9	555.2	488.2	-.1455	1.3987	1.2532		
-26	15.499	17.33	0.5770	0.5770	-65.9	604.1	538.2	554.3	488.4	-.1429	1.3935	1.2506		
-25	15.940	16.88	0.05924	0.05924	-64.8	603.2	538.4	553.3	488.6	-0.1403	1.3883	1.2480		
-24	16.391	16.43	-0.6080	-0.6080	-63.5	602.3	538.7	552.4	488.8	-.1370	1.3831	1.2455		
-23	16.853	16.03	0.6238	0.6238	-62.5	601.5	539.0	551.5	489.0	-.1350	1.3779	1.2429		
-22	17.320	15.62	-0.6402	-0.6402	-61.3	600.6	539.2	550.5	489.2	-.1324	1.3777	1.2403		
-21	17.810	15.22	0.6570	0.6570	-60.2	599.7	539.5	549.6	489.4	-.1298	1.3676	1.2378		
-20	18.304	14.84	0.06738	0.06738	-59.1	598.9	530.8	548.6	489.5	-0.1272	1.3625	1.2353		
-19	18.808	14.47	-0.66910	-0.66910	-58.0	598.0	540.0	547.6	489.6	-.1246	1.3574	1.2328		
-18	19.320	14.11	0.7087	0.7087	-56.9	597.1	540.2	546.7	489.8	-.1221	1.3533	1.2302		
-17	19.840	13.76	-0.7267	-0.7267	-55.7	596.2	540.5	545.7	490.0	-.1195	1.3477	1.2277		
-16	20.371	13.43	0.7446	0.7446	-54.6	595.4	540.8	544.8	490.2	-.1169	1.3422	1.2253		
-15	20.912	13.11	0.07628	0.07628	-53.5	594.5	541.0	543.8	490.3	-0.1144	1.3372	1.2228		
-14	21.465	12.79	-0.7819	-0.7819	-52.3	593.6	541.3	542.8	490.5	-.1119	1.3322	1.2203		
-13	22.037	12.48	0.8013	0.8013	-51.2	592.5	541.5	541.9	490.7	-.1093	1.3272	1.2179		
-12	22.619	12.18	-0.8210	-0.8210	-50.0	591.8	541.8	540.9	490.9	-.1068	1.3222	1.2154		
-11	23.210	11.89	0.8410	0.8410	-48.9	590.9	542.0	539.9	491.0	-.1042	1.3172	1.2130		
-10	23.810	11.60	0.08620	0.08620	-47.8	590.1	542.3	538.9	491.1	-0.1017	1.3123	1.2106		
-9	24.420	11.32	-0.8834	-0.8834	-46.7	589.2	542.5	537.9	491.2	-.0993	1.3075	1.2082		
-8	25.052	11.05	0.9050	0.9050	-45.6	588.3	542.7	537.0	491.4	-.0968	1.3026	1.2058		
-7	25.693	10.79	-0.9259	-0.9259	-44.4	587.4	543.0	536.0	491.6	-.0943	1.2977	1.2034		
-6	26.352	10.54	0.9479	0.9479	-43.3	586.5	543.2	535.1	491.8	-.0918	1.2928	1.2010		
-5	27.027	10.30	0.09709	0.09709	-42.1	585.6	543.5	534.1	492.0	-0.0893	1.2880	1.1987		
-4	27.710	10.06	-0.9940	-0.9940	-41.0	584.7	543.7	533.1	492.1	-.0868	1.2831	1.1963		
-3	28.411	9.84	1.016	1.016	-39.8	583.8	544.0	532.1	492.3	-.0843	1.2783	1.1940		
-2	29.122	9.61	-1.040	-1.040	-38.7	582.9	544.2	531.2	492.5	-.0818	1.2735	1.1917		
-1	29.841	9.39	1.065	1.065	-37.6	582.0	544.4	530.2	492.6	-.0793	1.2687	1.1894		

TEMPERATURE.—Continued

Temp. °F.	Pressure, lb. per sq. in.	Sp. vol., cu. ft. per lb.	$\frac{1}{v_1}$	Density, lb. per cu. ft.	Heat content of liquid	Latent heat of evap.	Heat content of vapor	Internal energy, B.t.u.		Entropy	
								Q_1	L	Q_1	U_1
t	p	v_1	v_1	$\frac{1}{v_1}$	Q_1	L	Q_1	U_2	U_1	ϕ_2	$\frac{L}{T}$
0	30.57	9.18	0.1089	0.1115	-36.5	581.1	544.6	520.2	492.7	-0.0768	1.1872
+ 1	31.33	8.97	0.1115	0.1129	-35.3	580.1	544.8	528.2	492.9	-0.0743	1.1853
2	32.09	8.77	0.1129	0.1165	-34.2	579.2	545.0	527.2	493.0	-0.0718	1.1828
3	32.87	8.58	0.1165	0.1192	-33.0	578.3	545.3	526.2	493.2	-0.0693	1.1807
4	33.66	8.39	0.1192		-31.9	577.4	545.5	525.3	493.4	-0.0669	1.1785
5	34.47	8.20	0.1220		-30.8	576.5	545.7	524.3	493.5	-0.0644	1.1763
6	35.29	8.02	0.1247		-29.7	575.6	545.9	523.3	493.6	-0.0619	1.1741
7	36.14	7.84	0.1275		-28.5	574.6	546.1	522.3	493.8	-0.0595	1.1719
8	37.00	7.67	0.1304		-27.4	573.7	546.3	521.4	494.0	-0.0570	1.1697
9	37.87	7.51	0.1333		-26.2	572.8	546.6	520.4	494.2	-0.0546	1.1675
10	38.76	7.35	0.1360		-25.1	571.9	546.8	519.4	494.3	-0.0522	1.1654
11	39.68	7.19	0.1391		-24.0	571.0	547.0	518.4	494.4	-0.0498	1.1632
12	40.61	7.03	0.1422		-22.8	570.0	547.2	517.4	494.6	-0.0474	1.1610
13	41.54	6.88	0.1453		-21.7	569.1	547.4	516.4	494.7	-0.0450	1.1589
14	42.49	6.73	0.1486		-20.6	568.2	547.6	515.4	494.8	-0.0426	1.1568
15	43.47	6.594	0.1516		-19.4	567.2	547.8	514.4	495.0	-0.0402	1.1547
16	44.45	6.456	0.1549		-18.3	566.2	547.9	513.4	495.1	-0.0378	1.1527
17	45.46	6.320	0.1582		-17.2	565.3	548.1	512.4	495.2	-0.0354	1.1506
18	46.50	6.189	0.1616		-16.0	564.3	548.3	511.4	495.4	-0.0331	1.1484
19	47.54	6.061	0.1650		-14.8	563.3	548.5	510.4	495.5	-0.0307	1.1463
20	48.60	5.934	0.1685		-13.7	562.4	548.6	509.4	495.6	-0.0283	1.1442
21	49.68	5.811	0.1721		-12.6	561.5	548.8	508.4	495.7	-0.0259	1.1422
22	50.78	5.691	0.1757		-11.4	560.5	549.1	507.3	495.9	-0.0235	1.1401
23	51.90	5.576	0.1793		-10.3	559.6	549.3	506.3	496.0	-0.0212	1.1381
24	53.03	5.463	0.1830		-9.2	558.7	549.5	505.3	496.1	-0.0188	1.1361
25	54.19	5.352	0.1868		-8.1	557.8	549.7	504.3	496.2	-0.0165	1.1340
26	55.37	5.245	0.1906		-6.9	556.8	549.8	503.2	496.3	-0.0141	1.1320
27	56.57	5.139	0.1946		-5.7	555.8	550.1	502.2	496.5	-0.0117	1.1300
28	57.79	5.037	0.1985		-4.6	554.8	550.2	501.2	496.6	-0.0094	1.1280
29	59.03	4.936	0.2026		-3.5	553.9	552.4	500.2	496.7	-0.0071	1.1260

TEMPERATURE.—Continued

TABLE I

Temp. °F.	Pressure, lb. per sq. in.	Sp. vol., cu. ft. per lb.	Density, lb. per cu. ft.	$\frac{1}{v_1}$	Heat content of liquid	Latent heat of evap.	Heat content of vapor	Internal energy, B.t.u.		Entropy	
								Q ₁	L	U ₁	Evap. $\frac{L}{T}$
30	60.29	4.838	0.2067	—	—	552.8	550.5	499.2	496.9	—	1.1241
31	61.57	4.743	.2108	—	2.3	551.9	550.7	498.1	496.9	—	1.1222
32	62.86	4.650	.2151	+	1.2	550.9	550.9	497.1	497.1	+	1.1202
33	64.17	4.559	.2193	—	0.0	549.9	551.0	496.2	497.3	—	1.1183
34	65.51	4.469	.2238	—	2.2	549.0	551.2	495.1	497.3	—	1.1164
35	66.87	4.382	.2281	—	3.4	547.9	551.3	494.1	497.5	—	1.1145
36	68.26	4.298	.2327	—	4.5	547.0	551.5	493.1	497.0	—	1.1126
37	69.66	4.216	.2377	—	5.6	546.0	551.6	492.0	497.6	—	1.1107
38	71.08	4.135	.2418	—	6.8	545.0	551.8	491.0	497.8	—	1.1088
39	72.54	4.056	.2465	—	8.0	543.9	551.9	490.0	498.0	—	1.1068
40	74.01	3.979	.2513	—	9.2	542.9	552.1	488.9	498.1	—	1.1049
41	75.51	3.904	.2561	—	10.3	542.0	552.3	487.8	498.1	—	1.1030
42	77.03	3.830	.2611	—	11.5	541.0	552.5	486.7	498.2	—	1.1011
43	78.57	3.758	.2661	—	12.6	540.0	552.6	485.7	498.3	—	1.0993
44	80.14	3.688	.2711	—	13.7	539.0	552.7	484.7	498.4	—	1.0975
45	81.73	3.619	.2763	—	14.9	537.9	552.8	483.6	498.5	—	1.0956
46	83.35	3.552	.2815	—	16.0	537.0	553.0	482.6	498.6	—	1.0938
47	84.99	3.486	.2869	—	17.2	535.9	553.1	481.5	498.7	—	1.0920
48	86.65	3.422	.2922	—	18.3	534.9	553.2	480.5	498.8	—	1.0902
49	88.33	3.359	.2977	—	19.5	533.8	553.3	479.5	499.0	—	1.0884
50	90.04	3.298	.3032	—	20.7	532.8	553.5	478.4	499.1	—	1.0866
51	91.77	3.239	.3087	—	21.8	531.8	553.7	477.3	499.2	—	1.0848
52	93.53	3.181	.3144	—	23.0	530.8	553.8	476.3	499.3	—	1.0830
53	95.30	3.124	.3201	—	24.1	529.8	553.9	475.3	499.4	—	1.0812
54	97.10	3.067	.3259	—	25.3	528.7	554.0	474.2	499.5	—	1.0794
55	98.93	3.011	.3321	—	26.4	527.7	554.1	473.1	499.5	—	1.0776
56	100.79	2.957	.3382	—	27.5	526.7	554.2	472.0	499.5	—	1.0759
57	102.68	2.904	.3444	—	28.7	525.6	554.3	470.9	499.6	—	1.0741
58	104.60	2.852	.3506	—	29.9	524.5	554.4	469.9	499.8	—	1.0724
59	106.55	2.802	.3569	—	31.0	523.5	554.5	468.8	499.8	—	1.0706

TEMPERATURE.—Continued

Temp. °R.	Pressure, lb. per sq. in.	Sp. vol., cu. ft. per lb.	$\frac{1}{V_1}$	Density, lb. per cu. ft.	Heat content of liquid	Latent heat of evap.	Heat content of vapor	Internal energy, B.t.u.		Entropy	
								Q_1	L	Q_1	U_1
t	p	V_1	$\frac{1}{V_1}$	Q_1	$+32.2$	522.4	554.6	467.7	466.6	499.9	$+0.0636$
60	108.54	2.752	0.3633	-32.2	521.4	554.7	467.7	466.6	465.5	499.9	1.0688
61	110.53	2.704	0.3698	33.3	521.4	554.8	467.7	466.6	465.5	499.9	1.0671
62	112.51	2.657	0.3764	34.5	520.3	554.8	467.7	466.6	465.5	499.9	1.0654
63	114.60	2.610	0.3831	35.6	519.3	554.9	467.7	466.4	465.4	499.9	1.0637
64	116.7	2.564	0.3900	36.7	518.2	554.9	467.7	466.4	465.4	499.9	1.0620
65	118.9	2.520	0.3968	37.9	517.1	555.0	467.3	466.3	465.3	500.2	0.9747
66	121.1	2.477	0.4037	39.0	516.0	555.1	467.3	466.3	465.3	500.3	0.9736
67	123.3	2.434	0.4108	40.2	514.9	555.1	467.2	466.2	465.2	500.4	0.9725
68	125.4	2.393	0.4179	41.4	513.8	555.2	467.2	466.1	465.1	500.5	0.9714
69	127.6	2.352	0.4252	42.6	512.7	555.3	467.2	466.0	465.0	500.6	0.9703
70	129.9	2.311	0.4327	43.8	511.6	555.4	467.1	466.0	465.0	500.7	0.9692
71	132.2	2.272	0.4401	44.9	510.5	555.4	467.1	466.0	465.0	500.7	0.9681
72	134.6	2.233	0.4478	46.1	509.4	555.5	467.0	466.0	465.0	500.8	0.9670
73	136.9	2.196	0.4554	47.2	508.3	555.5	467.0	466.0	465.0	500.8	0.9659
74	139.3	2.159	0.4632	48.4	507.2	555.6	467.0	466.0	465.0	500.9	0.9648
75	141.7	2.123	0.4710	49.5	506.1	555.6	467.0	466.0	465.0	500.9	0.9637
76	144.2	2.087	0.4792	50.7	505.0	555.7	467.0	466.0	465.0	501.0	0.9626
77	146.7	2.052	0.4873	51.9	503.9	555.8	467.0	466.0	465.0	501.0	0.9615
78	149.2	2.018	0.4955	53.1	502.8	555.9	467.0	466.0	465.0	501.1	0.9604
79	151.8	1.984	0.5040	54.2	501.7	555.9	467.0	466.0	465.0	501.1	0.9593
80	154.4	1.952	0.5123	55.4	500.6	556.0	467.0	466.0	465.0	501.2	0.9582
81	157.0	1.920	0.5208	56.6	499.5	556.1	467.0	466.0	465.0	501.3	0.9571
82	159.6	1.889	0.5294	57.7	498.4	556.1	467.0	466.0	465.0	501.3	0.9560
83	162.3	1.859	0.5379	58.9	497.3	556.2	467.0	466.0	465.0	501.4	0.9549
84	165.0	1.829	0.5470	60.0	496.2	556.3	467.0	466.0	465.0	501.4	0.9538

TABLE I

TEMPERATURE.—Continued

Temp. °F.	Pressure, lb. per sq. in.	Sp. vol., cu. ft. per lb.	$\frac{1}{v_1}$	Density, lb. per cu. ft.	Heat content of liquid	Latent heat of evap.	Heat content of vapor	Internal energy, B.t.u.		Entropy $\frac{L}{T}$	Vapor ϕ_1
								Q_1	U_1		
85	167.8	1.799	0.5559	61.2	495.1	556.3	440.2	501.4	501.5	0.1184	0.9088
86	170.5	1.770	.5650	62.4	494.0	556.4	439.1	501.5	501.5	.1206	.9050
87	173.3	1.742	.5744	63.6	492.8	556.4	437.9	501.5	501.5	.1227	.9013
88	176.2	1.714	.5834	64.8	491.7	556.5	436.7	501.5	501.5	.1249	.8976
89	179.1	1.687	.5928	66.0	490.5	556.5	435.6	501.6	501.6	.1270	.8939
90	182.0	1.660	0.6028	67.1	489.4	556.5	434.5	501.6	501.6	0.1291	0.8902
91	185.0	1.633	.6124	68.3	488.2	556.5	433.3	501.6	501.6	.1312	.8865
92	188.0	1.607	.6223	69.5	487.0	556.5	432.2	501.7	501.7	.1333	.8828
93	191.0	1.581	.6325	70.7	485.9	556.6	431.0	501.7	501.7	.1354	.8789
94	194.1	1.556	.6427	71.9	484.7	556.6	429.8	501.7	501.7	.1375	.8752
95	197.2	1.532	0.6527	73.1	483.5	556.6	428.7	501.8	501.8	0.1396	0.8715
96	200.4	1.508	.6636	74.2	482.4	556.6	427.6	501.8	501.8	.1417	.8678
97	203.6	1.484	.6739	75.4	481.2	556.6	426.4	501.8	501.8	.1438	.8641
98	206.9	1.461	.6845	76.6	480.0	556.6	425.2	501.8	501.8	.1459	.8605
99	210.1	1.437	.6959	77.7	478.8	556.5	424.1	501.8	501.8	.1480	.8569
100	213.4	1.414	0.707	78.9	477.6	556.5	422.9	501.8	501.8	0.1500	0.8533
101	216.7	1.392	.718	80.0	476.5	556.5	421.8	501.8	501.8	.1521	.8497
102	220.1	1.370	.730	81.2	475.3	556.5	420.6	501.8	501.8	.1542	.8461
103	223.5	1.350	.741	82.4	474.1	556.5	419.4	501.8	501.8	.1563	.8425
104	226.9	1.330	.752	83.6	472.9	556.5	418.2	501.8	501.8	.1584	.8389
105	230.4	1.310	0.763	84.7	471.7	556.4	417.1	501.8	501.8	0.1605	0.8353
106	233.9	1.290	.776	85.9	470.5	556.4	415.9	501.8	501.8	.1625	.8317
107	237.4	1.270	.787	87.1	469.2	556.3	414.7	501.8	501.8	.1646	.8281
108	241.0	1.251	.799	88.3	468.0	556.3	413.5	501.8	501.8	.1666	.8245
109	244.7	1.232	.812	89.4	466.8	556.2	412.3	501.7	501.7	.1687	.8209

TEMPERATURE.—Continued

Temp. °F.	Pressure, lb. per sq. in.	Sp. vol., cu. ft. per lb.	$\frac{1}{v_1}$	Density, lb. per cu. ft.	Heat content of liquid	Latent heat of evap.	Heat content of vapor	Internal energy, B.t.u.		Entropy	
								Q_1	L	U_1	Liquid ϕ_1
110	248.4	1.213	0.824	90.6	465.5	556.1	411.1	501.7	0.1708	0.8173	0.9881
111	252.1	1.195	.837	91.7	464.3	556.0	410.0	501.7	.1720	.8137	.9866
112	255.9	1.176	.850	92.9	463.0	555.9	408.8	501.7	.1749	.8101	.9850
113	259.8	1.158	.864	94.1	461.8	555.9	407.5	501.6	.1770	.8065	.9835
114	263.7	1.141	.876	95.3	460.5	555.8	406.3	501.6	.1790	.8029	.9819
115	267.6	1.123	0.891	96.5	459.2	555.7	405.1	501.6	0.1811	0.7993	0.9804
116	271.6	1.106	.904	97.7	457.9	555.6	403.9	501.6	.1832	.7957	.9789
117	275.6	1.089	.918	98.9	456.6	555.5	402.6	501.5	.1853	.7921	.9774
118	279.6	1.073	.933	100.1	455.4	555.5	401.4	501.5	.1874	.7885	.9759
119	283.7	1.057	.946	101.2	454.1	555.3	400.1	501.3	.1895	.7849	.9744
120	287.9	1.042	0.960	102.4	452.9	555.3	398.8	501.2	0.1916	0.7813	0.9729
121	292.1	1.026	.975	103.6	451.6	555.2	397.6	501.2	.1937	.7777	.9714
122	296.3	1.011	.989	104.8	450.3	555.1	396.4	501.2	.1957	.7741	.9698
123	300.5	0.996	1.004	106.0	449.0	555.0	395.2	501.2	.1978	.7705	.9683
124	304.8	.982	1.018	107.2	447.7	554.9	393.9	501.1	.1998	.7669	.9667
125	309.2	0.968	1.033	108.4	446.4	554.8	392.7	501.1	0.2019	0.7633	0.9652
126	313.6	.954	1.048	109.6	445.1	554.7	391.4	501.0	.2040	.7597	.9637
127	318.0	.940	1.065	110.8	443.8	554.6	390.2	501.0	.2060	.7561	.9621
128	322.5	.926	1.080	112.0	442.5	554.5	388.9	500.9	.2081	.7525	.9606
129	327.1	.913	1.095	113.3	441.1	554.4	387.6	500.9	.2102	.7490	.9592
130	331.7	0.899	1.112	114.5	439.8	554.3	386.4	500.9	0.2123	0.7454	0.9577
131	336.4	.886	1.129	115.7	438.5	554.2	385.1	500.8	.2144	.7418	.9562
132	341.1	.873	1.145	116.9	437.2	554.1	383.8	500.7	.2165	.7382	.9547
133	345.8	.860	1.163	118.1	435.8	553.9	382.5	500.6	.2185	.7347	.9532
134	350.6	.848	1.180	119.4	434.4	553.8	381.2	500.6	.2205	.7312	.9517

TABLE I

TEMPERATURE.—Concluded

Temp. °F.	Pressure, lb. per sq. in.	Sp. vol., cu. ft. per lb.	Density, lb. per cu. ft.	$\frac{1}{v_1}$	Heat content of liquid	Latent heat of evap.	Q ₁	L	Internal energy, B.t.u.		Entropy
									Evap. U ₁	Vapor U ₁	
t	p	v ₁							ϕ ₁	Evap. $\frac{L}{T}$	Vapor ϕ ₁
135	355.5	0.836	1.106	1.20.6	433.1	553.7	379.9	500.5	0.2226	0.7277	0.9503
136	360.4	.824	1.214	121.8	431.7	553.5	378.6	500.4	.2247	.7242	.9489
137	365.4	.812	1.233	123.0	430.3	553.3	377.3	500.3	.2267	.7207	.9474
138	370.4	.800	1.252	124.2	428.9	553.1	376.0	500.2	.2287	.7172	.9459
139	375.4	.788	1.269	125.4	427.6	553.0	374.7	500.1	.2308	.7137	.9445
140	380.6	0.777	1.287	126.6	426.2	552.8	373.4	500.0	0.2328	0.7102	0.9430
141	385.7	.765	1.307	127.8	424.8	552.6	372.1	499.9	.2349	.7067	.9416
142	390.9	.754	1.326	129.0	423.4	552.4	370.8	499.8	.2370	.7032	.9402
143	396.2	.743	1.346	130.2	422.0	552.2	369.5	499.7	.2391	.6997	.9387
144	401.5	.732	1.366	131.5	420.5	552.0	368.1	499.6	.2410	.6962	.9372
145	406.9	0.722	1.386	132.8	419.0	551.8	366.7	499.5	0.2430	0.6947	0.9357
146	412.3	.712	1.406	134.0	417.6	551.6	365.4	499.4	.2450	.6892	.9342
147	417.8	.702	1.426	135.2	416.2	551.4	364.0	499.3	.2470	.6857	.9327
148	423.3	.692	1.446	136.4	414.8	551.2	362.7	499.2	.2490	.6822	.9312
149	428.9	.682	1.467	137.7	413.3	551.0	361.3	499.1	.2510	.6787	.9297
150	434.5	0.672	1.488	139.0	411.8	550.8	360.0	499.0	0.2530	0.6752	0.9282
155	463.6	.626	1.597	145.0	404.5	549.5	353.1	498.1	.2635	.6577	.9212
160	494.0	.584	1.712	151.3	396.9	548.2	346.0	497.3	.2739	.6401	.9140
165	526.0	.544	1.838	157.7	389.1	546.8	338.8	496.5	.2843	.6225	.9068
170	559.4	.506	1.976	164.2	381.1	545.3	331.4	495.7	.2947	.6049	.8996
175	594.4	0.472	2.119	170.9	372.8	543.7	323.9	494.9	0.3052	0.5872	0.8924
180	631.2	.441	2.268	177.7	364.3	542.0	316.3	494.1	.3157	.5695	.8852
185	669.5	.412	2.427	184.8	355.6	540.4	308.4	493.3	.3265	.5516	.8781
190	709.5	.385	2.597	192.2	346.5	538.7	300.3	492.5	.3376	.5335	.8711
195	751.4	.358	2.793	199.9	337.2	537.1	291.9	491.8	.3489	.5152	.8641
200	795.1	.334	2.994	207.9	327.5	535.4	283.2	491.1	.3605	.4966	

PRESSURE TABLE

PRESSURE TABLE.—Continued

TABLE II

Pressure, lb.	Temp., °F.	Sp. vol., cu. ft. per lb.	Density, lb. per cu. ft.	$\frac{1}{v_1}$	Heat content of liquid v_1	Latent heat of liquid L	Heat content of vapor Q_1	Internal energy, B.t.u.			Entropy $E_{\text{vap.}} \frac{L}{T}$	Vapor ϕ_1
								Evap.	Vapor U_1	Liquid ϕ_2		
31	+0.57	9.061	0.1104	-36.1	580.2	544.8	493.1	-0.0755	1.2610	1.1855		
32	1.88	8.798	.1137	-34.6	579.1	545.1	493.3	-.0722	1.2549	1.1827		
33	3.17	8.541	.1170	-33.1	577.9	545.4	493.5	-.0690	1.2490	1.1800		
34	4.43	8.302	.1205	-31.7	576.8	545.6	493.7	-.0659	1.2431	1.1772		
35	5.65	8.084	.1238	-30.3	575.7	545.9	493.9	-.0628	1.2374	1.1746		
36	6.83	7.870	.1271	-28.9	574.7	546.2	492.8	-0.0599	1.2321	1.1722		
37	8.00	7.667	.1305	-27.6	573.6	546.4	494.3	-.0571	1.2269	1.1688		
38	9.14	7.478	.1337	-26.3	572.6	546.6	494.4	-.0543	1.2218	1.1675		
39	10.25	7.292	.1368	-25.0	571.6	546.8	494.5	-.0516	1.2168	1.1652		
40	11.34	7.117	.1403	-23.7	570.6	547.0	494.6	-.0490	1.2119	1.1629		
41	12.42	6.951	.1438	-22.4	569.6	547.3	494.7	-0.0464	1.2067	1.1603		
42	13.48	6.797	.1471	-21.2	568.6	547.5	494.8	-.0438	1.2018	1.1580		
43	14.53	6.646	.1505	-20.0	567.7	547.7	494.9	-.0413	1.1969	1.1556		
44	15.56	6.504	.1538	-18.8	566.7	547.9	495.0	-.0388	1.1924	1.1535		
45	16.58	6.366	.1570	-17.7	565.8	548.0	495.1	-.0364	1.1879	1.1514		
46	17.59	6.239	0.1603	-16.6	564.8	548.2	495.2	-0.0341	1.1835	1.1494		
47	18.56	6.116	.1635	-15.5	563.9	548.4	495.3	-.0318	1.1792	1.1474		
48	19.48	5.998	.1667	-14.4	563.0	548.6	495.4	-.0295	1.1749	1.1454		
49	20.37	5.885	.1699	-13.3	562.1	548.8	495.5	-.0273	1.1708	1.1435		
50	21.29	5.776	.1731	-12.2	561.2	549.0	495.6	-.0252	1.1671	1.1416		
51	22.20	5.671	0.1763	-11.3	560.3	549.1	495.7	-0.0231	1.1628	1.1397		
52	23.10	5.569	.1796	-10.2	559.4	549.2	495.8	-.0210	1.1590	1.1379		
53	23.97	5.471	.1828	-9.2	558.5	549.3	496.1	-.0189	1.1552	1.1361		
54	24.84	5.376	.1860	-8.2	557.7	549.5	496.0	-.0169	1.1514	1.1343		
55	25.70	5.284	.1893	-7.2	556.9	549.7	496.1	-.0149	1.1476	1.1346		
56	26.53	5.195	0.1925	-6.2	556.1	549.9	496.4	-.0129	1.1439	1.1359		
57	27.34	5.109	.1957	-5.3	555.3	550.0	496.3	-.0110	1.1403	1.1352		
58	28.14	5.024	.1990	-4.4	554.5	550.1	496.4	-.0091	1.1367	1.1275		
59	28.95	4.941	.2024	-3.5	553.7	550.2	496.5	-.0072	1.1331	1.1259		
60	29.76	4.861	.2057	-2.6	552.9	550.3	496.6	-.0053	1.1299	1.1243		

PRESSURE TABLE.—Continued

Pressure, lb.	Temp., °F.	t	Sp. vol., cu. ft. per lb.	Density, lb. per cu. ft.	$\frac{1}{v_1}$	Heat content of liquid	Latent heat of liquid	Heat content of vapor	Internal energy, B.t.u.		Entropy	
									Q ₁	L	Evap. U ₁	Vapor U ₁
p			v ₁									
61	30.56	4.784	0.2090	- 1.7	552.1	550.4	498.6	496.7	- .0035	1.1261	1.1227	
62	31.34	4.711	.2123	- 0.8	551.3	550.5	497.8	496.8	- .0017	1.1227	1.1211	
63	32.11	4.640	.2155	+ 0.1	550.6	550.7	497.0	496.9	+ .0002	1.1194	1.1190	
64	32.87	4.571	.2188	1.0	549.9	550.9	496.2	497.0	.0020	1.1162	1.1151	
65	33.62	4.505	.2220	1.9	549.2	551.1	495.4	497.1	.0038	1.1130	1.1116	
66	34.36	4.440	0.2252	2.7	548.5	551.2	494.6	497.2	0.0055	1.1098	1.1151	
67	35.10	4.378	.2284	3.5	547.8	551.3	493.9	497.3	.0072	1.1066	1.1137	
68	35.81	4.317	.2316	4.3	547.1	551.4	493.2	497.4	.0089	1.1036	1.1123	
69	36.51	4.256	.2349	5.1	546.4	551.5	492.5	497.5	.0105	1.1005	1.1100	
70	37.22	4.197	.2383	5.9	545.7	551.6	491.8	497.6	.0121	1.0984	1.1096	
71	37.93	4.139	0.2416	6.7	545.0	551.7	491.1	497.7	0.0137	1.0947	1.1083	
72	38.64	4.084	.2449	7.5	544.3	551.8	490.4	497.8	.0153	1.0918	1.1070	
73	39.34	4.030	.2481	8.3	543.6	551.9	489.7	497.9	.0169	1.0889	1.1057	
74	40.00	3.977	.2514	9.1	542.9	552.0	489.0	498.0	.0185	1.0860	1.1044	
75	40.66	3.927	.2546	9.9	542.2	552.1	488.3	498.1	.0200	1.0832	1.1032	
76	41.32	3.879	0.2578	10.7	541.5	552.2	487.6	498.2	0.0215	1.0804	1.1020	
77	41.97	3.832	.2610	11.5	540.9	552.4	486.9	498.3	.0230	1.0777	1.1008	
78	42.67	3.785	.2642	12.2	540.3	552.5	486.2	498.3	.0245	1.0750	1.0996	
79	43.28	3.740	.2674	12.9	539.7	552.6	485.5	498.4	.0260	1.0723	1.0984	
80	43.88	3.696	.2706	13.6	539.1	552.7	484.8	498.4	.0275	1.0706	1.0972	
81	44.54	3.653	0.2737	14.3	538.5	553.8	484.2	498.5	0.0289	1.0671	1.0917	
82	45.17	3.610	.2770	15.0	537.9	552.9	483.5	498.5	.0303	1.0646	1.0950	
83	45.80	3.568	.2803	15.7	537.3	553.0	482.9	498.6	.0317	1.0621	1.0949	
84	46.41	3.527	.2835	16.4	536.6	553.0	482.2	498.6	.0331	1.0596	1.0928	
85	47.02	3.485	.2868	17.2	535.9	553.1	481.5	498.7	.0345	1.0572	1.0917	
86	47.63	3.446	0.2902	17.9	535.3	553.2	480.9	498.8	0.0359	1.0547	1.0906	
87	48.23	3.407	.2935	18.5	534.8	553.3	480.3	498.8	.0372	1.0523	1.0895	
88	48.82	3.370	.2967	19.2	534.2	553.4	479.8	499.0	.0385	1.0499	1.0884	
89	49.39	3.335	.2999	20.0	533.5	553.5	479.0	499.0	.0398	1.0475	1.0873	
90	49.97	3.300	.3030	20.7	532.8	553.5	478.4	499.1	.0411	1.0462	1.0863	

PRESSURE TABLE.—Continued

TABLE II

Pressure, lb.	Temp., °F.	t	Sp. vol., cu.in. per lb.	ψ_1	Density, lb. per cu. ft.	$\frac{1}{v_1}$	Heat content of liquid	Latent heat of liquid	Heat content of vapor	Q ₁	Internal energy, B.t.u.			Entropy ϕ_1
											Evap.	U_1	Vapor U_1	Liquid ϕ_1
91	50.55	3.265	0.3663	21.4	532.2	553.6	477.7	499.1	0.0424	1.0429	1.0853	1.0406	1.0660	1.0660
92	51.13	3.231	0.3695	21.9	531.8	553.7	477.3	499.2	0.0437	1.0437	1.0843	1.0450	1.0659	1.0659
93	51.70	3.199	0.3126	22.6	531.2	553.8	476.6	499.2	0.0450	1.0450	1.0833	1.0463	1.0658	1.0658
94	52.27	3.166	0.3159	23.3	530.5	553.8	476.0	499.3	0.0463	1.0463	1.0823	1.0470	1.0657	1.0657
95	52.83	3.134	0.3191	23.9	530.0	553.9	475.5	499.4	0.0470	1.0470	1.0813	1.0473	1.0656	1.0656
96	53.39	3.102	0.3223	24.5	529.4	553.9	474.9	499.4	0.0489	1.0489	1.0803	1.0494	1.0655	1.0655
97	53.95	3.070	0.3257	25.3	528.7	554.0	474.2	499.5	0.0501	1.0495	1.0794	1.0501	1.0654	1.0654
98	54.50	3.039	0.3291	25.9	528.1	554.0	473.6	499.5	0.0513	1.0513	1.0785	1.0513	1.0653	1.0653
99	55.04	3.009	0.3323	26.4	527.7	554.1	473.1	499.5	0.0525	1.0525	1.0776	1.0525	1.0652	1.0652
100	55.58	2.980	0.3356	27.1	527.0	554.1	472.4	499.5	0.0537	1.0537	1.0767	1.0537	1.0651	1.0651
101	56.10	2.952	0.3388	27.6	526.6	554.2	471.9	499.5	0.0549	1.0549	1.0758	1.0549	1.0650	1.0650
102	56.62	2.924	0.3420	28.2	526.0	554.2	471.4	499.6	0.0561	1.0561	1.0749	1.0561	1.0649	1.0649
103	57.13	2.897	0.3452	28.8	525.5	554.3	470.8	499.6	0.0573	1.0573	1.0740	1.0573	1.0648	1.0648
104	57.63	2.870	0.3484	29.4	525.0	554.4	470.1	499.7	0.0584	1.0584	1.0731	1.0584	1.0647	1.0647
105	58.14	2.844	0.3516	30.0	524.4	554.4	469.7	499.7	0.0596	1.0596	1.0722	1.0596	1.0646	1.0646
106	58.65	2.817	0.3550	30.6	523.9	554.5	469.2	499.8	0.0607	1.0607	1.0713	1.0607	1.0645	1.0645
107	59.15	2.791	0.3583	31.1	523.4	554.5	468.7	499.8	0.0618	1.0618	1.0704	1.0618	1.0644	1.0644
108	59.66	2.765	0.3617	31.8	522.8	554.6	468.1	499.9	0.0629	1.0629	1.0695	1.0629	1.0643	1.0643
109	60.16	2.740	0.3650	32.4	522.2	554.6	467.5	499.9	0.0640	1.0640	1.0686	1.0640	1.0642	1.0642
110	60.67	2.716	0.3682	33.0	521.7	554.7	466.9	499.9	0.0651	1.0651	1.0677	1.0651	1.0641	1.0641
111	61.18	2.692	0.3715	33.5	521.2	554.7	466.4	499.9	0.0662	1.0662	1.0666	1.0662	1.0640	1.0640
112	61.68	2.669	0.3747	34.1	520.7	554.8	465.8	499.9	0.0673	1.0673	1.0660	1.0673	1.0639	1.0639
113	62.19	2.646	0.3779	34.7	520.1	554.8	465.3	500.0	0.0684	1.0684	1.0649	1.0684	1.0638	1.0638
114	62.69	2.624	0.3811	35.3	519.5	554.8	464.7	500.0	0.0695	1.0695	1.0640	1.0695	1.0637	1.0637
115	63.19	2.602	0.3843	35.8	519.1	554.9	464.2	500.0	0.0706	1.0706				
116	63.61	2.580	0.3876	36.3	518.6	554.9	463.7	500.0	0.0717	1.0717	1.0624	1.0717	1.0636	1.0636
117	64.14	2.559	0.3908	36.8	518.1	554.9	463.3	500.1	0.0728	1.0728	1.0616	1.0728	1.0635	1.0635
118	64.59	2.538	0.3940	37.4	517.6	555.0	462.7	500.1	0.0738	1.0738	1.0608	1.0738	1.0634	1.0634
119	65.02	2.517	0.3973	37.9	517.1	555.0	462.3	500.2	0.0748	1.0748	1.0600	1.0748	1.0633	1.0633
120	65.49	2.497	0.4005	38.5	516.6	555.1	461.7	500.2	0.0758	1.0758	1.0592	1.0758	1.0632	1.0632

PRESSURE TABLE.—Continued

Pressure, lb.	Temp., °F.	Sp. vol., cu. ft. per lb.	$\frac{v}{v_1}$	Density, lb. per cu. ft.	Heat content of liquid Q_1	Latent heat of liquid L	Heat content of vapor Q_2	Heat content of vapor U_1	Internal energy, B.t.u.		Entropy $\frac{L}{T}$	Vapor ϕ_1
									Evap. U_2	Vapor U_1		
121	65.94	2.477	0.4037	39.0	516.1	555.1	461.3	500.3	.0768	.0818	1.0535	
122	66.41	2.457	.4970	39.5	515.6	555.1	460.8	500.3	.0778	.9800	1.0577	
123	66.85	2.437	.4103	40.0	515.1	555.1	460.4	500.4	.0788	.9782	1.0570	
124	67.33	2.418	.4136	40.5	514.6	555.1	459.9	500.4	.0793	.9765	1.0563	
125	67.81	2.400	.4167	41.1	514.1	555.2	459.4	500.5	.0808	.9747	1.0555	
126	68.27	2.382	0.4198	41.7	513.5	555.2	458.8	500.5	.0818	.9729	1.0547	
127	68.7	2.363	.4232	42.2	513.1	555.3	458.4	500.6	.0828	.9711	1.0539	
128	69.2	2.345	.4266	42.8	512.5	555.3	457.8	500.6	.0838	.9693	1.0531	
129	69.6	2.327	.4297	43.3	512.0	555.3	457.3	500.6	.0848	.9676	1.0524	
130	70.1	2.309	.4330	43.9	511.4	555.3	456.8	500.7	.0858	.9659	1.0517	
131	70.5	2.292	0.4363	44.3	511.1	555.4	456.4	500.7	.0868	.9642	1.0510	
132	70.9	2.275	.4396	44.8	510.6	555.4	455.9	500.7	.0878	.9625	1.0503	
133	71.3	2.259	.4427	45.2	510.2	555.4	455.5	500.7	.0887	.9608	1.0495	
134	71.7	2.243	.4458	45.7	509.7	555.4	455.1	500.8	.0896	.9591	1.0487	
135	72.2	2.227	.4490	46.3	509.2	555.5	454.5	500.8	.0905	.9575	1.0480	
136	72.6	2.201	0.4543	46.8	508.7	555.5	454.0	500.8	.0914	.9559	1.0473	
137	73.0	2.196	.4554	47.2	508.3	555.5	453.6	500.8	.0923	.9543	1.0466	
138	73.4	2.180	.4587	47.7	507.8	555.5	453.1	500.8	.0932	.9527	1.0459	
139	73.8	2.164	.4621	48.2	507.4	555.6	452.7	500.9	.0941	.9510	1.0451	
140	74.2	2.149	.4653	48.7	506.9	555.6	452.2	500.9	.0950	.9497	1.0444	
141	74.7	2.134	0.4686	49.2	506.4	555.6	451.7	500.9	.0959	.9478	1.0437	
142	75.1	2.119	.4719	49.6	506.0	555.6	451.2	500.9	.0968	.9462	1.0430	
143	75.5	2.105	.4751	50.1	505.5	555.7	450.8	501.0	.0977	.9447	1.0424	
144	75.9	2.090	.4785	50.6	505.1	555.7	450.4	501.0	.0986	.9432	1.0418	
145	76.3	2.076	.4817	51.1	504.7	555.7	449.9	501.0	.0995	.9417	1.0412	
146	76.7	2.063	0.4847	51.6	504.2	555.8	449.4	501.0	.1004	.9402	1.0406	
147	77.1	2.049	.4880	52.0	503.8	555.8	449.0	501.0	.1013	.9387	1.0400	
148	77.5	2.035	.4914	52.5	503.3	555.8	448.6	501.1	.1022	.9372	1.0394	
149	77.9	2.021	.4948	53.0	502.9	555.9	448.1	501.1	.1030	.9357	1.0387	
150	78.3	2.008	.4980	53.4	502.5	555.9	447.7	501.1	.1039	.9342	1.0381	

TABLE II

PRESSURE TABLE.—Continued

Pressure, lb.	Temp., °F.	t	Sp. vol., cu.it. per lb.	$\frac{1}{v_1}$	Density, lb. per cu. ft.	Heat content of liquid	Latent heat of liquid	Heat content of vapor	Q ₁	Internal energy, B.t.u.		Entropy $\frac{L}{T}$	Vapor ϕ_1
										Evap.	U ₂		
151	78.7	1.995	1.982	0.5013	53.8	502.1	555.9	447.3	501.1	0.1048	0.9326	1.0374	
152	79.1	1.982	1.969	.5045	54.3	501.6	555.9	446.8	501.1	.1056	.9311	1.0367	
153	79.4	1.969	1.957	.5079	54.8	501.2	556.0	446.4	501.2	.1065	.9306	1.0361	
154	79.8	1.957	1.944	.5110	55.2	500.8	556.0	446.0	501.2	.1074	.9281	1.0355	
155	80.2	1.944	1.931	.5144	55.7	500.3	556.0	445.5	501.2	.1082	.9266	1.0348	
156	80.6	1.931	1.919	0.5179	56.1	499.9	556.0	445.1	501.2	0.1090	0.9252	1.0342	
157	81.0	1.919	1.907	.5211	56.6	499.5	556.1	444.7	501.3	.1098	.9238	1.0336	
158	81.3	1.907	1.895	.5244	57.0	499.1	556.1	444.3	501.3	.1106	.9224	1.0330	
159	81.7	1.895	1.885	.5277	57.4	498.7	556.1	443.9	501.3	.1114	.9210	1.0324	
160	82.1	1.885	1.873	.5305	57.8	498.3	556.1	443.5	501.3	.1122	.9199	1.0318	
161	82.5	1.873	1.861	0.5339	58.3	497.9	556.2	443.0	501.3	0.1130	0.9182	1.0312	
162	82.9	1.861	1.849	.5373	58.8	497.4	556.2	442.6	501.4	.1138	.9168	1.0306	
163	83.3	1.849	1.838	.5408	59.2	497.0	556.2	442.2	501.4	.1146	.9154	1.0300	
164	83.7	1.838	1.828	.5441	59.6	496.6	556.2	441.8	501.4	.1154	.9140	1.0294	
165	84.0	1.828	1.817	.5470	60.1	496.2	556.3	441.3	501.4	.1162	.9126	1.0288	
166	84.4	1.817	1.806	0.5504	60.5	495.8	556.3	440.9	501.4	0.1170	0.9113	1.0283	
167	84.8	1.806	1.795	.5537	60.9	495.4	556.3	440.5	501.4	.1178	.9099	1.0277	
168	85.1	1.795	1.784	.5568	61.3	495.0	556.3	440.1	501.4	.1186	.9085	1.0271	
169	85.4	1.784	1.773	.5599	61.7	494.6	556.4	439.7	501.4	.1194	.9071	1.0265	
170	85.8	1.773	1.762	.5631	62.2	494.2	556.4	439.3	501.5	.1201	.9062	1.0259	
171	86.2	1.766	1.755	0.5663	62.7	493.7	556.4	438.8	501.5	0.1208	0.9045	1.0253	
172	86.5	1.756	1.746	.5695	63.1	493.3	556.4	438.4	501.5	.1216	.9032	1.0248	
173	86.9	1.746	1.736	.5727	63.5	492.9	556.4	438.0	501.5	.1224	.9019	1.0243	
174	87.2	1.736	1.726	.5760	63.9	492.5	556.4	437.6	501.5	.1231	.9006	1.0237	
175	87.6	1.726	1.716	.5794	64.3	492.1	556.4	437.2	501.5	.1239	.8992	1.0231	
176	87.9	1.716	1.706	0.5828	64.7	491.8	556.5	436.8	501.5	0.1247	0.8979	1.0226	
177	88.3	1.706	1.696	.5862	65.1	491.4	556.5	436.4	501.5	.1255	.8965	1.0220	
178	88.6	1.696	1.687	.5896	65.5	491.0	556.5	436.0	501.6	.1263	.8952	1.0215	
179	89.0	1.687	1.677	.5928	65.9	490.6	556.5	435.6	501.6	.1270	.8939	1.0209	
180	89.3	1.677	1.667	.5963	66.3	490.2	556.5	435.3	501.6	.1277	.8920	1.0204	

PRESSURE TABLE.—Continued

Pressure, lb.	Temp., °F.	Sp. vol., cu. ft. per lb.	v_1	Density, lb. per cu. ft.	$\frac{1}{v_1}$	Heat content of liquid	Latent heat of liquid	L	Q_1	Heat content of vapor	Internal energy, B.t.u.	Entropy		
												Evap. U_2	Vapor U_1	Liquid ϕ_2
181	89.7	1.668	0.5905	66.7	0.5905	489.8	556.5	435.9	501.6	0.1284	0.8914	1.0198		
182	90.0	1.659	0.6028	67.1	0.6028	489.4	556.5	434.5	501.6	.1291	.8902	1.0193		
183	90.4	1.650	0.6061	67.5	0.6061	489.0	556.5	434.1	501.6	.1298	.8889	1.0187		
184	90.7	1.641	0.6094	67.9	0.6094	488.6	556.5	433.7	501.6	.1205	.8877	1.0182		
185	91.0	1.633	0.6124	68.3	0.6124	488.2	556.5	433.3	501.6	.1312	.8865	1.0177		
186	91.4	1.624	0.6158	68.7	0.6158	487.8	556.5	433.0	501.7	0.1319	0.8852	1.0171		
187	91.7	1.615	0.6192	69.1	0.6192	487.4	556.5	432.6	501.7	.1327	.8839	1.0166		
188	92.0	1.606	0.6227	69.5	0.6227	487.0	556.5	432.2	501.7	.1334	.8827	1.0161		
189	92.4	1.597	0.6262	69.9	0.6262	486.6	556.5	431.8	501.7	.1341	.8814	1.0155		
190	92.7	1.589	0.6293	70.3	0.6293	486.2	556.5	431.4	501.7	.1348	.8804	1.0149		
191	93.0	1.581	0.6325	70.7	0.6325	485.9	556.6	431.0	501.7	0.1354	0.8789	1.0143		
192	93.4	1.573	0.6357	71.1	0.6357	485.5	556.6	430.6	501.7	.1361	.8777	1.0138		
193	93.7	1.565	0.6390	71.5	0.6390	485.1	556.6	430.2	501.7	.1368	.8765	1.0133		
194	94.0	1.557	0.6423	71.9	0.6423	484.7	556.6	429.8	501.7	.1374	.8753	1.0127		
195	94.3	1.549	0.6456	72.3	0.6456	484.3	556.6	429.4	501.7	.1381	.8741	1.0122		
196	94.6	1.541	0.6489	72.7	0.6489	483.9	556.6	429.0	501.7	0.1388	0.8729	1.0117		
197	94.9	1.533	0.6523	73.1	0.6523	483.5	556.6	428.6	501.7	.1394	.8718	1.0112		
198	95.2	1.525	0.6557	73.5	0.6557	483.1	556.6	428.3	501.8	.1401	.8706	1.0107		
199	95.5	1.517	0.6592	73.8	0.6592	482.8	556.6	428.0	501.8	.1408	.8694	1.0102		
200	95.9	1.510	0.6623	74.1	0.6623	482.5	556.6	427.7	501.8	.1414	.8685	1.0097		
202	96.5	1.495	0.6669	74.9	0.6669	481.7	556.6	426.9	501.8	.1428	.8659	1.0087		
204	97.1	1.480	0.6757	75.6	0.6757	481.0	556.6	426.2	501.8	.1441	.8636	1.0077		
206	97.8	1.466	0.6821	76.3	0.6821	480.3	556.6	425.5	501.8	.1454	.8613	1.0067		
208	98.4	1.452	0.6887	77.0	0.6887	479.6	556.6	424.8	501.8	.1467	.8591	1.0058		
210	99.0	1.438	0.6954	77.7	0.6954	478.8	556.5	424.1	501.8	.1479	.8569	1.0048		

TABLE II

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PRESSURE TABLE.—Continued

Pressure, lb.	Temp., °F.	Sp. vol., cu. ft. per lb.	V ₁	Density, lb. per cu. ft.	$\frac{1}{V_1}$	Heat content of liquid	Latent heat of liquid	Heat content of vapor	Internal energy, B.t.u.		Entropy $\frac{L}{T}$	Vapor ϕ_1
									Q ₁	L		
212	99.6	1.424	0.7022	78.4	478.1	556.5	423.4	501.8	0.1492	0.8547	1.0039	
214	100.2	1.410	.7092	79.1	477.4	556.5	422.7	501.8	.1505	.8225	1.0030	
216	100.8	1.397	.7158	79.8	476.7	556.5	422.0	501.8	.1517	.8504	1.0021	
218	101.4	1.384	.7225	80.5	476.0	556.5	421.3	501.8	.1529	.8283	1.0012	
220	102.0	1.371	.7294	81.2	475.3	556.5	420.6	501.8	.1541	.8462	1.0003	
222	102.6	1.358	0.7364	81.9	474.6	556.5	419.9	501.8	0.1554	0.8440	0.9994	
224	103.2	1.346	.7429	82.6	473.9	556.5	419.2	501.8	.1566	.8419	.9985	
226	103.7	1.334	.7496	83.3	473.2	556.5	418.5	501.8	.1578	.8398	.9976	
228	104.3	1.323	.7559	84.0	472.5	556.5	417.8	501.8	.1590	.8377	.9967	
230	104.9	1.312	.7622	84.6	471.8	556.4	417.2	501.8	.1602	.8357	.9959	
232	105.5	1.300	0.7692	85.3	471.1	556.4	416.5	501.8	0.1614	0.8336	0.9950	
234	106.0	1.289	.7758	85.9	470.5	556.4	415.9	501.8	.1626	.8316	.9942	
236	106.6	1.278	.7825	86.6	469.8	556.4	415.2	501.8	.1637	.8296	.9933	
238	107.2	1.267	.7893	87.2	469.1	556.3	414.6	501.8	.1649	.8276	.9925	
240	107.7	1.256	.7962	87.9	468.4	556.3	413.9	501.8	.1660	.8256	.9916	
242	108.2	1.245	0.8032	88.5	467.7	556.2	413.2	501.7	0.1672	0.8336	0.9908	
244	108.8	1.234	.8104	89.2	467.0	556.2	412.5	501.7	.1683	.8316	.9899	
246	109.3	1.224	.8170	89.8	466.3	556.1	411.9	501.7	.1695	.8196	.9891	
248	109.9	1.214	.8237	90.5	465.6	556.1	411.2	501.7	.1706	.8176	.9882	
250	110.4	1.204	.8306	91.1	464.9	556.0	410.6	501.7	.1717	.8157	.9874	
252	111.0	1.195	0.8368	91.7	464.3	556.0	410.0	501.7	0.1728	0.8138	0.9866	
254	111.5	1.185	.8439	92.3	463.7	555.0	409.4	501.7	.1739	.8119	.9858	
256	112.0	1.176	.8503	92.9	463.0	555.9	408.8	501.7	.1750	.8100	.9850	
258	112.5	1.167	.8569	93.6	462.3	555.9	408.1	501.7	.1761	.8081	.9842	
260	113.1	1.158	.8636	94.2	461.7	555.9	407.5	501.7	.1771	.8063	.9834	

PRESSURE TABLE.—Continued

Pressure, lb. p	Temp., °F. t	Sp. vol., cu. ft. per lb. v ₁	Density, lb. per cu. ft. $\frac{1}{v_1}$	Heat content of liquid Q ₁	Latent heat of liquid L	Heat content of vapor Q ₁	Internal energy, B.t.u.		Entropy		
							Evap. U ₁	Vapor U ₁	Liquid ϕ ₁	Evap. L $\frac{L}{T}$	Vapor ϕ ₁
262	113.6	1.149	0.8703	94.8	461.0	555.8	406.9	501.7	0.1782	0.8044	0.9826
264	114.1	1.140	.8772	95.4	460.4	555.8	406.2	501.6	.1792	.8026	.9818
266	114.6	1.131	.8842	96.0	459.7	555.7	405.6	501.6	.1803	.8007	.9810
268	115.1	1.122	.8913	96.6	459.1	555.7	405.0	501.6	.1813	.7989	.9802
270	115.6	1.113	.8985	97.2	458.5	555.7	404.4	501.6	.1824	.7971	.9795
272	116.1	1.104	0.9038	97.8	457.8	555.6	403.8	501.6	0.1834	0.7953	0.9787
274	116.6	1.096	.9124	98.4	457.2	555.6	403.1	501.5	.1845	.7935	.9780
276	117.1	1.088	.9191	99.0	456.5	555.5	402.5	501.5	.1855	.7917	.9772
278	117.6	1.080	.9259	99.6	455.9	555.5	401.9	501.5	.1866	.7899	.9765
280	118.1	1.072	.9328	100.2	455.3	555.5	401.2	501.4	.1876	.7881	.9757
282	118.6	1.064	0.9398	100.8	454.6	555.4	400.6	501.4	0.1887	0.7863	0.9750
284	119.1	1.056	.9470	101.3	454.0	555.3	400.0	501.3	.1897	.7846	.9743
286	119.5	1.049	.9533	101.9	453.4	555.3	399.4	501.3	.1906	.7829	.9735
288	120.0	1.043	.9606	102.4	452.9	555.3	398.8	501.2	.1916	.7812	.9728
290	120.5	1.034	.9671	103.0	452.2	555.2	398.2	501.2	.1926	.7795	.9721
292	121.0	1.026	0.9747	103.6	451.6	555.2	397.6	501.2	0.1936	0.7778	0.9714
294	121.4	1.019	.9814	104.1	451.0	555.1	397.1	501.2	.1946	.7761	.9707
296	121.9	1.012	.9881	104.7	450.4	555.1	396.5	501.2	.1956	.7744	.9700
298	122.4	1.005	.9930	105.3	449.8	555.1	395.9	501.2	.1966	.7727	.9693
300	122.9	0.998	1.002	105.9	449.2	555.0	395.3	501.2	.1976	.7710	.9686
310	125.2	0.965	1.036	108.6	446.2	554.8	392.5	501.1	0.2023	0.7626	0.9649
320	127.4	.932	1.073	111.3	443.2	554.5	389.6	500.9	.2069	.7545	.9614
330	129.6	.904	1.106	114.1	440.3	554.4	386.8	500.9	.2115	.7468	.9583
340	131.8	.876	1.142	116.7	437.4	554.1	384.1	500.8	.2160	.7390	.9550
350	133.9	.849	1.173	119.2	434.6	553.8	381.4	500.6	.2203	.7316	.9519

TABLE II

PRESSURE TABLE.—Concluded

Pressure, lb.	Temp., °F.	t	Sp. vol., cu. ft. per lb.	Heat content of liquid per cu. ft.	$\frac{1}{v_1}$	Q ₁	Latent heat of liquid	Heat content of vapor	Q ₁	Internal energy, B.t.u.		Entropy
										Evap. U ₁	Vapor U ₁	
p			v ₁							U	\bar{L}	\bar{T}
360	135.9		.824	1.214	121.6	431.8	553.4	378.8	500.4	0.2245	0.7244	0.9489
370	137.9		.800	1.250	124.1	429.0	553.1	376.1	500.2	.2285	.7175	.9400
380	139.9		.777	1.287	126.5	426.3	552.8	373.5	500.0	.2336	.7106	.9432
390	141.8		.756	1.323	128.8	423.6	552.4	371.0	499.8	.2386	.7038	.9404
400	143.7		.736	1.359	131.1	420.9	552.0	368.5	499.6	.2404	.6972	.9376
410	145.6		.717	1.395	133.5	418.2	551.7	365.9	499.4	0.2441	0.6908	0.9349
420	147.4		.698	1.433	135.7	415.6	551.3	363.4	499.1	.2478	.6844	.9322
430	149.2		.680	1.471	137.9	413.0	550.9	361.9	498.9	.2515	.6780	.9295
440	151.0		.662	1.511	140.1	410.4	550.5	358.6	498.7	.2551	.6718	.9269
450	152.7		.646	1.548	142.2	407.9	550.1	356.3	498.5	.2587	.6657	.9244
460	154.4		.631	1.585	144.2	405.4	549.6	354.0	498.2	0.2622	0.6598	0.9220
470	156.1		.617	1.621	146.3	402.9	549.2	351.7	498.0	.2657	.6540	.9197
480	157.8		.603	1.658	148.4	400.4	548.8	349.3	497.7	.2691	.6482	.9173
490	159.4		.589	1.698	150.4	397.9	548.3	347.0	497.4	.2726	.6424	.9150
500	160.9		.576	1.736	152.5	395.4	547.9	344.6	497.1	.2759	.6368	.9127
525	164.8		.545	1.835	157.5	389.2	546.7	339.0	496.5	0.2838	0.6232	0.9070
550	168.6		.517	1.934	162.4	383.2	545.6	333.5	495.9	.2916	.6099	.9015
575	172.3		.492	2.033	167.2	377.3	544.5	328.1	495.3	.2993	.5970	.8963
600	175.8		.467	2.141	171.9	371.5	543.4	322.8	494.7	.3068	.5845	.8913
625	179.2		.447	2.237	176.6	365.7	542.3	317.6	494.2	0.3141	0.5724	0.8865
650	182.5		.427	2.342	181.3	359.9	541.2	312.4	493.7	.3213	.5605	.8818
675	185.7		.408	2.451	185.9	354.2	540.1	307.3	493.2	.3283	.5489	.8772
700	188.8		.391	2.558	190.5	348.6	539.1	302.2	492.7	.3350	.5378	.8728

SUPERHEATED AMMONIA (Nos. 5 to 12)

	Liquid	Vapor	10°	20°	30°	40°	50°	60°	70°	80°	90°	100°
5 Temp.	-62.13	-42.13	-32.13	-22.13	-12.13	-2.13	+7.87	+17.87	+27.87	+37.87	+37.87	
Q Vol.	527.5	532.5	537.4	542.4	547.3	552.3	557.3	562.4	567.4	572.5	577.5	
φ	0.02260	0.02260	0.02260	0.02260	0.02260	0.02260	0.02260	0.02260	0.02260	0.02260	0.02260	
6 Temp.	-56.84	-46.84	-36.84	-26.84	-16.84	-6.84	+3.16	+13.16	+26.16	+36.16	+46.16	
Q Vol.	529.0	534.0	539.0	543.9	548.9	553.9	558.9	563.9	569.0	574.0	579.0	
φ	1.35333	1.35333	1.35333	1.35333	1.35333	1.35333	1.35333	1.35333	1.35333	1.35333	1.35333	
7 Temp.	-52.23	-42.23	-32.23	-22.23	-12.23	-2.23	+7.77	+17.77	+27.77	+37.77	+47.77	
Q Vol.	530.3	535.3	540.3	545.3	550.3	555.3	560.3	565.3	570.4	575.4	580.4	
φ	0.02286	0.02286	0.02286	0.02286	0.02286	0.02286	0.02286	0.02286	0.02286	0.02286	0.02286	
8 Temp.	-48.13	-38.13	-28.13	-18.13	-8.13	+1.87	+11.87	+21.87	+31.87	+41.87	+51.87	
Q Vol.	531.4	536.4	541.4	546.4	551.4	556.4	561.4	566.4	571.5	576.5	581.5	
φ	32.05	32.05	32.05	32.05	32.05	32.05	32.05	32.05	32.05	32.05	32.05	
9 Temp.	-44.40	-34.40	-24.40	-14.40	-4.40	+5.60	+15.60	+25.60	+35.60	+45.60	+55.60	
Q Vol.	532.6	537.6	542.6	547.7	552.7	557.7	562.7	567.7	572.7	577.8	582.8	
φ	28.75	29.131	29.872	30.610	31.349	31.349	32.081	32.814	33.545	34.276	35.005	
10 Temp.	-40.01	-30.01	-20.01	-10.01	-0.01	+9.19	+19.19	+29.19	+39.19	+49.19	+59.19	
Q Vol.	533.8	538.8	543.8	548.9	553.9	558.9	563.9	569.0	574.0	579.1	584.1	
φ	26.05	26.647	27.203	27.930	28.581	29.210	29.109	29.467	30.094	31.719	32.345	
11 Temp.	-37.76	-27.76	-17.76	-7.76	+2.24	+12.24	+22.24	+32.24	+42.24	+52.24	+62.24	
Q Vol.	534.8	539.8	544.8	549.9	554.9	559.9	564.9	570.0	575.0	580.1	585.1	
φ	23.84	24.416	25.012	25.597	26.181	26.764	27.307	27.927	28.507	29.086	29.664	
12 Temp.	-34.83	-24.83	-14.83	-4.83	+5.17	+15.17	+25.17	+35.17	+45.17	+55.17	+65.17	
Q Vol.	535.8	538.8	545.8	550.9	555.9	560.9	566.0	571.0	576.1	581.1	586.2	
φ	21.99	22.53	23.06	23.60	24.13	24.67	25.30	25.73	26.27	26.80	27.33	
	1.2742	1.2865	1.2902	1.3097	1.3212	1.3342	1.3431	1.3540	1.3643	1.3747	1.3850	

SUPERHEATED AMMONIA (Nos. 5 to 12).—Continued

TABLE III

	Liquid	Vapor	110°	120°	130°	140°	150°	160°	180°	200°	250°	300°
5	Temp.	—62.13	+47.87	+57.87	+67.87	+77.87	+87.87	+97.87	+117.87	+137.87	+187.87	+237.87
Q	Vol.	-106.2	527.5	582.6	587.6	597.7	597.7	602.8	607.9	618.04	628.2	653.8
φ	—	0.02260	49.72	63.73	64.060	66.35	67.499	68.764	69.987	72.526	75.064	81.402
	Temp.	—56.84	+56.16	+66.16	+76.16	+86.16	+96.16	+106.16	+126.16	+146.16	+196.16	+246.16
6	Q	-100.4	529.0	584.1	589.1	594.2	599.2	604.8	609.4	619.6	629.8	655.4
Vol.	—	0.02273	41.79	52.547	53.512	54.476	55.438	56.400	57.549	59.844	62.138	67.866
φ	—	.2261	1.3400	1.4018	1.4718	1.4815	1.4910	1.5002	1.5098	1.5274	1.5450	1.5860
7	Temp.	—52.23	+57.77	+67.77	+77.77	+87.77	+97.77	+107.77	+127.77	+147.77	+197.77	+247.77
Q	Vol.	-95.3	530.3	589.5	595.6	599.5	604.8	608.0	610.8	621.0	631.42	656.8
φ	—	0.02286	36.23	46.210	47.179	48.087	48.994	49.900	50.811	52.642	54.449	58.903
8	Temp.	—48.13	+61.87	+71.87	+81.87	+91.87	+101.87	+111.87	+131.87	+151.87	+201.87	+251.87
Q	Vol.	-90.8	531.4	586.6	591.7	596.7	601.8	606.9	612.0	622.2	632.4	658.0
φ	—	0.02298	32.05	40.839	41.634	42.446	43.218	44.01	44.812	46.413	48.012	51.968
9	Temp.	—44.40	+65.60	+75.60	+85.60	+95.60	+105.60	+115.60	+135.60	+155.60	+205.60	+255.60
Q	Vol.	-86.5	532.6	587.9	592.9	598.0	603.0	608.1	613.2	623.5	633.7	659.4
φ	—	0.02308	28.75	36.460	37.186	37.890	38.585	39.281	40.055	41.449	42.803	46.368
10	Temp.	—40.91	+60.19	+70.19	+80.19	+90.19	+100.19	+110.19	+130.19	+150.19	+200.19	+250.19
Q	Vol.	-82.9	533.8	589.2	594.2	598.9	603.9	609.4	614.5	624.8	635.0	660.7
φ	—	0.02318	26.05	33.060	33.603	34.286	34.938	35.560	36.455	37.093	38.929	42.015
11	Temp.	—37.76	+72.24	+82.24	+92.24	+102.24	+112.24	+122.24	+142.24	+162.24	+212.24	+262.24
Q	Vol.	-70.3	534.8	590.2	595.3	600.2	605.4	610.5	615.6	625.9	636.1	661.8
φ	—	0.02326	23.84	30.239	30.817	31.4335	31.4327	31.569	32.541	33.233	34.376	38.388
12	Temp.	—34.83	+75.17	+85.17	+95.17	+105.17	+115.17	+125.17	+145.17	+165.17	+215.17	+265.17
Q	Vol.	-76.0	535.8	591.3	596.4	601.4	606.5	611.6	616.7	627.0	637.2	662.9
φ	—	0.02333	21.99	27.86	28.39	28.93	29.46	29.99	30.53	31.00	32.67	35.35
		-106.5	1.2742	1.3949	1.4047	1.4144	1.4241	1.4324	1.4422	1.4600	1.4774	1.5182

SUPERHEATED AMMONIA (Nos. 13 to 20)

	Liquid	Vapor	10°	20°	30°	40°	50°	60°	70°	80°	90°	100°	
13 Temp. Q Vol. ϕ	-32.10 -73.1 0.02340 - .1592	536.6 541.6 20.40 1.2667	-22.10 -12.10 20.90 1.2790	-12.10 551.7 21.39 1.2907	-2.10 550.7 21.88 1.3021	+ 7.90 561.7 22.38 1.3137	+17.90 566.8 22.88 1.3246	+27.90 571.9 23.37 1.3355	+37.90 576.9 23.87 1.3462	+47.90 582.0 24.36 1.3567	+57.90 587.1 24.85 1.3670	+67.90 587.1 25.35 1.3773	
14 Temp. Q Vol. ϕ	-29.55 -70.2 0.02345 - .1533	537.5 547.6 19.02 1.2599	-10.55 -9.55 547.6 19.48	+ 0.45 +10.45 557.7 20.41	+10.45 +20.45 567.7 20.87	+30.45 +40.45 567.8 21.78	+40.45 +50.45 577.9 22.25	+50.45 +60.45 577.9 22.71	+60.45 +70.45 582.9 23.17	+70.45 588.0 23.63 1.3703	+70.45 588.0 23.63 1.3703		
15 Temp. Q Vol. ϕ	-27.16 -67.5 0.02351 - .1438	538.5 543.5 17.83 1.2537	-17.16 -7.16 548.6 18.28	+ 2.84 +12.84 553.6 19.13	+12.84 +22.84 558.7 19.55	+32.84 +42.84 563.7 19.98	+32.84 +42.84 568.8 20.41	+42.84 +52.84 573.9 20.85	+62.84 +72.84 578.9 21.28	+62.84 +72.84 584.0 21.72	+62.84 +72.84 584.0 21.72	+62.84 +72.84 584.0 21.72	
16 Temp. Q Vol. ϕ	-24.87 -65.0 0.02356 - .1399	538.7 543.7 16.78 1.2473	-14.87 -4.87 548.8 17.60	+ 5.13 +15.13 553.9 18.01	+15.13 +25.13 558.9 18.41	+35.13 +45.13 560.9 18.82	+35.13 +45.13 560.9 19.23	+35.13 +45.13 560.9 19.63	+55.13 +65.13 570.1 20.04	+55.13 +65.13 570.1 20.45	+55.13 +65.13 570.1 20.85	+55.13 +65.13 570.1 20.85	
17 Temp. Q Vol. ϕ	-22.69 -62.4 0.02361 - .1343	539.5 544.6 15.85 1.2416	-12.69 -12.69 549.6 16.25	- 2.69 + 7.31 554.7 16.64	+ 7.31 +17.31 559.7 17.02	+27.31 +37.31 564.8 17.41	+27.31 +37.31 569.9 17.79	+47.31 +57.31 575.0 18.17	+57.31 +67.31 580.0 18.56	+57.31 +67.31 580.0 18.94	+57.31 +67.31 585.1 19.32	+57.31 +67.31 585.1 19.71	+57.31 +67.31 585.1 19.71
18 Temp. Q Vol. ϕ	-20.61 -60.1 0.02366 - .1299	540.0 545.1 15.015 1.2361	-10.61 - 0.61 550.1 15.74	+ 0.61 + 9.39 555.2 16.11	+ 9.39 +19.39 560.2 16.41	+29.39 +39.39 565.3 16.83	+29.39 +39.39 570.4 17.20	+39.39 +49.39 575.5 17.56	+49.39 +59.39 580.5 17.92	+49.39 +59.39 580.5 18.28	+49.39 +59.39 580.5 18.65	+49.39 +59.39 580.5 18.65	
19 Temp. Q Vol. ϕ	-18.63 -57.9 0.02371 - .1238	540.5 545.6 14.30 1.2311	- 8.63 - 8.63 550.6 14.64	+ 1.37 +11.37 555.7 15.34	+ 1.37 +21.37 560.7 15.67	+31.37 +41.37 565.8 16.01	+31.37 +41.37 570.9 16.36	+51.37 +61.37 576.0 16.70	+61.37 +71.37 581.1 17.05	+61.37 +71.37 581.1 17.38	+81.37 +91.37 586.2 17.72	+81.37 +91.37 586.2 17.72	
20 Temp. Q Vol. ϕ	-16.70 -55.7 0.02376 - .1188	541.0 546.1 13.66 1.2382	- 6.70 - 6.70 551.1 14.399	+ 3.30 +13.30 556.2 14.632	+ 3.30 +23.30 561.2 14.955	+33.30 +43.30 566.4 15.277	+33.30 +43.30 571.5 15.601	+63.30 +73.30 576.60 15.923	+63.30 +73.30 581.7 16.245	+63.30 +73.30 586.8 16.566	+83.30 +91.30 591.9 16.888	+83.30 +91.30 591.9 16.888	

SUPERHEATED AMMONIA (NOS. 13 to 20).—Continued

TABLE III

	Liquid	Vapor	.110°	.120°	.130°	.140°	.150°	.160°	.180°	.200°	.250°	.300°	
13	Temp.	-32.10	+77.90	+ 87.00	+ 97.00	+107.00	+117.00	+127.00	+147.00	+167.00	+217.90	+267.90	
	Q Vol.	-73.1 0.02340	536.6 20.40	592.2 25.84	597.3 26.33	607.4 27.32	612.5 27.80	617.6 28.30	627.9 28.90	638.1 30.27	663.8 32.74	689.8 35.21	
φ		-1.2667	1.3872	1.3970	1.4067	1.4165	1.4252	1.4345	1.4522	1.4697	1.5105	1.5485	
14	Temp.	-29.55	+80.45	+ 90.45	+100.45	+110.45	+120.45	+130.45	+150.45	+170.45	+220.45	+270.45	
	Q Vol.	-70.2 0.02345	537.5 19.02	593.1 24.09	598.2 24.55	603.3 25.01	608.4 25.47	613.5 25.93	618.6 26.39	623.7 26.71	630.1 27.31	664.8 28.22	690.0 30.51
φ		-1.2599	1.3802	1.3900	1.3998	1.4094	1.4182	1.4274	1.4451	1.4627	1.5033	1.5411	
15	Temp.	-27.16	+82.84	+ 92.84	+102.84	+112.84	+122.84	+132.84	+152.84	+172.84	+222.84	+272.84	
	Q Vol.	-67.5 0.02351	538.5 17.83	590.3 22.58	594.2 23.01	600.4 23.45	606.4 23.88	612.5 24.31	619.6 24.74	626.9 25.59	640.2 26.44	665.9 28.57	692.0 30.69
φ		-1.2537	1.3737	1.3837	1.3934	1.4029	1.4118	1.4209	1.4387	1.4562	1.4967	1.5347	
16	Temp.	-24.87	+85.13	+ 95.13	+105.13	+115.13	+125.13	+135.13	+155.13	+175.13	+225.13	+275.13	
	Q Vol.	-65.0 0.02356	538.7 16.78	594.4 21.26	599.5 21.67	604.6 22.07	609.7 22.47	614.8 22.88	619.9 23.28	630.2 24.07	640.4 24.87	666.2 26.86	692.3 28.86
φ		-1.2473	1.3672	1.3772	1.3868	1.3963	1.4052	1.4143	1.4321	1.4495	1.4901	1.5278	
17	Temp.	-22.69	+87.31	+ 97.31	+107.31	+117.31	+127.31	+137.31	+157.31	+177.31	+227.31	+277.31	
	Q Vol.	-62.4 0.02361	539.5 15.85	595.3 20.08	600.4 20.46	605.5 20.84	610.6 21.22	615.7 21.60	620.8 21.96	630.2 22.73	640.4 23.49	666.2 25.36	692.3 27.24
φ		-1.2416	1.3614	1.3714	1.3809	1.3904	1.3993	1.4084	1.4261	1.4435	1.4841	1.5216	
18	Temp.	-20.61	+89.39	+ 99.39	+109.39	+119.39	+129.39	+139.39	+159.39	+179.39	+229.39	+279.39	
	Q Vol.	-60.1 0.02366	540.0 15.05	595.8 19.01	600.9 19.37	606.1 19.73	611.3 20.09	616.3 20.46	621.4 20.81	631.7 21.52	641.9 22.33	667.7 24.00	693.3 25.77
φ		-1.2361	1.3558	1.3657	1.3751	1.3846	1.3936	1.4026	1.4203	1.4378	1.4784	1.5159	
19	Temp.	-18.63	+91.37	+101.37	+111.37	+121.37	+131.37	+141.37	+161.37	+181.37	+231.37	+281.39	
	Q Vol.	-57.9 0.02371	540.5 14.30	596.4 18.07	601.5 18.41	606.6 18.74	611.7 19.08	616.8 19.43	621.9 19.76	632.2 20.44	642.5 21.12	668.4 22.80	694.5 24.48
φ		-1.2311	1.3506	1.3605	1.3709	1.3805	1.3904	1.4003	1.4174	1.4356	1.4731	1.5105	
20	Temp.	-16.70	+93.30	+103.30	+113.30	+123.30	+133.30	+143.30	+163.30	+183.30	+233.30	+283.30	
	Q Vol.	-55.7 0.02376	541.0 13.66	597.0 17.200	602.1 17.529	607.2 17.850	612.3 18.170	617.4 18.49	622.5 18.80	632.8 19.506	643.1 20.183	669.0 21.75	696.1 23.29
φ		-1.2262	1.3454	1.3552	1.3647	1.3743	1.3832	1.3922	1.4099	1.4274	1.4677	1.5052	

SUPERHEATED AMMONIA (Nos. 21 to 28)

	Liquid	Vapor	10°	20°	30°	40°	50°	60°	70°	80°	90°	100°
21	Temp.	-14.84	-4.84	+ 5.16	+15.16	+25.16	+35.16	+45.16	+55.16	+65.16	+75.16	+85.16
Q Vol.	-53.6	541.4	546.6	551.5	559.5	561.6	566.8	571.9	577.0	582.1	587.2	591.3
φ	0.02381	13.025	13.335	13.644	13.955	14.275	14.573	14.882	15.191	15.499	15.807	16.114
22	Temp.	-13.07	-3.07	+ 6.93	+16.93	+26.93	+36.93	+46.93	+56.93	+66.93	+76.93	+86.93
Q Vol.	-51.6	541.7	546.8	551.8	556.9	561.9	567.1	572.2	577.3	582.4	587.5	592.6
φ	0.02386	12.470	12.767	13.063	13.359	13.655	13.950	14.246	14.541	14.835	15.129	15.423
23	Temp.	-11.36	-1.36	+ 8.64	+18.64	+28.64	+38.64	+48.64	+58.64	+68.64	+78.64	+88.64
Q Vol.	-49.6	542.2	547.3	552.4	557.4	562.5	567.6	572.7	577.8	583.0	588.1	593.2
φ	0.02390	11.965	12.251	12.536	12.820	13.105	13.389	13.672	14.238	14.521	14.803	14.803
24	Temp.	-9.69	* 0.31	+10.31	+20.31	+30.31	+40.31	+50.31	+60.31	+70.31	+80.31	+90.31
Q Vol.	-47.7	542.6	547.7	552.8	557.8	562.9	568.0	573.1	578.3	583.4	588.6	593.7
φ	0.02394	11.490	11.763	12.035	12.307	12.579	12.850	13.121	13.392	13.663	13.933	14.203
25	Temp.	-8.08	+ 1.92	+11.92	+21.92	+31.92	+41.92	+51.92	+61.92	+71.92	+81.92	+91.92
Q Vol.	-45.9	543.0	548.1	553.2	558.3	563.4	568.5	573.6	578.7	583.9	589.0	594.1
φ	0.02399	11.065	11.327	11.589	11.85	12.11	12.372	12.632	13.153	13.412	13.672	13.672
26	Temp.	-6.54	+ 3.46	+13.46	+23.46	+33.46	+43.46	+53.46	+63.46	+73.46	+83.46	+93.46
Q Vol.	-44.2	543.3	548.4	553.5	558.6	563.7	568.8	573.9	579.0	584.2	589.3	594.4
φ	0.02403	10.682	10.933	11.184	11.434	11.685	11.935	12.185	12.434	12.683	12.932	13.181
27	Temp.	-5.06	+ 4.94	+14.94	+24.94	+34.94	+44.94	+54.94	+64.94	+74.94	+84.94	+94.94
Q Vol.	-42.5	543.6	548.7	553.8	559.0	564.1	569.2	574.3	579.4	584.6	589.7	594.8
φ	0.02407	10.308	10.551	10.793	11.035	11.277	11.2441	11.559	11.760	12.001	12.242	12.482
28	Temp.	-3.59	+ 6.41	+16.41	+26.41	+36.41	+46.41	+56.41	+66.41	+76.41	+86.41	+96.41
Q Vol.	-40.8	543.9	549.0	554.1	559.3	564.4	569.5	574.6	579.8	584.9	590.1	595.2
φ	0.02410	9.969	10.203	10.446	10.669	10.902	11.204	11.367	11.599	11.831	12.063	12.294
		- .0859	1.1946	1.2181	1.2292	1.2404	1.2617	1.2724	1.2827	1.2929	1.3027	1.3103

SUPERHEATED AMMONIA (NOS. 21 to 28). — *Continued*

TABLE III

	Liquid	Vapor	110°	120°	130°	140°	150°	160°	170°	180°	190°	200°	210°	220°	300°
21	Temp.	-53.6	-14.84	+95.16	+106.16	+115.16	+122.16	+135.16	+145.16	+165.16	+185.16	+235.16	+285.16	+285.16	
Q	Vol.	0.02381	541.4	597.4	602.5	607.7	612.8	617.9	623.0	633.3	643.6	669.5	695.7	695.7	
φ		.1141	13.025	16.422	16.729	17.936	17.344	17.65	17.956	18.559	19.180	20.707	22.23	22.23	
			1.2217	1.3408	1.3506	1.3601	1.3695	1.3785	1.3875	1.4050	1.4227	1.4629	1.5004	1.5004	
22	Temp.	-51.6	-13.97	+96.93	+106.93	+116.93	+126.93	+136.93	+146.93	+166.93	+186.93	+236.93	+286.93	+286.93	
Q	Vol.	0.02386	541.7	597.7	602.8	608.0	613.1	618.2	623.4	633.7	644.0	669.9	696.1	696.1	
φ		.1096	12.470	15.717	16.011	16.304	16.597	16.890	17.192	17.766	18.350	19.866	21.26	21.26	
			1.2173	1.3363	1.3461	1.3554	1.3651	1.3738	1.3828	1.4003	1.4181	1.4383	1.4956	1.4956	
23	Temp.	-49.6	-11.36	+98.64	+108.64	+118.64	+128.64	+138.64	+148.64	+168.64	+188.64	+238.64	+288.64	+288.64	
Q	Vol.	0.02390	542.2	598.3	603.4	608.6	613.7	618.8	624.0	634.2	644.6	670.5	696.7	696.7	
φ		.1043	11.965	15.085	15.367	15.638	15.958	16.20	16.486	17.038	17.596	18.989	20.102	20.102	
			1.2139	1.3328	1.3424	1.3519	1.3614	1.3702	1.3791	1.3967	1.4143	1.4544	1.4919	1.4919	
24	Temp.	-47.7	-9.69	+100.31	+110.31	+120.31	+130.31	+140.31	+150.31	+160.31	+160.31	+160.31	+240.31	+290.31	+290.31
Q	Vol.	0.02394	542.6	598.8	603.9	609.1	614.2	619.3	624.5	634.8	645.1	671.0	697.2	697.2	
φ		.10111	11.490	14.473	14.743	15.012	15.281	15.55	15.816	16.347	16.877	18.20	19.52	19.52	
			1.2091	1.3278	1.3374	1.3469	1.3564	1.3652	1.3741	1.3917	1.4091	1.4493	1.4867	1.4867	
25	Temp.	-45.9	-8.08	+101.92	+111.92	+121.92	+131.92	+141.92	+151.92	+171.92	+191.92	+241.92	+291.92	+291.92	
Q	Vol.	0.02399	543.0	599.2	604.3	609.3	614.6	619.7	624.9	635.2	645.5	671.5	697.7	697.7	
φ		.0970	11.065	13.931	14.190	14.448	14.707	14.965	15.278	15.791	16.303	17.586	18.81	18.81	
			1.2053	1.3238	1.3334	1.3429	1.3524	1.3612	1.3702	1.3877	1.4052	1.4453	1.4826	1.4826	
26	Temp.	-44.2	-6.54	+103.46	+113.46	+123.46	+133.46	+143.46	+153.46	+173.46	+193.46	+243.46	+293.46	+293.46	
Q	Vol.	0.02403	543.3	599.5	604.7	609.8	615.0	620.1	625.3	635.6	645.9	671.9	698.2	698.2	
φ		.0932	10.682	13.449	13.677	13.915	14.173	14.42	14.67	15.163	15.657	16.890	18.12	18.12	
			1.2018	1.3200	1.3298	1.3393	1.3487	1.3574	1.3665	1.3840	1.4015	1.4416	1.4788	1.4788	
27	Temp.	-42.5	-5.06	+104.94	+114.94	+124.94	+134.94	+144.94	+154.94	+174.94	+194.94	+244.94	+294.94	+294.94	
Q	Vol.	0.02407	543.6	596.9	605.1	610.2	614.4	620.5	625.7	635.0	645.3	671.9	698.6	698.6	
φ		.0895	10.308	12.963	13.202	13.442	13.881	13.92	14.158	14.635	15.110	16.296	17.48	17.48	
			1.1982	1.3162	1.3261	1.3355	1.3450	1.3535	1.3627	1.3801	1.3977	1.4377	1.4749	1.4749	
28	Temp.	-40.8	-3.59	+106.41	+116.41	+126.41	+136.41	+146.41	+156.41	+176.41	+196.41	+246.41	+296.41	+296.41	
Q	Vol.	0.02410	543.9	600.3	605.4	610.6	615.7	620.8	626.0	636.3	646.7	672.7	699.1	699.1	
φ		.0839	9.969	12.526	12.757	12.988	13.229	13.45	13.680	14.141	14.600	15.747	16.89	16.89	
			1.1946	1.3124	1.3223	1.3316	1.3411	1.3587	1.3763	1.3938	1.4338	1.4709	1.4709	1.4709	

SUPERHEATED AMMONIA (Nos. 30 to 44)

	Liquid	Vapor	10°	20°	30°	40°	50°	60°	70°	80°	90°	100°	
30	Temp.	-37.6 Q Vol.	- 0.78 0.02418	544.5 9.344	+ 9.22 540.6	+19.22 554.7	+39.22 559.9	+49.22 565.0	+59.22 570.1	+69.22 575.3	+79.22 580.4	+89.22 585.6	
	φ	-	.0789	1.1884	9.563 1.2002	9.782 1.2118	10.00 1.2229	10.218 1.2340	10.436 1.2445	10.653 1.2553	10.871 1.2659	11.088 1.2763	
32	Temp.	-34.6 Q Vol.	+ 1.88 0.02425	545.1 8.798	+11.88 550.2	+21.88 555.3	+31.88 560.5	+41.88 565.6	+51.88 570.7	+61.88 575.9	+71.88 581.1	+81.88 586.2	
	φ	-	.0722	1.1827	9.003 1.1045	9.229 9.413	9.413 9.618	9.618 1.2169	9.822 1.2281	10.026 1.2387	10.230 1.2495	10.433 1.2600	10.636 1.2703
34	Temp.	-31.7 Q Vol.	+ 4.43 0.02432	545.6 8.302	+14.43 550.7	+24.43 555.9	+34.43 560.0	+44.43 566.2	+54.43 571.3	+64.43 576.5	+74.43 581.6	+84.43 586.8	
	φ	-	.0659	1.1772	8.407 1.1890	8.601 1.2005	8.885 1.2113	9.078 1.2224	9.271 1.2332	9.465 1.2437	9.66 1.2543	9.85 1.2646	10.042 1.2746
36	Temp.	-28.9 Q Vol.	+ 6.83 0.02438	546.2 7.870	+16.83 551.4	+26.83 556.4	+36.83 561.7	+46.83 566.8	+56.83 572.4	+66.83 577.2	+76.83 582.4	+86.83 587.5	
	φ	-	.0559	1.1722	8.054 1.1840	8.248 1.1934	8.421 1.1954	8.604 1.2062	8.789 1.2172	8.969 1.2281	9.152 1.2385	9.333 1.2491	9.515 1.2594
38	Temp.	-26.3 Q Vol.	+ 9.14 0.02445	546.6 7.478	+19.14 551.4	+29.14 556.4	+39.14 561.7	+49.14 567.2	+59.14 572.4	+69.14 576.5	+79.14 582.4	+89.14 587.5	
	φ	-	.0543	1.1675	7.654 1.1792	7.830 1.1906	8.005 1.2015	8.180 1.2125	8.355 1.2232	8.529 1.2335	8.703 1.2442	8.876 1.2545	9.049 1.2693
40	Temp.	-23.7 Q Vol.	+11.34 0.02450	547.0 7.117	+21.34 552.2	+31.34 557.4	+41.34 562.4	+51.34 567.7	+61.34 572.4	+71.34 578.1	+81.34 583.3	+90.14 588.4	
	φ	-	.0490	1.1629	7.286 1.1745	7.455 1.1860	7.624 1.1968	7.797 1.2078	7.959 1.2184	8.126 1.2288	8.293 1.2394	8.458 1.2497	8.623 1.2545
42	Temp.	-21.2 Q Vol.	+13.48 0.02457	547.5 6.797	+23.48 552.7	+33.48 557.9	+43.48 562.9	+53.48 568.2	+63.48 573.4	+73.48 578.6	+83.48 583.8	+93.48 588.4	
	φ	-	.0438	1.1580	6.959 1.1095	7.118 1.1810	7.280 1.1918	7.440 1.2028	7.600 1.2132	7.759 1.2237	7.917 1.2343	8.076 1.2445	8.234 1.2544
44	Temp.	-18.8 Q Vol.	+15.56 0.02463	547.9 6.594	+25.56 553.1	+35.56 558.3	+45.56 563.5	+55.56 568.7	+65.56 573.9	+75.56 579.1	+85.56 584.3	+95.56 589.5	
	φ	-	.0388	1.1535	6.650 1.1650	6.814 1.1765	6.968 1.1871	7.121 1.1982	7.274 1.2086	7.426 1.2190	7.577 1.2295	7.728 1.2398	7.878 1.2496

TABLE III

SUPERHEATED AMMONIA (NOS. 30 to 44).—Continued

	Liquid	Vapor	110°	120°	130°	140°	150°	160°	180°	200°	250°	300°
30	Temp.	- 0.78	+ 109.22	+ 119.22	+ 129.22	+ 139.22	+ 149.22	+ 159.22	+ 179.22	+ 190.22	+ 249.22	+ 290.22
	Q Vol.	- 544.5	544.5	601.0	611.3	616.4	621.5	627.5	637.1	647.5	673.5	699.5
	φ	0.02418	9.344	11.737	12.169	12.385	12.816	13.246	13.677	14.750	14.4270	15.82
		- .0789	1.1884	1.3059	1.3250	1.3346	1.3433	1.3521	1.3696	1.3872	1.4270	1.4641
32	Temp.	+ 1.88	+ 111.88	+ 121.88	+ 131.88	+ 141.88	+ 151.88	+ 161.88	+ 181.88	+ 201.88	+ 251.88	+ 301.88
	Q Vol.	- 34.6	545.1	601.7	606.9	612.0	617.3	622.3	627.5	648.2	674.3	700.6
	φ	0.02425	8.798	11.042	11.244	11.446	11.85	12.053	12.459	12.863	13.873	14.88
		- .0722	1.1827	1.2999	1.3097	1.3190	1.3285	1.3371	1.3459	1.3809	1.4207	1.4577
34	Temp.	+ 4.43	+ 114.43	+ 124.43	+ 134.43	+ 144.43	+ 154.43	+ 164.43	+ 184.43	+ 204.43	+ 254.43	+ 304.43
	Q Vol.	- 31.7	545.6	602.3	607.5	612.6	617.8	623.0	628.2	638.5	648.9	701.4
	φ	0.02432	8.302	1.1772	1.2940	1.426	10.617	10.808	10.999	11.381	12.143	14.04
		- .0659	1.1772		1.3131	1.3040	1.3226	1.3311	1.3400	1.3574	1.3749	1.4515
36	Temp.	+ 6.83	+ 116.83	+ 126.83	+ 136.83	+ 146.83	+ 156.83	+ 166.83	+ 186.83	+ 206.83	+ 256.83	+ 306.83
	Q Vol.	- 28.9	546.2	603.1	608.2	613.4	618.5	623.7	628.9	639.3	649.7	698.2
	φ	0.02438	7.870	1.1722	1.2887	1.059	10.239	10.420	10.861	11.163	11.524	13.30
		- .0599	1.1722		1.3077	1.2986	1.3172	1.3257	1.3344	1.3520	1.3904	1.4459
38	Temp.	+ 9.14	+ 119.14	+ 129.14	+ 139.14	+ 149.14	+ 159.14	+ 169.14	+ 189.14	+ 209.14	+ 259.14	+ 309.14
	Q Vol.	- 26.3	546.6	603.5	608.7	613.8	619.0	624.2	629.4	639.8	675.8	702.8
	φ	0.02445	7.478	1.1675	9.395	9.566	9.737	9.909	10.08	10.251	10.502	11.782
		- .0543	1.1675		1.2837	1.2935	1.3025	1.3122	1.3205	1.3293	1.3468	1.4039
40	Temp.	+ 11.34	+ 121.34	+ 131.34	+ 141.34	+ 151.34	+ 161.34	+ 171.34	+ 191.34	+ 211.34	+ 261.34	+ 311.34
	Q Vol.	- 23.7	547.0	604.0	609.2	614.3	619.5	624.7	629.9	640.4	650.4	677.0
	φ	0.02450	7.117	1.1629	8.952	9.116	9.280	9.443	9.606	10.004	10.418	11.227
		- .0490	1.1629		1.2788	1.2884	1.2975	1.3072	1.3156	1.3242	1.3417	1.3987
42	Temp.	+ 13.48	+ 123.48	+ 133.48	+ 143.48	+ 153.48	+ 163.48	+ 173.48	+ 193.48	+ 213.48	+ 263.48	+ 313.48
	Q Vol.	- 21.2	547.5	604.6	609.8	615.0	620.2	625.4	630.6	641.0	651.4	677.7
	φ	0.02457	6.797	1.1880	8.548	8.685	8.862	9.018	9.174	9.330	9.604	10.723
		- .0438	1.1880		1.2736	1.2832	1.2923	1.3019	1.3102	1.3189	1.3363	1.3930
44	Temp.	+ 15.56	+ 125.56	+ 135.56	+ 145.56	+ 155.56	+ 165.56	+ 175.56	+ 195.56	+ 215.56	+ 265.56	+ 315.56
	Q Vol.	- 18.8	547.9	605.1	610.3	615.5	620.7	625.9	631.1	641.6	652.0	678.2
	φ	0.02463	6.504	8.178	8.326	8.475	8.623	8.771	8.946	9.241	9.535	10.270
		- .0388	1.1535	1.2688	1.2785	1.2875	1.2970	1.3063	1.3140	1.3313	1.3486	1.3880

SUPERHEATED AMMONIA (Nos. 46 to 75)

	Liquid	Vapor	10°	20°	30°	40°	50°	60°	70°	80°	90°	100°
46	Temp. Q Vol. ϕ	-16.6 +0.02468 -0.0341	+17.59 548.2 6.239 1.1494	+27.59 553.4 6.388 1.1608	+37.59 558.6 6.537 1.1724	+47.59 563.8 6.685 1.1829	+57.59 569.0 6.832 1.1939	+67.59 574.2 6.978 1.2044	+77.59 579.4 7.125 1.2147	+87.59 584.6 7.270 1.2252	+107.59 589.9 7.415 1.2354	+117.59 595.1 7.559 1.2452
48	Temp. Q Vol. ϕ	-14.4 0.02474 -0.0295	+19.48 548.6 5.998 1.1454	+29.48 553.8 6.142 1.1567	+39.48 559.0 6.287 1.1683	+49.48 564.3 6.428 1.1787	+59.48 569.5 6.570 1.1896	+69.48 574.7 6.711 1.2002	+79.48 579.9 6.851 1.2106	+89.48 585.1 6.990 1.2207	+99.48 590.3 7.129 1.2312	+109.48 595.1 7.268 1.2409
50	Temp. Q Vol. ϕ	-12.2 0.02479 -0.0232	+21.29 549.0 5.776 1.1416	+31.29 554.2 5.916 1.1528	+41.29 559.5 6.054 1.1645	+51.29 564.7 6.191 1.1748	+61.29 570.0 6.328 1.1857	+71.29 575.2 6.462 1.1963	+81.29 580.4 6.597 1.2066	+91.29 585.6 6.731 1.2171	+101.29 590.8 6.864 1.2271	+111.29 596.0 6.997 1.2368
55	Temp. Q Vol. ϕ	-7.2 0.02492 -0.0150	+25.70 549.7 5.284 1.1326	+35.70 554.9 5.410 1.1347	+45.70 560.2 5.535 1.1354	+55.70 565.4 5.659 1.1356	+65.70 570.7 5.783 1.1765	+75.70 575.9 5.905 1.1868	+85.70 581.1 6.028 1.1971	+95.70 586.4 6.149 1.2076	+105.70 591.6 6.270 1.2176	+115.70 596.9 6.391 1.2274
60	Temp. Q Vol. ϕ	-2.6 0.02594 -0.0033	+29.76 550.3 4.861 1.1243	+39.76 555.6 5.008 1.1353	+49.76 560.1 5.132 1.1470	+59.76 566.1 5.268 1.1572	+69.76 571.3 5.392 1.1679	+79.76 576.6 5.435 1.1782	+89.76 581.9 5.548 1.1883	+99.76 587.1 5.660 1.1988	+109.76 592.4 5.772 1.2088	+119.76 597.6 5.883 1.2185
65	Temp. Q Vol. ϕ	+1.9 0.02515 +0.0038	+33.62 551.1 4.595 1.1166	+43.62 556.4 4.613 1.1276	+53.62 561.7 4.720 1.1390	+63.62 566.9 4.826 1.1492	+73.62 572.2 4.932 1.1598	+83.62 577.5 5.037 1.1701	+93.62 582.8 5.141 1.1802	+103.62 588.0 5.245 1.1906	+113.62 593.3 5.349 1.2005	+123.62 598.6 5.452 1.2102
70	Temp. Q Vol. ϕ	+5.9 0.02526 +0.0121	+37.22 551.6 4.197 1.1096	+47.22 556.9 4.298 1.1205	+57.22 562.2 4.398 1.1318	+67.22 567.4 4.498 1.1420	+77.22 572.7 4.596 1.1526	+87.22 578.0 4.695 1.1628	+97.22 583.3 4.792 1.1728	+107.22 588.6 4.890 1.1831	+117.22 593.8 4.986 1.1930	+127.22 599.1 5.081 1.2026
75	Temp. Q Vol. ϕ	+9.9 0.02536 +0.0200	+40.66 552.1 3.927 1.1032	+50.66 567.4 4.022 1.1141	+60.66 568.6 4.116 1.1252	+70.66 573.3 4.210 1.1354	+80.66 578.6 4.303 1.1470	+90.66 583.9 4.395 1.1562	+100.66 589.2 4.487 1.1661	+110.66 594.5 4.579 1.1762	+120.66 599.8 4.669 1.1862	+130.66 605.1 4.760 1.1957

TABLE III

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SUPERHEATED AMMONIA (Nos. 46 to 75).—Continued

	Liquid	Vapor	110°	120°	130°	140°	150°	160°	180°	200°	250°	300°
46	Temp. Q Vol. ϕ	-16.6 0.02468 - .0341 -	+17.59 548.2 6.239 1.1494	+127.59 605.5 7.946 1.2644	+137.59 610.7 7.989 1.2742	+147.59 615.9 8.132 1.2832	+157.59 621.1 8.275 1.2925	+167.59 626.3 8.417 1.3018	+177.59 631.5 8.559 1.3094	+187.59 642.0 8.843 1.3267	+217.59 654.4 9.117 1.3442	+267.59 678.7 9.833 1.3834
48	Temp. Q Vol. ϕ	-14.4 0.02474 - .0295 -	+19.48 548.6 5.998 1.1454	+129.48 605.9 7.343 1.2602	+139.48 611.1 7.681 1.2698	+149.48 616.4 7.817 1.2788	+159.48 621.6 7.954 1.2882	+169.48 626.8 8.090 1.2974	+179.48 632.0 8.232 1.3052	+199.48 642.5 8.516 1.3223	+219.48 652.9 8.800 1.3396	+269.48 679.3 9.463 1.3792
50	Temp. Q Vol. ϕ	-12.2 0.02479 - .0232 -	+21.29 549.0 5.776 1.1416	+131.29 606.4 7.260 1.2559	+141.29 611.6 7.391 1.2559	+151.29 616.9 7.522 1.2746	+161.29 622.1 7.652 1.2840	+171.29 627.3 7.782 1.2931	+181.29 632.5 8.185 1.3009	+201.29 643.0 8.452 1.3180	+221.29 653.5 8.799 1.3354	+271.29 679.9 9.103 1.3745
55	Temp. Q Vol. ϕ	-7.2 0.02492 - .0150 -	+25.70 549.7 5.284 1.1326	+135.70 607.3 6.630 1.2461	+145.70 612.6 6.749 1.2558	+155.70 617.8 6.867 1.2648	+165.70 623.1 6.985 1.2741	+175.70 628.3 7.102 1.2823	+185.70 633.5 7.223 1.2908	+205.70 644.0 7.463 1.3078	+225.70 654.5 7.690 1.3251	+255.70 680.9 8.255 1.3641
60	Temp. Q Vol. ϕ	-2.6 0.02504 - .0033 -	+29.76 550.3 4.861 1.1243	+139.76 608.1 6.104 1.2371	+149.76 613... 6.213 1.2468	+159.76 618.6 6.323 1.2556	+169.76 623.9 6.431 1.2648	+179.76 629.1 6.540 1.2731	+189.76 634.4 6.652 1.2815	+209.76 644.9 6.874 1.2985	+229.76 655.4 7.095 1.3158	+255.70 688.9 8.255 1.3546
65	Temp. Q Vol. ϕ	+ 1.9 0.02515 + .0038 +	+33.62 551.1 4.505 1.1166	+143.62 609.1 5.656 1.1286	+153.62 614.4 5.758 1.1384	+163.62 619.6 5.859 1.1471	+173.62 624.9 5.960 1.1564	+183.62 630.2 6.060 1.1644	+193.62 635.5 6.151 1.1728	+213.62 646.0 6.340 1.1898	+233.62 656.5 6.552 1.2069	+283.62 683.0 7.079 1.3456
70	Temp. Q Vol. ϕ	+ 5.9 0.02556 + .0121 +	+37.22 551.6 4.197 1.1096	+147.22 609.7 5.272 1.2209	+157.22 615.0 5.367 1.2306	+167.22 620.3 5.462 1.2395	+177.22 625.6 5.555 1.2486	+187.22 630.9 5.649 1.2566	+197.22 636.2 5.747 1.2648	+217.22 646.7 5.941 1.2818	+237.22 657.3 6.135 1.2989	+287.22 683.9 6.619 1.3372
75	Temp. Q Vol. ϕ	+ 9.9 0.02536 + .0200 +	+40.66 552.1 3.927 1.1032	+150.66 610.4 5.028 1.2137	+160.66 615.7 5.117 1.2236	+170.66 621.0 5.205 1.2323	+180.66 626.3 5.293 1.2413	+190.66 631.6 5.384 1.2492	+200.66 636.9 5.565 1.2577	+220.66 647.4 5.746 1.2744	+240.66 658.0 5.946 1.2915	+340.66 684.7 6.197 1.3296

SUPERHEATED AMMONIA (Nos. 80 to 115)

	Liquid	Vapor	10°	20°	30°	40°	50°	60°	70°	80°	90°	100°
80	Temp. Q Vol. φ	+13.6 0.02547 +	+43.88 552.7 3.006 1.0972	+63.88 558.0 3.874 1.1081	+73.88 568.7 3.903 1.1190	+83.88 574.0 4.050 1.1294	+93.88 584.6 4.137 1.1408	+103.88 594.3 4.223 1.1501	+113.88 599.9 4.300 1.1597	+123.88 595.2 4.395 1.1698	+133.88 600.6 4.480 1.1802	+143.88 605.8 4.564 1.1900
85	Temp. Q Vol. φ	+17.2 0.02556 +	+47.02 553.1 3.485 1.0917	+57.02 563.7 3.572 1.1025	+77.02 569.1 3.740 1.1134	+87.02 574.4 3.823 1.1236	+97.02 579.8 3.906 1.1350	+107.02 585.1 3.988 1.1440	+117.02 590.4 4.069 1.1538	+127.02 595.8 4.150 1.1639	+137.02 601.1 4.230 1.1737	+147.02 606.4 4.310 1.1832
90	Temp. Q Vol. φ	+20.7 0.02566 +	+49.97 553.5 3.300 1.0863	+59.97 564.2 3.385 1.0971	+69.97 568.6 3.465 1.1079	+79.97 574.9 3.544 1.1181	+89.97 580.3 3.622 1.1293	+99.97 585.6 3.700 1.1384	+109.97 590.9 3.778 1.1482	+119.97 596.3 3.855 1.1583	+129.97 601.6 3.931 1.1678	+139.97 606.9 4.007 1.1773
95	Temp. Q Vol. φ	+23.9 0.02575 +	+52.83 553.9 3.134 1.0876	+62.83 564.7 3.211 1.0920	+72.83 570.0 3.287 1.1028	+82.83 575.4 3.363 1.1128	+92.83 580.8 3.438 1.1230	+102.83 586.1 3.512 1.1332	+112.83 591.4 3.586 1.1428	+122.83 596.8 3.660 1.1528	+132.83 602.1 3.733 1.1623	+142.83 607.4 3.806 1.1718
100	Temp. Q Vol. φ	+27.1 0.02584 +	+55.58 554.1 2.980 1.0937	+65.58 559.5 3.053 1.0872	+75.58 564.9 3.125 1.0980	+85.58 570.2 3.197 1.1080	+95.58 575.6 3.268 1.1194	+105.58 581.0 3.339 1.1284	+115.58 586.4 3.409 1.1379	+125.58 591.7 3.479 1.1478	+135.58 597.1 3.549 1.1573	+145.58 602.4 3.618 1.1667
105	Temp. Q Vol. φ	+30.0 0.02592 +	+58.14 554.4 2.844 1.0767	+68.14 559.8 2.914 1.0826	+78.14 565.2 2.983 1.0934	+88.14 570.5 3.051 1.1034	+98.14 575.9 3.119 1.1146	+108.14 581.3 3.187 1.1236	+118.14 586.7 3.254 1.1331	+128.14 592.1 3.321 1.1430	+138.14 597.4 3.388 1.1524	+148.14 602.8 3.454 1.1618
110	Temp. Q Vol. φ	+33.0 0.02601 +	+60.67 554.7 2.716 1.0677	+70.67 560.1 2.783 1.0780	+80.67 565.5 2.849 1.0887	+90.67 570.9 2.915 1.0987	+100.67 576.3 2.980 1.1099	+110.67 581.7 3.045 1.1188	+120.67 587.1 3.110 1.1283	+130.67 592.5 3.174 1.1382	+140.67 597.9 3.238 1.1475	+150.67 603.3 3.302 1.1569
115	Temp. Q Vol. φ	+35.8 0.02609 +	+63.19 554.9 2.666 1.0632	+73.19 560.3 2.730 1.0734	+83.19 565.7 2.793 1.0841	+93.19 571.2 2.856 1.0942	+103.19 576.6 2.918 1.1052	+113.19 582.0 2.980 1.1142	+123.19 587.4 3.042 1.1235	+133.19 598.2 3.103 1.1333	+143.19 603.6 3.164 1.1426	+153.19 609.0 3.225 1.1520

SUPERHEATED AMMONIA (Nos. 80 to 115).—Continued

TABLE III

	Liquid	Vapor	110°	120°	130°	140°	150°	160°	170°	180°	190°	200°	210°	220°	230°	240°	250°	260°	270°	280°	290°	300°
80	Temp. Q Vol. φ	+13.6 0.02547 + .0275	+43.88 552.7 3.066 1.0672	+153.88 611.1 4.648 1.2072	+163.88 62.8 4.816 1.2255	+173.88 62.1 4.869 1.2344	+183.88 63.4 4.982 1.2422	+193.88 63.7 5.087 1.2508	+203.88 64.3 5.237 1.2674	+223.88 65.9 5.406 1.2844	+243.88 66.3 5.582 1.3224	+293.88 68.6 5.828 1.3580	+343.88 71.3 6.25									
85	Temp. Q Vol. φ	+17.2 0.02556 + .0345	+47.02 553.5 3.485 1.0917	+157.02 611.7 6.70 1.2012	+167.02 622.4 4.467 1.2107	+177.02 62.7 4.545 1.2195	+187.02 63.0 4.623 1.2282	+197.02 63.3 4.701 1.2362	+207.02 63.6 4.781 1.2446	+227.02 64.9 4.942 1.2611	+247.02 65.4 5.103 1.2779	+297.02 66.3 5.501 1.3159	+347.02 71.1 5.900 1.3514									
90	Temp. Q Vol. φ	+20.7 0.02566 + .0411	+49.97 553.5 3.300 1.0863	+159.97 612.2 617.5 4.159 1.1051	+169.97 622.9 628.2 4.309 1.2048	+179.97 62.9 4.383 1.2133	+189.97 63.5 4.458 1.2222	+199.97 63.8 4.533 1.2299	+209.97 64.5 4.602 1.2383	+229.97 65.2 4.692 1.2448	+249.97 66.0 4.842 1.2715	+299.97 66.9 5.214 1.3095	+349.97 71.3 5.586 1.3447									
95	Temp. Q Vol. φ	+23.9 0.02575 + .0476	+52.83 553.9 3.134 1.0813	+162.83 612.7 618.1 4.022 1.1091	+172.83 623.4 628.8 4.094 1.2076	+182.83 62.4 4.166 1.2104	+192.83 63.1 4.237 1.2243	+202.83 63.8 4.300 1.2328	+212.83 64.5 4.390 1.2383	+232.83 65.1 4.452 1.2489	+252.83 66.0 4.504 1.2655	+302.83 66.9 4.950 1.3033	+352.83 71.4 5.305 1.3385									
100	Temp. Q Vol. φ	+27.1 0.02584 + .0537	+55.58 554.1 2.980 1.0767	+165.58 613.1 618.5 3.755 1.1842	+175.58 623.8 629.2 3.891 1.1939	+185.58 62.8 3.949 1.2025	+195.58 63.5 4.057 1.2112	+205.58 64.2 4.199 1.2190	+215.58 64.9 4.305 1.2272	+235.58 65.6 4.452 1.2435	+255.58 66.3 4.504 1.2602	+305.58 68.6 4.950 1.2977	+355.58 71.5 5.307 1.3387									
105	Temp. Q Vol. φ	+30.0 0.02592 + .0596	+58.14 554.4 2.844 1.0722	+168.14 613.5 618.9 3.585 1.1792	+178.14 624.2 629.6 3.651 1.1888	+188.14 62.6 3.716 1.1972	+198.14 63.3 3.781 1.2060	+208.14 64.0 3.846 1.2137	+218.14 64.7 3.911 1.2222	+238.14 65.0 4.040 1.2382	+258.14 66.8 4.169 1.2547	+308.14 68.6 4.490 1.2922	+358.14 71.5 4.810 1.3272									
110	Temp. Q Vol. φ	+33.0 0.02601 + .0631	+60.67 554.7 2.716 1.0677	+170.67 614.0 3.429 1.1742	+180.67 619.4 3.492 1.1837	+190.67 62.7 3.555 1.1922	+200.67 63.1 3.618 1.2008	+210.67 63.5 3.680 1.2087	+220.67 64.0 3.753 1.2272	+240.67 65.5 3.849 1.2329	+260.67 66.2 3.975 1.2495	+310.67 68.0 4.277 1.2867	+360.67 71.6 4.604 1.3217									
115	Temp. Q Vol. φ	+35.8 0.02609 + .0706	+63.19 554.9 2.602 1.0632	+173.19 614.4 3.286 1.1692	+183.19 619.7 3.346 1.1787	+193.19 62.5 3.407 1.1872	+203.19 63.0 3.467 1.1957	+213.19 63.8 3.520 1.2034	+223.19 64.2 3.586 1.2117	+243.19 65.0 3.705 1.2277	+263.19 66.2 3.823 1.2442	+313.19 68.7 4.117 1.2812	+363.19 71.6 4.411 1.3161									

SUPERHEATED AMMONIA (Nos. 120 to 150)

	Liquid	Vapor	10°	20°	30°	40°	50°	60°	70°	80°	90°	100°
120	Temp. Q Vol. φ	+38.5 0.02618 + .0758	+65.49 555.1 2.497 1.0392	+75.49 560.5 2.559 1.0093	+85.49 560.0 2.620 1.0800	+95.49 571.4 2.741 1.1011	+105.49 576.9 2.801 1.1099	+115.49 582.3 2.860 1.1192	+125.49 587.7 2.919 1.1290	+135.49 593.1 2.979 1.1382	+145.49 598.6 3.037 1.1476	+155.49 604.0 3.096 1.1560
125	Temp. Q Vol. φ	+41.1 0.02626 + .0808	+67.81 555.2 2.490 1.0555	+77.81 560.7 2.528 1.0762	+87.81 566.1 2.577 1.0863	+97.81 571.6 2.635 1.0973	+107.81 577.0 2.693 1.1058	+117.81 582.5 2.750 1.1154	+127.81 587.9 2.808 1.1250	+137.81 593.3 2.865 1.1341	+147.81 598.8 2.921 1.1435	+157.81 604.2 2.978 1.1518
130	Temp. Q Vol. φ	+43.9 0.02634 + .0838	+70.1 555.3 2.399 1.0517	+80.1 560.8 2.366 1.0617	+90.1 566.3 2.423 1.0723	+100.1 571.7 2.480 1.0822	+110.1 577.2 2.536 1.0932	+120.1 582.7 2.592 1.1018	+130.1 588.1 2.648 1.1112	+140.1 593.6 2.704 1.1209	+150.1 599.0 2.759 1.1301	+160.1 604.5 2.814 1.1392
135	Temp. Q Vol. φ	+46.3 0.02642 + .0905	+72.2 555.5 2.227 1.0480	+82.2 561.0 2.282 1.0580	+92.2 566.5 2.338 1.0685	+102.2 572.0 2.392 1.0784	+112.2 577.5 2.337 1.0883	+122.2 583.0 2.501 1.0980	+132.2 588.4 2.554 1.1072	+142.2 593.9 2.608 1.1169	+152.2 599.3 2.661 1.1260	+162.2 604.8 2.714 1.1352
140	Temp. Q Vol. φ	+48.7 0.02649 + .0950	+74.2 555.6 2.149 1.0444	+84.2 561.1 2.203 1.0544	+94.2 566.6 2.256 1.0648	+104.2 572.1 2.309 1.0746	+114.2 577.6 2.362 1.0855	+124.2 583.1 2.415 1.0944	+134.2 588.6 2.467 1.1034	+144.2 594.1 2.519 1.1131	+154.2 599.5 2.571 1.1222	+164.2 605.0 2.624 1.1313
145	Temp. Q Vol. φ	+51.1 0.02657 + .0995	+76.3 555.7 2.076 1.0412	+86.3 561.2 2.128 1.0512	+96.3 566.7 2.180 1.0616	+106.3 572.3 2.232 1.0713	+116.3 577.8 2.284 1.0822	+126.3 583.3 2.335 1.0910	+136.3 588.8 2.386 1.1000	+146.3 594.3 2.436 1.1096	+156.3 599.8 2.486 1.1187	+166.3 605.3 2.536 1.1278
150	Temp. Q Vol. φ	+53.4 0.02664 + .1039	+78.3 555.9 2.008 1.0381	+88.3 561.4 2.050 1.0481	+98.3 567.0 2.100 1.0584	+108.3 572.5 2.160 1.0681	+118.3 578.1 2.210 1.0781	+128.3 583.6 2.259 1.0876	+138.3 589.1 2.308 1.0968	+148.3 594.6 2.357 1.1063	+158.3 600.1 2.406 1.1153	+168.3 605.6 2.454 1.1244

TABLE III

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SUPERHEATED AMMONIA (Nos. 120 to 150). — Continued

	Liquid	Vapor	110°	120°	130°	140°	150°	160°	170°	180°	190°	200°	210°	220°	300°
120	Temp.	+38.5	+65.49	+175.49	+185.49	+195.49	+205.49	+215.49	+225.49	+245.49	+265.49	+315.49	+365.49	+365.49	+365.49
	Q	+0.02618	555.1	614.8	620.1	625.5	630.8	636.2	641.6	655.5	663.3	690.2	717.4	717.4	717.4
	Vol.	+.0758	2.497	3.154	3.213	3.271	3.328	3.385	3.434	3.549	3.663	3.946	4.235	4.235	4.235
	φ		1.0592	1.1647	1.1747	1.1826	1.1912	1.1990	1.2072	1.2230	1.2392	1.2702	1.3111	1.3111	1.3111
125	Temp.	+41.1	+67.81	+177.81	+187.81	+197.81	+207.81	+217.81	+227.81	+247.81	+267.81	+317.81	+367.81	+367.81	+367.81
	Q	.02626	555.2	615.0	620.4	625.7	631.1	636.5	641.9	652.8	663.6	690.6	717.8	717.8	717.8
	Vol.	+.0868	2.400	3.034	3.090	3.146	3.202	3.257	3.312	3.422	3.531	3.803	4.074	4.074	4.074
	φ		1.0555	1.1165	1.1700	1.1784	1.1868	1.1945	1.2030	1.2185	1.2347	1.2777	1.3064	1.3064	1.3064
130	Temp.	+43.9	+70.1	+180.1	+190.1	+200.1	+210.1	+220.1	+230.1	+250.1	+270.1	+320.1	+370.1	+370.1	+370.1
	Q	+0.02634	555.3	615.3	620.7	626.0	631.4	636.8	642.2	653.1	663.9	691.0	718.2	718.2	718.2
	Vol.	+.0858	2.309	2.923	2.977	3.031	3.084	3.138	3.191	3.297	3.402	3.663	3.923	3.923	3.923
	φ		1.0517	1.1503	1.1539	1.1573	1.1825	1.1902	1.1985	1.2142	1.2302	1.2669	1.3017	1.3017	1.3017
135	Temp.	+46.3	+72.2	+182.2	+192.2	+202.2	+212.2	+222.2	+232.2	+252.2	+272.2	+322.2	+372.2	+372.2	+372.2
	Q	+0.02642	555.5	615.6	621.0	626.4	631.8	637.2	642.6	653.5	664.4	691.5	718.8	718.8	718.8
	Vol.	+.0905	2.227	2.819	2.872	2.924	2.975	3.027	3.078	3.180	3.282	3.536	3.787	3.787	3.787
	φ		1.0480	1.1532	1.1609	1.1690	1.1782	1.1869	1.1942	1.2100	1.2260	1.2635	1.2970	1.2970	1.2970
140	Temp.	+48.7	+74.2	+184.2	+194.2	+204.2	+214.2	+224.2	+234.2	+254.2	+274.2	+324.2	+374.2	+374.2	+374.2
	Q	+0.02649	555.6	615.9	621.3	626.7	632.1	637.5	642.9	653.8	664.7	691.9	719.3	719.3	719.3
	Vol.	+.0950	2.149	2.726	2.776	2.827	2.877	2.927	2.975	3.075	3.174	3.418	3.660	3.660	3.660
	φ		1.0444	1.1483	1.1575	1.1658	1.1742	1.1819	1.1902	1.2056	1.2216	1.2582	1.2924	1.2924	1.2924
145	Temp.	+51.1	+76.3	+186.3	+196.3	+206.3	+216.3	+226.3	+236.3	+256.3	+276.3	+326.3	+376.3	+376.3	+376.3
	Q	+0.02657	555.7	616.2	621.6	627.0	632.4	637.8	643.3	654.2	665.1	692.4	719.7	719.7	719.7
	Vol.	+.0995	2.076	2.636	2.685	2.734	2.783	2.832	2.876	2.974	3.069	3.304	3.538	3.538	3.538
	φ		1.0412	1.1447	1.1540	1.1632	1.1705	1.1783	1.1864	1.2020	1.2177	1.2542	1.2884	1.2884	1.2884
150	Temp.	+53.4	+78.3	+188.3	+198.3	+208.3	+218.3	+228.3	+238.3	+258.3	+278.3	+328.3	+378.3	+378.3	+378.3
	Q	+0.02664	555.9	616.5	622.0	2.598	2.646	2.693	2.740	2.881	2.973	3.201	3.425	3.425	3.425
	Vol.	+.1039	2.008	2.550	1.0381	1.1504	1.1556	1.1669	1.1749	1.1829	1.2141	1.2504	1.2884	1.2884	1.2884

SUPERHEATED AMMONIA (Nos. 155 to 200)

	Liquid	Vapor	10°	20°	30°	40°	50°	60°	70°	80°	90°	100°
155 Temp. Q Vol. ϕ	+ 55.7 + 0.02671 + .1082 +	+ 80.2 556.0 1.944 1.0348	+ 90.2 561.6 1.993 1.0448	+ 100.2 567.1 2.043 1.0550	+ 110.2 572.7 2.091 1.0648	+ 120.2 583.8 2.139 1.0748	+ 130.2 589.3 2.188 1.0841	+ 140.2 594.8 2.235 1.0933	+ 150.2 594.8 2.283 1.1028	+ 160.2 600.4 2.330 1.1119	+ 170.2 605.9 2.377 1.1208	+ 180.2 611.4 2.423 1.1291
160 Temp. Q Vol. ϕ	+ 57.8 + 0.02678 + .1122 +	+ 82.1 556.1 1.885 1.0318	+ 92.1 561.7 1.983 1.0418	+ 102.1 567.3 1.981 1.0520	+ 112.1 572.8 2.028 1.0618	+ 122.1 578.4 2.076 1.0717	+ 132.1 584.0 2.122 1.0810	+ 142.1 589.5 2.169 1.0901	+ 152.1 595.1 2.215 1.0998	+ 162.1 600.6 2.261 1.1088	+ 172.1 606.2 2.307 1.1177	+ 182.1 611.7 2.353 1.1259
165 Temp. Q Vol. ϕ	+ 60.1 + 0.02685 + .1162 +	+ 84.0 556.3 1.828 1.0288	+ 94.0 561.9 1.875 1.0388	+ 104.0 567.5 1.921 1.0489	+ 114.0 573.1 1.968 1.0588	+ 124.0 578.7 2.013 1.0687	+ 134.0 584.3 2.059 1.0778	+ 144.0 589.8 2.104 1.0870	+ 154.0 595.4 2.149 1.0966	+ 164.0 600.9 2.193 1.1056	+ 174.0 606.5 2.238 1.1145	+ 184.0 612.0 2.282 1.1228
170 Temp. Q Vol. ϕ	+ 62.2 + 0.02692 + .1201 +	+ 85.8 556.4 1.776 1.0359	+ 95.8 562.0 1.822 1.0460	+ 105.8 567.6 1.867 1.0559	+ 115.8 573.3 1.912 1.0559	+ 125.8 578.9 1.957 1.0657	+ 135.8 584.5 2.002 1.0749	+ 145.8 590.1 2.046 1.0840	+ 155.8 595.6 2.090 1.0936	+ 165.8 601.2 2.133 1.1025	+ 175.8 606.7 2.176 1.1114	+ 185.8 612.3 2.219 1.1198
180 Temp. Q Vol. ϕ	+ 66.3 + 0.02705 + .1277 +	+ 89.3 556.5 1.677 1.0204	+ 99.3 562.2 1.721 1.0304	+ 109.3 567.8 1.764 1.0405	+ 119.3 573.5 1.807 1.0504	+ 129.3 579.1 1.850 1.0601	+ 139.3 584.8 1.892 1.0694	+ 149.3 590.4 1.935 1.0784	+ 159.3 596.0 1.976 1.0880	+ 169.3 601.5 2.018 1.0968	+ 179.3 607.1 2.059 1.1057	+ 189.3 612.7 2.100 1.1141
190 Temp. Q Vol. ϕ	+ 70.3 + 0.02718 + .1348 +	+ 92.7 556.6 1.589 1.0149	+ 102.7 562.2 1.631 1.0249	+ 112.7 567.9 1.673 1.0350	+ 122.7 573.6 1.714 1.0449	+ 132.7 579.3 1.755 1.0544	+ 142.7 585.0 1.795 1.0639	+ 152.7 590.6 1.836 1.0729	+ 162.7 596.2 1.876 1.0824	+ 172.7 601.9 1.915 1.0910	+ 182.7 607.5 1.954 1.0999	+ 192.7 613.1 1.993 1.1084
200 Temp. Q Vol. ϕ	+ 74.1 + 0.02732 + .1414 +	+ 95.9 556.6 1.510 1.0097	+ 105.9 562.3 1.550 1.0197	+ 115.9 568.1 1.590 1.0299	+ 125.9 573.8 1.630 1.0397	+ 135.9 579.6 1.666 1.0492	+ 145.9 585.3 1.708 1.0587	+ 155.9 590.9 1.747 1.0677	+ 165.9 596.6 1.785 1.0772	+ 175.9 602.2 1.822 1.0858	+ 185.9 607.9 1.860 1.0947	+ 195.9 613.5 1.889 1.1032

SUPERHEATED AMMONIA (Nos. 155 to 200).—Continued

TABLE III

	Liquid	Vapor	110°	120°	130°	140°	150°	160°	170°	180°	190°	200°	210°	220°	230°	240°	250°	260°	270°	280°	290°
155	Temp. Q Vol. ϕ	+55.7 0.02671 + .1082	556.0 1.044 1.0348	+190.2 616.8 2.470 1.1378	+200.2 622.3 2.516 1.1468	+210.2 627.7 2.562 1.1548	+220.2 633.2 2.608 1.1632	+230.2 638.6 2.656 1.1711	+240.2 644.1 2.700 1.1792	+250.2 655.0 2.791 1.1946	+260.2 665.9 2.881 1.2103	+270.2 673.3 3.102 1.2466	+280.2 683.3 3.102 1.2865	+290.2 693.3 3.102 1.2865							
160	Temp. Q Vol. ϕ	+57.8 0.02678 + .1122	556.1 1.885 1.0318	+192.1 617.1 2.398 1.1346	+202.1 622.1 2.443 1.1436	+212.1 628.0 2.488 1.1517	+222.1 633.5 2.533 1.1598	+232.1 638.9 2.577 1.1678	+242.1 644.4 2.621 1.1758	+252.1 655.3 2.708 1.1913	+262.1 666.3 2.795 1.2068	+272.1 675.3 2.881 1.2429	+282.1 683.1 3.011 1.2758	+292.1 693.7 3.011 1.2758	+302.1 702.1 3.224 1.2758						
165	Temp. Q Vol. ϕ	+60.1 0.02685 + .1162	556.3 1.828 1.0288	+194.0 617.5 2.326 1.1313	+204.0 623.0 2.369 1.1401	+214.0 628.4 2.413 1.1485	+224.0 633.9 2.456 1.1566	+234.0 639.4 2.500 1.1645	+244.0 644.9 2.543 1.1725	+254.0 655.8 2.627 1.1878	+264.0 666.8 2.711 1.2033	+274.0 676.8 2.795 1.2393	+284.0 686.8 2.881 1.2733	+294.0 694.2 2.920 1.2733	+304.0 704.2 3.126 1.2733						
170	Temp. Q Vol. ϕ	+62.2 0.02692 + .1201	556.4 1.776 1.0259	+195.8 617.8 2.262 1.1281	+205.8 623.3 2.305 1.1370	+215.8 628.7 2.347 1.1452	+225.8 634.2 2.389 1.1534	+235.8 639.7 2.431 1.1612	+245.8 645.2 2.472 1.1692	+255.8 656.2 2.555 1.1846	+265.8 667.2 2.637 1.1999	+275.8 677.2 2.739 1.2359	+285.8 687.2 2.839 1.2759	+295.8 694.7 2.939 1.2759	+305.8 704.7 3.040 1.2759						
180	Temp. Q Vol. ϕ	+66.3 0.02705 + .1277	556.5 1.677 1.0204	+199.3 618.2 2.141 1.1224	+209.3 623.7 2.181 1.1313	+219.3 629.3 2.221 1.1393	+229.3 634.8 2.261 1.1474	+239.3 640.3 2.301 1.1554	+249.3 645.8 2.341 1.1634	+259.3 656.9 2.419 1.1784	+269.3 667.9 2.497 1.1937	+279.3 677.9 2.576 1.2296	+289.3 687.9 2.656 1.2296	+299.3 694.7 2.739 1.2296	+309.3 704.7 2.881 1.2296						
190	Temp. Q Vol. ϕ	+70.3 0.02718 + .1348	556.7 1.589 1.0149	+202.7 618.6 2.032 1.1169	+212.7 624.2 2.071 1.1254	+222.7 629.7 2.109 1.1335	+232.7 635.3 2.147 1.1417	+242.7 640.8 2.185 1.1497	+252.7 646.3 2.223 1.1574	+272.7 657.4 2.298 1.1727	+292.7 668.4 2.372 1.1878	+342.7 696.1 2.556 1.2234	+392.7 723.7 2.738 1.2569	+392.7 723.7 2.738 1.2569							
200	Temp. Q Vol. ϕ	+74.1 0.02732 + .1414	556.9 1.510 1.0097	+205.9 619.1 2.007 1.1117	+215.9 624.7 2.043 1.1202	+225.9 630.3 2.079 1.1282	+235.9 635.9 2.143 1.1363	+245.9 641.5 2.186 1.1442	+255.9 647.0 2.223 1.1522	+275.9 658.1 2.257 1.1672	+295.9 669.1 2.353 1.1822	+345.9 686.9 2.433 1.2177	+395.9 724.6 2.602 1.2512	+395.9 724.6 2.602 1.2512							

SUPERHEATED AMMONIA (Nos. 210 to 270)

	Liquid	Vapor	10°	20°	30°	40°	50°	60°	70°	80°	90°	100°
210 Temp.	+77.7 Q. Vol.	+99.0 0.02744 +	556.5 1.438 1.0048	+100.0 568.1 1.515 1.0252 1.0148	+110.0 573.8 1.553 1.0348	+120.0 579.6 1.591 1.0443	+130.0 585.4 1.628 1.0538	+140.0 591.1 1.665 1.0629	+150.0 596.8 1.702 1.0723	+160.0 602.4 1.738 1.0810	+170.0 608.1 1.774 1.0898	+180.0 613.8 1.810 1.0984
220 Temp.	+81.2 Q. Vol.	+102.0 0.02757 +	556.5 1.371 1.0003	+112.0 562.3 1.409 1.0103	+122.0 574.0 1.446 1.0208	+132.0 579.9 1.519 1.0303	+142.0 585.7 1.555 1.0400	+152.0 591.4 1.591 1.0493	+162.0 597.1 1.625 1.0586	+172.0 602.7 1.661 1.0678	+182.0 608.4 1.696 1.0766	+192.0 614.1 1.730 1.0854
230 Temp.	+84.6 Q. Vol.	+104.9 0.02769 +	556.4 1.312 0.9959	+114.9 566.1 1.348 1.0061	+124.9 574.0 1.384 1.0165	+134.9 579.8 1.420 1.0259	+144.9 585.7 1.455 1.0357	+154.9 591.4 1.490 1.0449	+164.9 597.2 1.524 1.0544	+174.9 602.9 1.558 1.0636	+184.9 608.7 1.592 1.0724	+194.9 614.4 1.658 1.0899
240 Temp.	+87.9 Q. Vol.	+107.7 0.02782 +	556.3 1.256 1.0001	+117.7 562.2 1.291 1.0019	+127.7 574.0 1.361 1.0123	+137.7 579.9 1.395 1.0217	+147.7 585.8 1.428 1.0316	+157.7 591.6 1.462 1.0408	+167.7 597.4 1.494 1.0504	+177.7 603.1 1.527 1.0596	+187.7 608.9 1.559 1.0686	+197.7 614.7 1.591 1.0773
250 Temp.	+91.1 Q. Vol.	+110.4 0.02794 +	556.0 1.204 0.9874	+120.4 567.9 1.239 0.9978	+130.4 577.9 1.273 1.0084	+140.4 579.8 1.306 1.0176	+150.4 585.7 1.339 1.0274	+160.4 591.5 1.372 1.0369	+170.4 597.3 1.404 1.0466	+180.4 603.1 1.436 1.0556	+190.4 608.9 1.468 1.0649	+200.4 614.7 1.530 1.0864
260 Temp.	+94.2 Q. Vol.	+113.1 0.02806 +	555.9 1.158 0.9834	+123.1 567.9 1.192 0.9939	+133.1 573.9 1.225 1.0044	+143.1 579.9 1.257 1.0139	+153.1 585.9 1.289 1.0237	+163.1 591.7 1.321 1.0334	+173.1 597.5 1.353 1.0396	+183.1 603.4 1.383 1.0522	+193.1 609.2 1.413 1.0614	+203.1 615.0 1.444 1.0702
270 Temp.	+97.2 Q. Vol.	+115.6 0.02817 +	555.7 1.113 0.9795	+125.6 567.8 1.145 0.9901	+135.6 573.8 1.176 1.0005	+145.6 579.9 1.207 1.0104	+155.6 585.9 1.238 1.0201	+165.6 591.7 1.268 1.0296	+175.6 597.6 1.298 1.0396	+185.6 603.4 1.327 1.0487	+195.6 609.3 1.356 1.0580	+205.6 615.1 1.385 1.0667

TABLE III

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SUPERHEATED AMMONIA (Nos. 210 to 270).—Continued

	Liquid	Vapor	110°	120°	130°	140°	150°	160°	170°	180°	190°	200°	210°	220°	230°	240°	250°	260°	270°
210	Temp.	+ 77.7 Q Vol. φ	+ 99.0 0.02744 + .1479	556.5 1.438 1.0048	+ 209.0 619.4 1.845 1.1069	+ 219.0 625.0 1.877 1.1153	+ 229.0 630.6 1.915 1.1233	+ 239.0 636.2 1.960 1.1314	+ 249.0 641.8 1.984 1.1393	+ 259.0 647.4 2.019 1.1473	+ 279.0 658.5 2.087 1.1623	+ 299.0 669.6 2.155 1.1773	+ 349.0 697.4 2.322 1.2123	+ 390.0 725.2 2.488 1.2458					
220	Temp.	+ 81.2 Q Vol. φ	+ 102.0 0.02757 + .1541	556.5 1.371 1.0003	+ 212.0 619.7 1.764 1.0026	+ 222.0 625.4 1.798 1.1110	+ 232.0 631.0 1.832 1.1188	+ 242.0 636.7 1.865 1.1270	+ 252.0 642.3 1.898 1.1350	+ 262.0 647.9 1.931 1.1428	+ 282.0 659.0 1.997 1.1579	+ 302.0 670.2 2.061 1.1728	+ 352.0 698.0 2.221 1.2075	+ 402.0 725.9 2.38 1.2411					
230	Temp.	+ 84.6 Q Vol. φ	+ 104.9 0.02769 + .1602	556.4 1.312 0.9959	+ 214.9 620.1 1.690 1.0987	+ 224.9 625.7 1.723 1.1068	+ 234.9 631.4 1.755 1.1147	+ 244.9 637.0 1.787 1.1229	+ 254.9 642.7 1.818 1.1309	+ 264.9 648.3 1.852 1.1384	+ 284.9 659.4 1.915 1.1537	+ 304.9 670.6 1.977 1.1682	+ 354.9 698.5 2.131 1.2029	+ 404.9 726.6 2.283 1.2364					
240	Temp.	+ 87.9 Q Vol. φ	+ 107.7 0.02782 + .1660	556.3 1.256 0.9916	+ 217.7 620.4 1.623 1.0948	+ 227.7 626.1 1.654 1.1029	+ 237.7 631.7 1.685 1.1106	+ 247.7 637.4 1.716 1.1189	+ 257.7 643.1 1.746 1.1268	+ 267.7 648.7 1.777 1.1345	+ 287.7 659.9 1.838 1.1406	+ 307.7 671.1 1.898 1.1641	+ 357.7 699.1 1.977 1.1988	+ 407.7 727.2 2.193 1.2321					
250	Temp.	+ 91.1 Q Vol. φ	+ 110.4 0.02794 + .1717	556.0 1.204 0.9874	+ 220.4 620.4 1.560 1.0914	+ 230.4 626.1 1.591 1.0994	+ 240.4 631.9 1.621 1.1071	+ 250.4 637.6 1.651 1.1152	+ 260.4 643.3 1.680 1.1233	+ 270.4 648.9 1.710 1.1307	+ 290.4 660.1 1.768 1.1462	+ 310.4 671.3 1.826 1.1604	+ 360.4 699.4 1.969 1.1948	+ 410.4 727.7 2.110 1.2244					
260	Temp.	+ 94.2 Q Vol. φ	+ 113.1 0.02806 + .1771	555.9 1.158 0.9834	+ 223.1 620.7 1.503 1.0879	+ 233.1 626.5 1.532 1.0959	+ 243.1 632.2 1.561 1.1037	+ 253.1 638.0 1.590 1.1118	+ 263.1 643.7 1.618 1.1200	+ 273.1 649.3 1.647 1.1274	+ 293.1 660.6 1.704 1.1427	+ 313.1 671.8 1.760 1.1529	+ 363.1 699.9 1.898 1.1923	+ 413.1 728.3 2.034 1.2244					
270	Temp.	+ 97.2 Q Vol. φ	+ 115.6 0.02817 + .1824	555.7 1.113 0.9795	+ 225.6 620.9 1.441 1.0848	+ 235.6 626.7 1.469 1.0930	+ 245.6 632.4 1.496 1.1005	+ 255.6 638.2 1.531 1.1087	+ 265.6 644.0 1.560 1.1169	+ 275.6 649.6 1.601 1.1243	+ 295.6 660.9 1.654 1.1396	+ 315.6 672.2 1.706 1.1537	+ 365.6 700.4 1.835 1.1880	+ 415.6 728.9 1.916 1.2210					

SUPERHEATED AMMONIA (Nos. 280 to 400)

	Liquid	Vapor	10°	20°	30°	40°	50°	60°	70°	80°	90°	100°
280	Temp.	+118.1	+128.1 561.6	+138.1 567.7	+148.1 573.7	+158.1 579.8	+168.1 585.9	+178.1 591.8	+188.1 597.7	+198.1 603.5	+208.1 609.4	+218.1 615.3
	Q Vol.	+100.2 0.02828	555.5 1.972	+104 0.9866	+135 0.9969	+166 1.0067	+196 1.0167	+226 1.0262	+256 1.0363	+314 1.0457	+342 1.0550	+370 1.0637
	φ	+ .1876	.957									.0749
290	Temp.	+122.9	+132.9 561.2	+142.9 567.4	+152.9 573.5	+162.9 579.7	+172.9 585.9	+182.9 591.8	+192.9 597.8	+202.9 603.7	+212.9 609.7	+222.9 615.6
	Q Vol.	+105.9 0.02851	555.0 0.998	+1028 0.9901	+1058 0.9998	+1117 1.0104	+1145 1.0204	+1173 1.0306	+1201 1.0409	+1229 1.0495	+1256 1.0581	+1282 1.0676
	φ	+ .1916	.986									
300	Temp.	+127.4	+137.4 560.8	+147.4 567.0	+157.4 573.3	+167.4 578.5	+177.4 585.8	+187.4 591.8	+197.4 597.8	+207.4 603.8	+217.4 609.8	+227.4 615.8
	Q Vol.	+111.3 0.02873	554.5 0.932	+961 0.990	+9726 0.9834	+9934 1.018	+1045 1.045	+1072 1.072	+1099 1.099	+1126 1.126	+1152 1.152	+1202 1.202
	φ	+ .2069	.9614									.0633
310	Temp.	+131.8	+141.8 560.5	+151.8 566.8	+161.8 562.2	+171.8 579.5	+181.8 585.9	+191.8 592.0	+201.8 598.0	+211.8 604.1	+221.8 610.1	+231.8 616.2
	Q Vol.	+116.7 0.02894	554.1 0.876	+1094 0.904	+931 0.958	+958 0.984	+984 1.010	+1036 1.036	+1062 1.062	+1086 1.086	+1110 1.104	+1135 1.135
	φ	+ .2160	.9550	+9665 0.9665	+9774 0.9774	+9879 .9879	+9985 1.0090	+10198 1.0198	+10292 1.0292	+10394 1.0394	+10483 1.0483	+10582 1.0582
320	Temp.	+135.9	+145.9 559.9	+155.9 566.3	+165.9 571.7	+175.9 579.2	+185.9 585.7	+195.9 591.8	+205.9 598.0	+215.9 604.1	+225.9 610.3	+235.9 616.4
	Q Vol.	+121.6 0.02910	553.4 0.824	+851 .9609	+877 0.902	+902 0.928	+928 0.953	+947 0.977	+1001 1.0151	+1025 1.0249	+1048 1.0354	+10545 1.0545
	φ	+ .2245	.9489									
330	Temp.	+139.8	+149.8 559.4	+159.8 565.9	+169.8 571.5	+179.8 579.0	+189.8 585.6	+199.8 591.8	+209.8 598.0	+219.8 604.2	+229.8 610.4	+239.8 616.6
	Q Vol.	+126.5 0.02938	552.8 0.777	+803 0.828	+853 0.853	+877 0.901	+890 0.925	+901 0.948	+911 1.0110	+921 1.0212	+931 1.0318	+941 1.0414
	φ	+ .2326	.9432	+9552 0.9664	+9780 .9780	+9890 1.0000	+10000 1.0110	+10110 1.0212	+10212 1.0318	+10318 1.0414	+10414 1.0514	
340	Temp.	+143.7	+153.7 558.6	+163.7 565.2	+173.7 571.8	+183.7 578.4	+193.7 585.3	+203.7 591.6	+213.7 597.9	+223.7 604.1	+233.7 610.4	+243.7 616.7
	Q Vol.	+131.1 0.02958	552.0 0.736	+661 0.761	+809 0.809	+833 0.833	+856 0.879	+876 0.901	+896 0.923	+916 0.945	+936 1.0284	+948 1.0488
	φ	+ .2404	.9376	+9501 0.9616	+9733 0.9733	+9846 0.9961	+9961 1.0071	+10071 1.0189	+10189 1.0386	+10386 1.0488	+10488 1.0582	

TABLE III

SUPERHEATED AMMONIA (Nos. 280 to 400).—Continued

	Liquid	Vapor	110°	120°	130°	140°	150°	160°	170°	180°	190°	200°	250°	300°
280 Temp. Q Vol. φ	+100.2 0.05828 +	+118.1 555.5 1.072 0.9757	+228.1 621.1 1.398 1.0819	+238.1 626.9 1.426 1.0900	+248.1 632.7 1.453 1.0977	+258.1 638.5 1.479 1.1059	+268.1 644.3 1.506 1.1140	+278.1 650.0 1.532 1.1214	+298.1 661.3 1.555 1.1369	+318.1 672.6 1.566 1.1308	+368.1 700.8 1.636 1.1841	+418.1 729.4 1.762 1.2177		
	+105.9 0.02851 +	+122.9 555.0 0.998 .9686	+232.9 621.4 1.309 1.0768	+242.9 627.3 1.335 1.0851	+252.9 633.1 1.361 1.0926	+262.9 639.0 1.386 1.1009	+272.9 644.8 1.410 1.1089	+282.9 650.5 1.437 1.1165	+302.9 661.9 1.488 1.1319	+322.9 673.3 1.536 1.1401	+372.9 705.6 1.657 1.1801	+422.9 730.4 1.777 1.2126		
	+111.3 0.05873 +	+127.4 554.5 0.932 .0014	+237.4 621.7 1.227 1.0719	+247.4 627.6 1.252 1.0802	+257.4 633.6 1.276 1.0881	+267.4 639.5 1.300 1.0963	+277.4 645.4 1.324 1.1044	+287.4 651.1 1.348 1.1118	+307.4 662.6 1.395 1.1284	+327.4 674.1 1.442 1.1417	+377.4 702.4 1.557 1.1754	+427.4 731.4 1.671 1.2074		
320 Temp. Q Vol. φ	+116.7 0.05894 +	+131.8 554.1 0.876 .9559	+241.8 622.2 1.168 1.0680	+251.8 628.1 1.181 1.0764	+261.8 634.1 1.205 1.0844	+271.8 640.0 1.227 1.0926	+281.8 646.0 1.250 1.1008	+291.8 651.8 1.273 1.1084	+311.8 663.4 1.317 1.1240	+331.8 675.0 1.360 1.1385	+381.8 703.3 1.467 1.1779	+431.8 732.5 1.579 1.2038		
	+121.6 0.05910 +	+135.9 553.4 0.824 .2245	+245.9 622.4 1.094 1.0642	+255.9 628.4 1.116 1.0729	+265.9 634.4 1.139 1.0812	+275.9 640.4 1.160 1.0896	+285.9 646.4 1.182 1.0979	+295.9 652.2 1.204 1.1057	+315.9 663.4 1.248 1.1214	+335.9 675.5 1.291 1.1359	+385.9 703.9 1.395 1.1693	+435.9 733.2 1.496 1.2008		
	+126.5 0.05938 +	+139.8 552.8 0.777 .9432	+249.8 622.6 1.037 1.0612	+259.8 628.7 1.059 1.0702	+269.8 634.7 1.080 1.0790	+279.8 640.8 1.101 1.0874	+289.8 646.8 1.122 1.0957	+299.8 652.7 1.143 1.1037	+319.8 664.4 1.184 1.1195	+339.8 676.2 1.225 1.1342	+389.8 704.5 1.325 1.1676	+439.8 734.3 1.422 1.1987		
380 Temp. Q Vol. φ	+131.1 0.05958 +	+143.7 552.0 0.736 .2404	+253.7 622.8 1.008 1.0586	+263.7 628.9 1.029 1.0676	+273.7 634.9 1.049 1.0771	+283.7 641.0 1.069 1.0858	+293.7 647.1 1.0944	+303.7 653.0 1.102 1.1026	+323.7 664.9 1.141 1.1181	+343.7 676.7 1.179 1.1328	+393.7 705.1 1.259 1.1666	+443.7 735.1 1.355 1.1971		
	+131.1 0.05958 +	+143.7 552.0 0.736 .2404	+253.7 622.8 1.008 1.0586	+263.7 628.9 1.029 1.0676	+273.7 634.9 1.049 1.0771	+283.7 641.0 1.069 1.0858	+293.7 647.1 1.0944	+303.7 653.0 1.102 1.1026	+323.7 664.9 1.141 1.1181	+343.7 676.7 1.179 1.1328	+393.7 705.1 1.259 1.1666	+443.7 735.1 1.355 1.1971		



